

**Nilkanthrao Shinde Science & Arts College,
Bhadrawati, Dist. Chandrapur (MS)**

2.5.1

**Mechanism of internal assessment is transparent
and robust**

Unit Test B.Sc. II Sem III, 21/9/2020

Timestamp	email of the student	What is mixture homogeneous?	What is solution?	What is the difference between molality & molarity?	What is normality?	Define colligative properties.
2020/09/21 10:09:42 AM GMT+5:30	Pratikshazade06@gmail.com	A homogeneous mixture is a mixture that has a same properties of its component throughout any given sample. A solution is a homogeneous mixture	A solution is a homogeneous mixture of two or more substance having uniform properties throughout.	Molality:- It is define as the number of moles of solute dissolve in 1000g of the solvent. Molarity:- It is define as the number of moles of solute dissolve in per litre of the solution.	The normality of the solution is define as the number of gram equivalent of the solute dissolve per litre of the solution.	Colligative properties define as those which depend entirely upon the number of a particle of the solute dissolve in a known volume of a given solvent and not a all upon the nature of a solute is called colligative properties.
2020/09/21 10:12:43 AM GMT+5:30	thakurtushar171@gmail.com	The reaction in which the catalyst is present in the same phase as the reactants, is called as homogeneous mixture	A solution is a homogenous mixture of two or more substances having uniform properties.	1) Molality: It is defined as the number of moles of the solute dissolve in 100kg of the solvent . 2) Molarity: it is defined as the number of moles of solute dissolved in per litre of solution .	Normality of the solution is defines as the number of the gram equivalent of solute dissolved per litre of solution.	Colligative property are those who depends entirely upon the no. of particles of solute dissolved in a known volume of a given solvent and not at all the nature of a solute.
2020/09/21 10:25:54 AM GMT+5:30	shababeda7@gmail.com	A mixture which is evenly distributed. e.g salt water is homogeneous mixture i.e. salt gets dissolved in water.	A homogeneous mixture containing of one phases and containing only two components is called as solutions. E.g solution of NaCl in water	Molality (m) = moles of solute / kg of solvent. And. Molarity = moles of solute / litre of solution.	Normality (N) = No. gram equivalent of solute / litre of solution	The change depends on the number of the solute particles in the solution, not the nature and size of solute particles is called as colligative properties.

Marks obtained

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<p>2020/09/21 10:25:35 AM GMT+5:30</p>	<p>harsbadaskar001@gmail.com</p>	<p>Homogeneous mixtures are also referred to as solutions. While we normally think of solutions as liquids, such as soft drinks and lemonade, they can actually be in the form of solids, liquids, and gases. Homogeneous mixtures can also be a combination of these forms, as in a liquid-gas mixture.</p>	<p>a solution is a special type of homogeneous mixture composed of two or more substances. In such a mixture, a solute is a substance dissolved in another substance, known as a solvent.</p>	<p>Molarity is the ratio of the moles of a solute to the total liters of a solution. The solution includes both the solute and the solvent. Molality, on the other hand, is the ratio of the moles of a solute to the kilograms of a solvent.</p>	<p>The normality of a solution is the gram equivalent weight of a solute per liter of solution. ... For example, the concentration of a hydrochloric acid solution might be expressed as 0.1 N HCl. A gram equivalent weight or equivalent is a measure of the reactive capacity of a given chemical species (ion, molecule, etc.).</p>	<p>colligative properties are those properties of solutions that depend on the ratio of the number of solute particles to the number of solvent molecules in a solution, and not on the nature of the chemical species present. ... The word colligative is derived from the Latin colligatus meaning bound together.</p>
<p>2020/09/21 10:26:04 AM GMT+5:30</p>	<p>mayuriwatkar062000@gmail.com</p>	<p>A reaction in which the catalyst is present in the same reaction as a reactant is called as homogeneous mixture.</p>	<p>A solution is a homogeneous mixture of two or more substance on molecular level.</p>	<p>Molarity-no. of moles of solute per litre of a solutions. Molality-no. of moles of solute per kilogram of solvent.</p>	<p>It is define as no. of equivalent of solute per litre of solutions.</p>	<p>Property which depends on particles of solute dissolved in give n solvent and not on chemical composition of solute.</p>
<p>2020/09/21 10:35:21 AM GMT+5:30</p>	<p>mayurgaurkar66@gmail.com</p>	<p>A homogeneous mixture is a mixture that has a same properties of its component throughout any given sample .</p>	<p>A solution is a homogeneous mixture of two or more substance having uniform properties.</p>	<p>Molality:- a) It is defined as the number of solute dissol in 100g of the solvent. B) It is doneted by m. Molarity:- a) number of mole of solute dissolved in per litter of the solution b) It is denoted by m.</p>	<p>The number of gram equivalent of the solute dissolved per liter of the solution .</p>	<p>Colligative properties defined as those which depends entirely upon the number of particles of the solute dissolved in a known volume of a given solvent and not a all upon the nature of the solute.</p>

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<p>2020/09/21 10:13:39 AM GMT+5:30</p>	<p>achalkarinaijiya 111@gmail.com</p>	<p>The reaction in which the catalyst is present in the same phase as the reactant, is called as homogenous catalysis.</p>	<p>A solution is a homogenous mixture of two or more substances having uniform properties.</p>	<p>Molarity-It is defined as the no. Of moles of the solute dissolve in 1000g of the solvents, and Molarity- It is defined as the no. Of moles of the solute dissolve per litre of the solution.</p>	<p>The Normality of the solution is defined as the no. Of gram equivalent of the solute dissolved per litre of the solution.</p>	<p>Colligative Properties are those which depend entirely upon the no. Of particles of the solute dissolved in a know volume of a given solvent and not at all upon the nature of the solute.</p>
<p>2020/09/21 10:17:24 AM GMT+5:30</p>	<p>guddikakde@gmail.com</p>	<p>Solution is homogenous if its composition is uniform through the body of solution.</p>	<p>A solution is a homogenous mixture of two or more substance having uniform properties throughout. Solutions are contain two component solute and solvent</p>	<p>The difference between molality and molarity is , in molality the number of moles of the solute dissolved in 100g of the solvent while in molarity the number of solute dissolved in per litre of solvent.</p>	<p>The normality of the solution defined as the number of gram equivalent of the dissolved per litre of the solution. Normality = Number of gram equivalent of the solute \times Volume of the solution in litres</p>	<p>Colligative properties are those which depend entirely upon the number of particles of the solute dissolved in a known volume of a given solvent and not at all upon the nature. The various colligative properties are- 1. Osmotic pressure 2. Lowering of vapour pressure 3. Elevation in boiling point 4. Depression in freezing point</p>

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<p>2020/09/21 10:37:47 AM GMT+5:30</p>	<p>Nutankorade02@gmail.com</p>	<p>A homogeneous mixture is a solid, liquid or gaseous mixture that has the same proportions of its components throughout any given sample. Conversely, a heterogeneous mixture has components in which proportions vary throughout the sample.</p>	<p>Solution, in chemistry, a homogenous mixture of two or more substances in relative amounts that can be varied continuously up to what is called the limit of solubility. The term solution is commonly applied to the liquid state of matter, but solutions of gases and solids are possible.</p>	<p>An important distinction between molality and molarity is the difference between a solution and a solvent. Molarity is the ratio of the moles of a solute to the total liters of a solution. ... Molality, on the other hand, is the ratio of the moles of a solute to the kilograms of a solvent.</p>	<p>Normality is a measure of concentration equal to the gram equivalent weight per litre of solution. Gram equivalent weight is the measure of the reactive capacity of a molecule. The solute's role in the reaction determines the solution's normality. Normality is also known as the equivalent concentration of a solution.</p>	<p>In chemistry, colligative properties are those properties of solutions that depend on the ratio of the number of solute particles to the number of solvent molecules in a solution, and not on the nature of the chemical species present. ... The word colligative is derived from the Latin colligatus meaning bound together.</p>
<p>2020/09/21 10:39:08 AM GMT+5:30</p>	<p>nikitakamatkar7@gmail.com</p>	<p>A homogeneous mixture is a mixture that has a same properties of its component throughout any given sample a solution is a homogeneous mixture are called as mixture homogeneous.</p>	<p>A solution is a homogeneous mixture of two or more than substance on molecule level</p>	<p>Molality 1. it is defined as the number of moles of solute dissolve in 1000g of the solvent 2.it is denoted by m molarity 1.it is defined as the number of moles of solute dissolve per litre of the solute 2.it is denoted by M</p>	<p>The normality of the solution is defined as the number of gram equivalent of the solute dissolved per liter of the solution. Normality=number of gram equivalent of solute/volume of solution in liters</p>	<p>Colligative properties defined as those which depends entirely upon the number of particles of the solute dissolved in a known volume of a given solvent and not all upon the nature of the solute this is called as colligative properties</p>

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2020/09/21 10:39:09 AM GMT+5:30	Prasadamrita875@gmail.com	A solution is a mixture homogeneous of two or more substance having uniform properties.	A solution is a homogeneous mixture of two or more substance .	Molarity: It is defined as the number of moles of the solute dissolved per litre of the solution. Molality: It is defined as the number of moles of the solute dissolved in 1000g of the solvent .	The normality of the solution is defined as the number of gram equivalent of the solute dissolved per litre of the solution	Colligative properties are those which depend upon the number of particle of the solute dissolved in known volume of given solvent and not at all upon the nature of the solute
2020/09/21 10:41:16 AM GMT+5:30	tusharasutkar2000@gmail.com	Homogeneous mixture is solid, liquid or gaseous mixture that has the same properties of its components throughout any given sample	Solution is a homogeneous mixture of two or more substances	Molarity:- the number of moles of solute per litre solvent. It is represented by M... Molality:- Number of moles of solute per kilogram of solvent. It is represented by m.	The number of equivalent of solute per litre of the solvent	The properties which depends on the number of moles of solute present in the solution not on the nature of the solute
2020/09/21 10:44:32 AM GMT+5:30	Pritisunilthawari26@gmail.com	It is a mixture in which the component are evenly distributed each other.	It is a homogeneous mixture of two or more substances.	Molality: molality is defined as a number of moles of solute dissolve in per kg of solvent. It is denoted by m. unit is mol/kg. Molarity: molarity is defined as a number of moles of solute dissolve in per litre of solvent. It is denoted by M. unit: mol/litre	Normality is defined as the gram equivalent weight of solute dissolve per litre of solution.	The properties which depend entirely on the number of particals of solute dissolve in given solvent and not upon chemical composition of solute are called colligative properties.

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2020/09/21 10:44:38 AM GMT+5:30	prasad.neha221 02001@gmail.com	A homogeneous mixture is a mixture in which the components are uniformly distributed throughout the mixture.	A solution is a homogeneous mixture of two or more substances on molecular level. The constituent of the mixture present in a smaller amount is called solute, and the one present in a large amount is called as solvent.	Molality is the ratio of number of moles of solute per kilogram of solvent and Molarity is the ratio of number of moles of solute per liter of solution.	Normality is defined as number of equivalent of solute per liter of the solution.	The properties which depends entirely upon number of particles of solute dissolved in given solvent and not upon chemical composition of solute are called as colligative properties.
2020/09/21 10:46:57 AM GMT+5:30	tiwarikritesh92 @gmail.com	The mixture of solid, liquid and gas which have same proportion of its component throughout the sample is called homogeneous mixture	The homogeneous mixture of two or more substance is called solution	Molarity is the ratio of the mole of a solute to the total liters of a solution. Morality is the ratio of the moles of a solute to the kilograms of a solvent.	The number of gram equivalent of the solute dissolve per litter of the solution	These are those properties of solution that depend on the ratio of the number of solute particles to the number of solvent molecules in a solution, and not on the nature of chemical species
2020/09/21 10:52:44 AM GMT+5:30	rupeshbhusari8 4@gmail.com	A homogenous mixture is a solid, liquid or gaseous mixture that has the same proportions of its components.	A solution is a homogeneous mixture of two or more substance.	Molarity 1. it is the number of moles of solute present in 1000g of the solvent. 2. It is denoted by m. 2. molarity 1. it is defined as the number of moles of solute dissolved per litre of the solution. 2. it is denoted by M.	It is the number of gram equivalent of solute per litre of the solution. It is denoted by N.	Colligative properties are those which depends upon the number of solute particles in solution and not on the nature of the solute particles are called colligative properties.

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<p>2020/09/21 11:10:04 AM GMT+5:30</p>	<p>sameer2001buy@gmail.com</p>	<p>A homogeneous mixture is a solid, liquid or gaseous mixture that has the same proportions of its components throughout any given mixture .</p>	<p>Solution is a homogeneous mixture of two or more substances. It is mixture of solute and solvent.</p>	<p>Molarity is defined as number of moles of solute dissolved in one litre volume of solution. It's SI unit is mol/ltr. Molality is defined as number of moles of solute dissolved in one Kg of solvent. It's SI unit is mol/kg.</p>	<p>Normality is defined as number of grams equivalent of solute dissolved in one litre volume of solution. Denoted by N.</p>	<p>The property which depends upon number of solute dissolved in given solvent and does not depend upon nature of solute is called as Colligative property.</p>
<p>2020/09/21 11:10:41 AM GMT+5:30</p>	<p>Pratikshachamatar@gmail.com</p>	<p>a homogeneous mixture is a mixture that has a same properties of its component throughout any given, sample a solution is a homogeneous mixture</p>	<p>A solution is a homogeneous mixture of two or more substance having uniform properties throughout</p>	<p>Molality it is defined as the number of moles of solute dissolve in 1000g of the solvent 2. It is denoted by m molarity it is defined as the number of moles of solute dissolve per litre of the solution 2.it is denoted by M</p>	<p>The normality of the solution is defined as the number of gram equivalent of the solute dissolved per litre of the solution</p>	<p>Colligative properties defined as those which depends entirely upon the number of particles of the solute dissolved in a known volume of a given solvent and not all upon the nature of the solute these is called as colligative properties</p>

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2020/09/21 11:22:51 AM GMT+5:30	Lilharekhusbu 7@gmail.com	The reaction in which the catalyst is present in the same phase as the reactant, is called as homogeneous catalysis.	Solution is a mixture of solute and solvent the component which is present in larger quantity is known as called solvent and other one is called solute.	Molality:- It is defined as moles of solute dissolved per kilogram of the solvent. It is denoted by m. Molarity:- The molarity of a solution is defined as the no. of moles of solute dissolved in one liter of solution .It is denoted by M.	The normality of the solution is defined as the no. of gram equivalent the solute dissolved per liter of the solution .	Colligative properties are those which depend upon the no. of particle of the solute dissolved in a volume of given solute and not at all upon the nature of the solute.
2020/09/21 11:40:14 AM GMT+5:30	Sapanamandal0 02@gmail.com	A mixture in which it's constituents are distributed uniformly is called as homogeneous mixture	A solution is a homogeneous mixture of two or more substance having uniform properties	Solution and a solvent	The number of gram equivalents of the solute dissolved per litre of the solution	Colligative properties are those which depend entirely upon the number of particles of the solute dissolved in a known volume of a given solvent and not at all upon the nature of the solute
2020/09/21 11:41:06 AM GMT+5:30	runal.samate1@ gmail.com	these are the type of mixture in which the component mixed are uniformly distributed throughout the mixture	homogeneous mixture of two or more substance the component not mixed and may be varied certain limit	molarity is defn as the ratio of no. of mole of solute to the volume of solution in dm ³ . it is temperature dependance quantity. and molality is defn as the ratio of mole of solute in kg to the mass of solvent in kg		colligative properties means collective properties. these properties of solution depend only no. of solute particle not on the nature of particle

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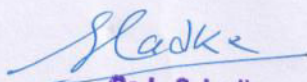
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2020/09/21 11:50:32 AM GMT+5:30	sidhanti.sonewane890@gmail.com	It is type of mixture in which the component mixed are uniformly distributed throughout the mixture. And generally homogeneous mixture means composition of same type of components.	Solution is composition of solute and solvent. Or it is homogeneous mixture of two or more pure substance whose composition can be varied.	Molarity : it is number of moles of the solute dissolved per litre of the solution. And represent by M . Molality: it is number of moles of the solute dissolved per 1kg of the solvent. It is denoted by m.	Normality : it is the number of gram equivalent of the solute dissolved per litre of solution. It is denoted by N.	Colligative properties are used to define vapour pressure osmosis freezing point and boiling points
2020/09/21 12:09:03 PM GMT+5:30	mehandichalkhure711@gmail.com	A homogeneous mixture is a solid, liquid or gaseous mixture that has the same proportion of its components Throughout any given sample.	A solution is special type of homogeneous mixture composed of two or more substance.	An important distinction between molality and molarity is the difference between a solution and solvent. molarity is the ratio of the moles of a solute to the total liters of a solution. Molality, on the other hands, is the ratio of the moles of a solute to the kilograms of a solvent.	Normality is a measure of concentration equal to the gram equivalents weight per liter of solutions. Gram equivalents weight is the measure of the reactive capacity of a molecule.	Colligative properties are those properties of solutions that depends on the ratio of the number of solute particles to the number of the solvent molecules in solutions, and not are a nature Of the chemical species present.

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<p>2020/09/21 12:33:08 PM GMT+5:30</p>	<p>Potepriti8@gmail.com</p>	<p>A solution is a homogeneous mixture</p>	<p>Two or more substance having uniform properties is called solution</p>	<p>Molality : it is defined as the number of moles of the solute dissolved in 1000g of the solvent. And molarity: it dissolved per litre of the solution</p>	<p>The normality of the solution is defined as the number of gram equivalent of the solut dissolved per litre of the solution</p>	<p>properties are those which depend entirely upon the number of particles of the solute dissolved in a known volume of a given solvent and not at all upon the nature of the solute. These properties depends upon the nature of the solvent. The various colligative properties are: 1. Lowering of vapour pressure of the solvent. 2. osmotic pressure of the solution.</p>
<p>2020/09/21 12:41:28 PM GMT+5:30</p>	<p>madkechueli@gmail.com</p>	<p>A homogeneous mixture is a mixture which solute substance are evenly spread throughout the entire the solution.</p>	<p>Solution is a homogeneous mixture of solute and solvent . Solute is a substance present in small quantity whereas solvent is a substance present in large quantity.</p>	<p>Molality :- molality difine as the number moles of solute dissolve in 1 KILOGRAM (kg) of solvent. Molarity :- molarity is define as number of moles of solute dossolve in 1 LITER of solvent.</p>	<p>Normality define as equivalent number of solute dissolve in per liter of solution.</p>	<p>The properties which entirely depends upon number of particles of solutes dissolved in given particular solvent but not upon chemical nature of solute</p>

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2020/09/21 2:07:07 PM GMT+5:30	shreyajunarkar@gmail.com	The reaction in which the catalyst is present in the same phase as the reactant is called homogeneous catalyst.	A solution is a homogeneous mixture of two or more substance on molecular level the constituent of the mixture present in a smaller amount is called solute.	1) molality = molarity is defined as the number of solute per kilogram of solvent. 2) molarity = molarity is defined as the number of moles of solute per litre of solution.	The normality of the solution is defined as the no of gram equivalents of the solute dissolved per litre of the solution.	The properties which depend upon the numbers of particles of solute dissolved in given solvent and note upon chemical composition of solute are called colligative properties.
2020/09/21 2:38:33 PM GMT+5:30	pooramadot@gmail.com	A homogeneous mixture is a solid, liquid or gaseous mixture that has the same proportions of its components throughout any given sample.	A solution is a special type of homogeneous mixture composed of two or more substance	An important distinction between a solution and a solvent . Molarity is the ratio of the mole of a solute to the total liters of solution... Molarity on the other hand is the ratio of the moles of a solute to the kilogram of a solvent	Normality is a measure of concentration equal to the gram equivalent weight per liter of solution. Gram equivalent weight is the measure of the reactive capacity of a molecule.	Colligative properties are those properties of solutions that depend on the ratio of the number of solvent molecules in a solution and not on the nature of the chemical species present.
2020/09/22 10:00:47 AM GMT+5:30	Pranadate3357@gmail.com	A smaller amount of solute is mixed with solvent is called homogeneous	The mixture of solute and solvent is called solution	Molarity is the number of moles of solute per liter of solution and molarity number of moles of solute per kilogram of solvent	Number of equivalent of solute per liter of the solution is called normality	The properties which depend entirely upon of particles of solute dissolved in solvent and chemical composition of solute are called colligative properties

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Unit Test B.Sc.II, Sem III, 01/10/2020

Timestamp	Email of student	Which of the following statements about the catalyst is true? (a) A catalyst accelerates the rate of reaction by bringing down the activation energy. (d) A catalyst makes equilibrium constant more favourable for forward reaction. (b) A catalyst does not participate in reaction mechanism.	What will be the fraction of molecules having greater energy than activation energy, E_a ? (a) K (d) 1.8×10^{-5}	The half life of the reaction having first order rate constant $K = 1.7 \times 10^{-5} \text{ s}^{-1}$ is (d) 1.8 h (c) 11.3 h	Which among the following is a false statement? RCOOR + H ₂ O → RCOOH + H ₂ O The order of reaction is this? (a) Second order (c) Half	The rate of a chemical reaction tells us about (c) how slow or fast the reaction is taking place. (d) none of the above	The reactant vs time concentration diagram for a reaction is a straight line with a negative pendulum. The reaction follows an equation for the intensity. (b) first order (a) zero order	In the reaction $2A + B \rightarrow A_2B$, if the concentration of A is doubled and that of B is halved, then the rate of the reaction will (a) increase 2 times (d) remain the same	Which of the following observations is incorrect about the order of a reaction? (a) Order of a reaction is always a whole number. (c) Order of reaction is the sum of power to express the rate of reaction to the concentration terms of the reactants. (b) The stoichiometric coefficient of reactants doesn't affect the order. (d) Order of reaction is the sum of power to express the rate of reaction to the concentration terms of the reactants.
2020/10/01 11:58:07 AM GMT+5:30	Sameer Masurkar	(a) A catalyst accelerates the rate of reaction by bringing down the activation energy. (d) A catalyst makes equilibrium constant more favourable for forward reaction. (b) A catalyst does not participate in reaction mechanism.	(a) K (d) 1.8×10^{-5}	(d) 1.8 h (c) 11.3 h	(a) Second order (c) Half	(c) how slow or fast the reaction is taking place. (d) none of the above	(b) first order (a) zero order	(a) increase 2 times (d) remain the same	(a) Order of a reaction is always a whole number. (c) Order of reaction is the sum of power to express the rate of reaction to the concentration terms of the reactants. (b) The stoichiometric coefficient of reactants doesn't affect the order. (d) Order of reaction is the sum of power to express the rate of reaction to the concentration terms of the reactants.
2020/10/01 12:01:07 PM GMT+5:30	Pratiksha zade	(a) A catalyst accelerates the rate of reaction by bringing down the activation energy. (d) A catalyst makes equilibrium constant more favourable for forward reaction. (b) A catalyst does not participate in reaction mechanism.	(a) K (d) 1.8×10^{-5}	(d) 1.8 h (c) 11.3 h	(a) Second order (c) Half	(c) how slow or fast the reaction is taking place. (d) none of the above	(b) first order (a) zero order	(a) increase 2 times (d) remain the same	(a) Order of a reaction is always a whole number. (c) Order of reaction is the sum of power to express the rate of reaction to the concentration terms of the reactants. (b) The stoichiometric coefficient of reactants doesn't affect the order. (d) Order of reaction is the sum of power to express the rate of reaction to the concentration terms of the reactants.
2020/10/01 12:02:54 PM GMT+5:30	Achal raui	(a) A catalyst accelerates the rate of reaction by bringing down the activation energy. (d) A catalyst makes equilibrium constant more favourable for forward reaction. (b) A catalyst does not participate in reaction mechanism.	(a) K (d) 1.8×10^{-5}	(d) 1.8 h (c) 11.3 h	(a) Second order (c) Half	(c) how slow or fast the reaction is taking place. (d) none of the above	(b) first order (a) zero order	(a) increase 2 times (d) remain the same	(a) Order of a reaction is always a whole number. (c) Order of reaction is the sum of power to express the rate of reaction to the concentration terms of the reactants. (b) The stoichiometric coefficient of reactants doesn't affect the order. (d) Order of reaction is the sum of power to express the rate of reaction to the concentration terms of the reactants.
2020/10/01 12:03:55 PM GMT+5:30	Parkshita Chandrapal Kakkde	(a) A catalyst accelerates the rate of reaction by bringing down the activation energy. (d) A catalyst makes equilibrium constant more favourable for forward reaction. (b) A catalyst does not participate in reaction mechanism.	(a) K (d) 1.8×10^{-5}	(d) 1.8 h (c) 11.3 h	(a) Second order (c) Half	(c) how slow or fast the reaction is taking place. (d) none of the above	(b) first order (a) zero order	(a) increase 2 times (d) remain the same	(a) Order of a reaction is always a whole number. (c) Order of reaction is the sum of power to express the rate of reaction to the concentration terms of the reactants. (b) The stoichiometric coefficient of reactants doesn't affect the order. (d) Order of reaction is the sum of power to express the rate of reaction to the concentration terms of the reactants.
2020/10/01 12:06:37 PM GMT+5:30	Tushar Thakur	(a) A catalyst accelerates the rate of reaction by bringing down the activation energy. (d) A catalyst makes equilibrium constant more favourable for forward reaction. (b) A catalyst does not participate in reaction mechanism.	(a) K (d) 1.8×10^{-5}	(d) 1.8 h (c) 11.3 h	(a) Second order (c) Half	(c) how slow or fast the reaction is taking place. (d) none of the above	(b) first order (a) zero order	(a) increase 2 times (d) remain the same	(a) Order of a reaction is always a whole number. (c) Order of reaction is the sum of power to express the rate of reaction to the concentration terms of the reactants. (b) The stoichiometric coefficient of reactants doesn't affect the order. (d) Order of reaction is the sum of power to express the rate of reaction to the concentration terms of the reactants.

Marks obtained

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2020/10/01 12:06:52 PM GMT+5:30	snehaskambhies2001@gmail.com	(a) A catalyst accelerates the rate of reaction by bringing down the activation energy. (a) K	(b) 9.7 h	c) Molecularity of a reaction may be zero or fraction. (b) Half life of a third order reaction is inversely proportional to square of initial concentration of the reactant.	(b) One	a) the reactants taking part in the reaction	(d) remain the same	(b) mol L ⁻¹ s ⁻¹	(b) The stoichiometric coefficient of the reactants doesn't affect the order	20
2020/10/01 12:10:19 PM GMT+5:30	Hilima askar	(b) A catalyst does not participate in reaction mechanism. (c) Ae-E _a /Rt	(a) 12.1 h	(a) Second order (b) Half life of a third order reaction is inversely proportional to square of initial concentration of the reactant.	(b) One	(d) none of the above	(c) decrease 2 times	(c) 1/2 mol ⁻² s ⁻¹	(b) The stoichiometric coefficient of the reactants doesn't affect the order	20
2020/10/01 12:10:34 PM GMT+5:30	Aabeda Fatema Isari Khan Pathan	(a) A catalyst accelerates the rate of reaction by bringing down the activation energy. (d) e-E _a /Rt	(c) 11.3 h	(c) Pseudo-unimolecular	(c) Half	(b) the products formed in the reaction	(a) increase 2 times	(a) s ⁻¹	(a) Order of a reaction is always a whole number	20
2020/10/01 12:17:06 PM GMT+5:30	Pritisunithawari26@gmail.com	(a) A catalyst accelerates the rate of reaction by bringing down the activation energy. (d) e-E _a /Rt	(c) 11.3 h	(c) Pseudo-unimolecular	(c) Half	(b) the products formed in the reaction	(a) increase 2 times	(b) mol L ⁻¹ s ⁻¹	9	19
2020/10/01 12:23:48 PM GMT+5:30	mayurivarkar062000@gmail.com	(a) A catalyst accelerates the rate of reaction by bringing down the activation energy. (d) e-E _a /Rt	(c) 11.3 h	(c) Pseudo-unimolecular	(b) One	(c) how slow or fast the reaction is taking place	(c) increase 2 times	(c) 1/2 mol ⁻² s ⁻¹	(a) Order of a reaction is always a whole number	20
2020/10/01 12:24:35 PM GMT+5:30	Pratishchamaskar@gmail.com	(a) A catalyst accelerates the rate of reaction by bringing down the activation energy. (d) e-E _a /Rt	(c) 11.3 h	(c) Pseudo-unimolecular	(c) Half	(b) the products formed in the reaction	(a) increase 2 times	(b) mol L ⁻¹ s ⁻¹	(d) Order can only be assessed experimentally	20
2020/10/01 12:26:09 PM GMT+5:30	Nikhilamaskar77@gmail.com	(a) A catalyst accelerates the rate of reaction by bringing down the activation energy. (d) e-E _a /Rt	(c) 11.3 h	(c) Pseudo-unimolecular	(c) Half	(b) the products formed in the reaction	(a) increase 2 times	(b) mol L ⁻¹ s ⁻¹	(d) Order can only be assessed experimentally	20

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flakes

2020/10/01 12:26:51 PM GMT+5:30	privankayeguide2001@gmail.com	(a) A catalyst accelerates the rate of reaction by bringing down the activation energy. (d) $e^{-E_a/RT}$	c) 11.3 h	c) Pseudo-unimolecular	(c) Half	(b) the products formed in the reaction	(a) zero order	(a) increase 2 times	(b) mol L ⁻¹ s ⁻¹	(d) Order can only be assessed experimentally	20
2020/10/01 12:27:04 PM GMT+5:30	Khushbu lillhare	(e) A catalyst accelerates the rate of reaction by bringing down the activation energy. (d) $e^{-E_a/RT}$	c) 11.3 h	c) Pseudo-unimolecular	(c) Half	(b) the products formed in the reaction	(a) zero order	(a) increase 2 times	(d) L mol ⁻¹ s ⁻¹	(d) Order can only be assessed experimentally	20
2020/10/01 12:36:14 PM GMT+5:30	Ritesh tiwari	(a) A catalyst accelerates the rate of reaction by bringing down the activation energy. (d) $e^{-E_a/RT}$	c) 11.3 h	(b) Unimolecular	(c) Half	(c) how slow or fast the reaction is taking place	(a) zero order	(d) remain the same	(a) s ⁻¹	(a) Order of a reaction is always a whole number	18
2020/10/01 12:36:16 PM GMT+5:30	prasad.neha22102001@gmail.com	(a) A catalyst accelerates the rate of reaction by bringing down the activation energy. (d) $e^{-E_a/RT}$	c) 11.3 h	(b) Unimolecular	(b) One	(c) how slow or fast the reaction is taking place	(c) second order	(d) remain the same	(a) s ⁻¹	(a) Order of a reaction is always a whole number	20
2020/10/01 12:41:12 PM GMT+5:30	mehandichalkhur e711@gmail.com	(a) A catalyst accelerates the rate of reaction by bringing down the activation energy. (b) K	c) 11.3 h	(b) Unimolecular	(d) Zero	(d) none of the above	(a) zero order	(a) increase 2 times	(b) mol L ⁻¹ s ⁻¹	(b) The stoichiometric coefficient of the reactants doesn't affect the order	20
2020/10/01 12:46:21 PM GMT+5:30	7385286869	(a) A catalyst accelerates the rate of reaction by bringing down the activation energy. (a) K	c) 11.3 h	(b) Unimolecular	(d) Zero	(d) none of the above	(a) zero order	(a) increase 2 times	(b) mol L ⁻¹ s ⁻¹	(b) The stoichiometric coefficient of the reactants doesn't affect the order	18
2020/10/01 1:07:30 PM GMT+5:30	rupeshhusari84@gmail.com	(a) A catalyst accelerates the rate of reaction by bringing down the activation energy. (a) K	(d) 1.8 h	c) Pseudo-unimolecular	(c) Half	(a) the reactants taking part in the reaction	(d) third order	(c) decrease 2 times	(b) mol L ⁻¹ s ⁻¹	(a) Order of a reaction is always a whole number	20

Slacks
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2020/10/01 8:27:53 PM GMT+5:30	achalkamasulija11@gmail.com	(a) A catalyst accelerates the rate of reaction by bringing down the activation energy. (d) $e-E_a/RT$	(c) 11.3 h	c) Molecularity of a reaction may be zero or fraction.	(a) Second order (c) Half	(c) how slow or fast the reaction is taking place	(a) zero order	(a) increase 2 times	(b) mol L ⁻¹ s ⁻¹	(d) Order can only be assessed experimentally
2020/10/01 7:35:39 PM GMT+5:30	Nutan sharad korade	(a) A catalyst accelerates the rate of reaction by bringing down the activation energy. (d) $e-E_a/RT$	(c) 11.3 h	c) Molecularity of a reaction may be zero or fraction.	(b) Unimolecular (c) Half	(c) how slow or fast the reaction is taking place	(d) third order	(a) increase 2 times	(a) s ⁻¹	(c) Order of reaction is the sum of power to express the rate of reaction to the concentration terms of the reactants.
2020/10/01 8:33:27 PM GMT+5:30	Prayasah652001@gmail.com	(c) A catalyst makes the reaction feasible by making $\Delta^{\ddagger}G$ more negative. (a) K	(d) 1.8 h	(d) For a first order reaction, $t_{1/2} = 0.693/K$	(b) Pseudo-unimolecular (b) One	(d) none of the above	Option 2	(a) increase 2 times	(b) mol L ⁻¹ s ⁻¹	(b) The stoichiometric coefficient of the reactants doesn't affect the order
2020/10/01 9:36:06 PM GMT+5:30	Chuvell Madke	(a) A catalyst accelerates the rate of reaction by bringing down the activation energy. (a) K	(b) 9.7 h	c) Molecularity of a reaction may be zero or fraction.	(a) Second order (b) One	(c) how slow or fast the reaction is taking place	(a) zero order	(b) increase 4 times	(b) mol L ⁻¹ s ⁻¹	(b) The stoichiometric coefficient of the reactants doesn't affect the order
2020/10/01 9:58:01 PM GMT+5:30	Harshad askar	(a) A catalyst accelerates the rate of reaction by bringing down the activation energy. (d) $e-E_a/RT$	(c) 11.3 h	c) Molecularity of a reaction may be zero or fraction.	(c) Half (b) Unimolecular	(c) how slow or fast the reaction is taking place	(d) third order	(a) increase 2 times	(a) s ⁻¹	(c) Order of reaction is the sum of power to express the rate of reaction to the concentration terms of the reactants.
2020/10/03 11:17:22 AM GMT+5:30		(a) A catalyst accelerates the rate of reaction by bringing down the activation energy. (a) K	(c) 11.3 h	c) Molecularity of a reaction may be zero or fraction.	(a) Second order (c) Half	(c) how slow or fast the reaction is taking place	(a) zero order	(a) increase 2 times	(b) mol L ⁻¹ s ⁻¹	(a) Order of a reaction is always a whole number

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Mr. Kulelip. V. Bhonyale

The screenshot shows a Google Forms interface for a 'Physics Quiz'. At the top, there is a header with the text 'a x W P=2l+2W' and a diagram of a rectangular loop in a magnetic field. Below the header, the title 'Physics Quiz' is displayed in large white letters on a dark background. The form is divided into sections. The first section is titled 'Section 1 of 2' and contains a sub-header 'Unit Test (B.Sc-2,Sem-III)'. Below this, there are two text input fields: 'MB, BE & PD statistics' and 'Full Name of Students'. A 'Short answer text' field is also present. At the bottom of the section, it says 'After section 1 - Continue to next section'. The browser's address bar shows the Google Forms URL.

The screenshot shows a Google Forms interface for an 'Untitled Section'. The section is titled 'Section 2 of 2'. It includes a sub-header 'Untitled Section' and two pieces of information: 'Max Mark - 30' and 'Time -40 min'. The first question is '1. Phase space is a _____' with four multiple-choice options: 'a) 3 Dimensional Space', 'b) 4 Dimensional Space', 'c) 5 Dimensional Space', and 'd) 6 Dimensional Space'. The second question is '2. Maxwell-Boltzmann statistics cannot be applied to _____' with two multiple-choice options: 'a) Atoms' and 'b) Molecules'. The browser's address bar shows the Google Forms URL.

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UNIT TEST- B.Sc-2 (SEM-3), (MB, BE & FD STATISTICS)																		
Timestamp	Full Name of Students	1. Phase space is a a) 1D b) 2D c) 3D d) 6 Dimensional Space	2. Maxwell-Boltzmann statistics cannot be applied to a) Lattice b) Photons c) Phonons d) 6 Dimensional Space	3. The function for the evaluation of β in M.B. distribution law is a) $\beta = kT$ b) $N/2mN^2$ c) $\beta = 1/kT$ d) $\beta = 1/kT$	4. Stirling's approximation of $\ln n!$ is a) $n \ln n$ b) $n \ln n - n$ c) $n \ln n + n$ d) $n \ln n - 1$	5. The average kinetic energy of each degree of freedom is a) kT b) $2kT$ c) $3kT$ d) $1/2 kT$	6. Bose-Einstein statistics is for a) Particles with integral spin b) Particles with half integral spin c) Fermions d) Particles with integral spin	7. Bose-Einstein statistics is for a) Particles with integral spin b) Particles with half integral spin c) Fermions d) Particles with integral spin	8. The difference between bosons and fermions wave function is a) Symmetric b) Antisymmetric c) Symmetric d) Antisymmetric	9. Bose-Einstein statistics can be applied to a) Electrons b) Photons c) Protons d) Neutrons	10. In Bose-Einstein energy state, the number of particles is a) True b) False c) True d) False	11. During Bose-Einstein condensation, the temperature is a) Ground state b) First excited state c) Ground state d) First excited state	12. Fermi-Dirac statistics is for a) Particles with half integral spin b) Particles with integral spin c) Particles with half integral spin d) Particles with integral spin	13. At $T > 0$, the probability of a state being occupied is a) True b) False c) True d) False	14. The Kinetic energy of a particle is a) True b) False c) True d) False	Fermi-Dirac Statistics a) True b) False c) True d) False	Score	
3/4/2021 13:09:24	Sameer Nasir	d) 6	b) Photons	a) $\beta = kT$	a) $n \ln n$	b) $2kT$	a) Particles with integral spin	a) Symmetric	a) Electrons	b) False	a) Ground state	a) Particles with half integral spin	b) False	b) False	a) True	b) False	6/30	
3/4/2021 13:24:20	Satish Prakash	d) 6	b) Photons	a) $\beta = 1/kT$	a) $n \ln n$	b) $2kT$	a) Particles with integral spin	a) Symmetric	a) Photons	a) True	a) Ground state	a) Particles with half integral spin	a) True	a) True	a) True	a) True	a) True	24/30
3/4/2021 13:36:13	Sajal Somnath	d) 6	b) Photons	a) $\beta = 1/kT$	a) $n \ln n$	b) $2kT$	a) Particles with integral spin	a) Symmetric	a) Photons	b) False	a) First excited state	a) Particles with half integral spin	b) False	b) False	a) True	b) False	22/30	
3/4/2021 13:50:23	Nikita Vijay	d) 6	b) Photons	a) $\beta = 1/kT$	a) $n \ln n$	b) $2kT$	a) Particles with integral spin	a) Symmetric	a) Photons	b) False	a) First excited state	a) Particles with half integral spin	b) False	b) False	a) True	b) False	24/30	
3/4/2021 16:35:56	Murali	d) 6	b) Photons	a) $\beta = 1/kT$	a) $n \ln n$	b) $2kT$	a) Particles with integral spin	a) Symmetric	a) Fermions	a) True	a) Ground state	a) Particles with half integral spin	a) True	a) True	a) True	a) True	a) True	18/30
3/4/2021 16:54:04	Meharaj	d) 6	b) Photons	a) $\beta = 1/kT$	a) $n \ln n$	b) $2kT$	a) Particles with integral spin	a) Symmetric	a) Photons	a) True	a) First excited state	a) Particles with half integral spin	a) True	a) True	a) True	a) True	a) True	8/30
3/5/2021 6:25:57	Priya Umesh	d) 6	b) Photons	a) $\beta = 1/kT$	a) $n \ln n$	b) $2kT$	a) Particles with integral spin	a) Symmetric	a) Photons	a) True	a) Ground state	a) Particles with half integral spin	a) True	a) True	a) True	a) True	a) True	24/30
3/5/2021 6:44:45	Pratiksha ellip	d) 6	b) Photons	a) $\beta = 1/kT$	a) $n \ln n$	b) $2kT$	a) Particles with integral spin	a) Symmetric	a) Photons	b) False	a) Ground state	a) Particles with half integral spin	a) True	a) True	a) True	a) True	a) True	18/30
3/5/2021 10:10:21	Priya Anand	d) 6	b) Photons	a) $\beta = 1/kT$	a) $n \ln n$	b) $2kT$	a) Particles with integral spin	a) Symmetric	a) Photons	b) False	a) Ground state	a) Particles with half integral spin	a) True	a) True	a) True	a) True	a) True	6/30
3/5/2021 12:00:58	Vishwaswar	d) 6	b) Photons	a) $\beta = 1/kT$	a) $n \ln n$	b) $2kT$	a) Particles with integral spin	a) Symmetric	a) Photons	b) False	a) Ground state	a) Particles with half integral spin	a) True	a) True	a) True	a) True	a) True	18/30
3/5/2021 20:45:37	Tushar	d) 6	b) Photons	a) $\beta = 1/kT$	a) $n \ln n$	b) $2kT$	a) Particles with integral spin	a) Symmetric	a) Photons	a) True	a) Ground state	a) Particles with half integral spin	a) True	a) True	a) True	a) True	a) True	24/30
3/5/2021 11:44:35	Dyngeshwar	d) 6	b) Photons	a) $\beta = 1/kT$	a) $n \ln n$	b) $2kT$	a) Particles with integral spin	a) Symmetric	a) Electrons	a) True	a) Ground state	a) Particles with half integral spin	a) True	a) True	a) True	a) True	a) True	10/30
3/5/2021 14:57:26	Shruti Abhay	d) 6	b) Photons	a) $\beta = 1/kT$	a) $n \ln n$	b) $2kT$	a) Particles with integral spin	a) Symmetric	a) Photons	b) False	a) Ground state	a) Particles with half integral spin	a) True	a) True	a) True	a) True	a) True	22/30
3/5/2021 15:27:35	Amisha sanjiv	d) 6	b) Photons	a) $\beta = 1/kT$	a) $n \ln n$	b) $2kT$	a) Particles with integral spin	a) Symmetric	a) Photons	a) True	a) Ground state	a) Particles with half integral spin	a) True	a) True	a) True	a) True	a) True	28/30
3/5/2021 9:33:46	Bhavana	d) 6	b) Photons	a) $\beta = 1/kT$	a) $n \ln n$	b) $2kT$	a) Particles with integral spin	a) Symmetric	a) Photons	a) True	a) Ground state	a) Particles with half integral spin	a) True	a) True	a) True	a) True	a) True	22/30
3/5/2021 9:36:49	Rahul sunil	d) 6	b) Photons	a) $\beta = 1/kT$	a) $n \ln n$	b) $2kT$	a) Particles with integral spin	a) Symmetric	a) Photons	a) True	a) Ground state	a) Particles with half integral spin	a) True	a) True	a) True	a) True	a) True	30/30
3/5/2021 9:37:37	Abhijit Anil	d) 6	b) Photons	a) $\beta = 1/kT$	a) $n \ln n$	b) $2kT$	a) Particles with integral spin	a) Symmetric	a) Fermions	a) True	a) Ground state	a) Particles with half integral spin	a) True	a) True	a) True	a) True	a) True	14/30
3/5/2021 9:38:03	Anmol Raju	d) 6	b) Photons	a) $\beta = 1/kT$	a) $n \ln n$	b) $2kT$	a) Particles with integral spin	a) Symmetric	a) Electrons	b) False	a) Ground state	a) Particles with half integral spin	a) True	a) True	a) True	a) True	a) True	16/30
3/5/2021 10:03:48	Neha sanjiv	d) 6	b) Photons	a) $\beta = 1/kT$	a) $n \ln n$	b) $2kT$	a) Particles with integral spin	a) Symmetric	a) Photons	a) True	a) Ground state	a) Particles with half integral spin	a) True	a) True	a) True	a) True	a) True	26/30
3/5/2021 10:07:18	Rohit kishor	d) 6	b) Photons	a) $\beta = 1/kT$	a) $n \ln n$	b) $2kT$	a) Particles with integral spin	a) Symmetric	a) Photons	a) True	a) Ground state	a) Particles with half integral spin	a) True	a) True	a) True	a) True	a) True	16/30
3/5/2021 10:25:46	Bhagya	d) 6	b) Photons	a) $\beta = 1/kT$	a) $n \ln n$	b) $2kT$	a) Particles with integral spin	a) Symmetric	a) Photons	a) True	a) Ground state	a) Particles with half integral spin	a) True	a) True	a) True	a) True	a) True	16/30

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3/9/2021 10:39:51	Vaishnavi Pramod Dhangar	Dimensional Space of 6	c) Photons	a) KT	c) nN	a) $\beta = 1/KT$	b) Pauli's Exclusion Principle	c) Particles with integral spin	c) Symmetric	b) Photons	a) True	a) Ground state	c) Particles with half integral spin	b) False	a) True	b) False	24 / 30
3/9/2021 10:42:49	Vaishnavi Dhanbare	Dimensional Space of 6	d) Lattice	c) KT2	a) $NnN-N$	a) $\beta = 1/KT$	b) Pauli's Exclusion Principle	b) Symmetrical Particles	c) Symmetric	b) Photons	a) True	a) Ground state	c) Particles with half integral spin	b) False	a) True	b) False	22 / 30
3/9/2021 11:44:06	Amrita Krishna Prasad	Dimensional Space of 6	c) Photons	a) KT	b) $N2/nN+2$	a) $\beta = 1/KT$	b) Pauli's Exclusion Principle	c) Particles with integral spin	c) Symmetric	b) Photons	a) True	a) Ground state	c) Particles with half integral spin	b) False	a) True	b) False	22 / 30
3/9/2021 12:00:46	Rushikesh Jagendra patkar	Dimensional Space of 6	d) Lattice	a) KT	b) $N2/nN+2$	a) $\beta = 1/KT$	b) Pauli's Exclusion Principle	c) Particles with integral spin	c) Symmetric	a) Electrons	a) True	a) Ground state	c) Particles with half integral spin	b) False	a) True	b) False	20 / 30
3/9/2021 12:18:28	Pavai sanjay bhavane	Dimensional Space of 6	a) Photons	a) KT	a) $NnN-N$	a) $\beta = 1/KT$	b) Pauli's Exclusion Principle	a) Continuous	a) Continuous	b) Photons	a) True	b) First excited state	b) Distinguishable particles	b) False	a) True	b) False	14 / 30
3/9/2021 12:40:10	Sujata Kakte	Dimensional Space of 6	b) Molecules	c) KT2	d) $NnN-N$	a) $\beta = 1/KT$	b) Pauli's Exclusion Principle	c) Particles with integral spin	c) Symmetric	c) Fermions	a) True	a) Ground state	c) Particles with half integral spin	b) False	a) True	b) False	20 / 30
3/9/2021 13:00:24	Nirnak Ramesh and Kanale	Dimensional Space of 6	c) Photons	c) KT2	a) $NnN-N$	b) $\beta = -1/KT$	b) Pauli's Exclusion Principle	a) Distinguishable particles	c) Symmetric	b) Photons	b) False	a) Ground state	c) Particles with half integral spin	a) True	a) True	b) False	22 / 30
3/9/2021 13:07:47	Bhushan L. Eschale	Dimensional Space of 6	c) Photons	c) KT2	b) $N2/nN+2$	c) $\beta = KT$	b) Pauli's Exclusion Principle	d) Particles with integral spin	c) Symmetric	b) Photons	b) True	a) Ground state	c) Particles with half integral spin	b) False	a) True	b) False	26 / 30
3/9/2021 13:54:12	Achut dhanraj Coye	Dimensional Space of 6	c) Photons	c) KT2	a) $NnN-N$	a) $\beta = 1/KT$	b) Pauli's Exclusion Principle	c) Particles with integral spin	c) Symmetric	d) Protons	a) True	a) Ground state	c) Particles with half integral spin	b) False	a) True	b) False	28 / 30
3/9/2021 18:47:56	Madhuri Jhondale	Dimensional Space of 6	c) Photons	c) KT2	c) nN	c) $\beta = KT$	b) Pauli's Exclusion Principle	c) Particles with integral spin	c) Symmetric	b) Photons	a) True	a) Ground state	c) Particles with half integral spin	b) False	a) True	b) False	22 / 30
3/10/2021 11:13:39	Kiran dattu satpute	Dimensional Space of 6	c) Photons	a) KT	c) nN	c) $\beta = KT$	b) Pauli's Exclusion Principle	c) Particles with integral spin	c) Symmetric	c) Fermions	a) True	a) Ground state	c) Particles with half integral spin	b) False	a) True	b) False	18 / 30
3/12/2021 8:50:15	Sohani y Sorewara	Dimensional Space of 6	c) Photons	c) KT2	a) $NnN-N$	a) $\beta = 1/KT$	b) Pauli's Exclusion Principle	c) Particles with integral spin	c) Symmetric	b) Photons	a) True	a) Ground state	c) Particles with half integral spin	b) False	a) True	b) True	28 / 30

P. Chougale
(Mr. K. V. Chougale)

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Mr. Kuldip V. Bhongale

My Drive - Google Drive x Copy of Unit Test (B.Sc-1 Sem-I) x

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Physics Quiz

Section 1 of 2

Unit Test (B.Sc-1, Sem-I)

Dynamics of Rigid Body & Rotational Motion

Name

Short answer: text

After section 1 Continue to next section

My Drive - Google Drive x Copy of Unit Test (B.Sc-1 Sem-I) x

docs.google.com/forms/d/1blOau8tma1bUgMu9H5FAY268h4q3hvt2LFGS110g/edit

All The Best.

Max Mark - 30 Time - 40 min

Section 2 of 2

1. Analogue of mass in rotational motion is _____

- a) Moment of inertia
- b) Angular momentum
- c) Gyration
- d) Angular acceleration

2. Moment of inertia of an object does not depend upon _____

- a) Mass of object
- b) Mass of distribution

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30/02/2021 12:40:15 Singh	Sandeep Suresh	a) Moment of inertia	a) Mass of object	d) Newton's first law of motion	c) $zz - bx = IY$	(b) m^2	A) Torque= Moment of inertia X angular acceleration	A) $(25) \text{M}^2 \text{S}^{-2}$	B) $(\text{M}^2 \text{S}^{-1})$	C) Energy	B) Angular impulse	A) $1 \text{Kg m}^2 \text{sec}^{-2}$	a) Isotropy of space	A) Rotational Inertia	A) 1639
31/02/2021 7:10:35	Gaurav	a) Moment of inertia	d) Axis of rotation	b) Newton's third law of motion	c) $zz - bx = IY$	(b) m^2	A) Torque= Moment of inertia X angular acceleration	B) $(25) \text{M}^2 \text{S}^{-2}$	C) $(\text{M}^2 \text{S}^{-1})$	A) Pressure	B) Angular impulse	C) $\text{Kg m}^2 \text{sec}^{-2}$	b) Homogeneity of time	A) Rotational Inertia	A) 1639
31/02/2021 17:57:58	Atis	a) Moment of inertia	a) Angular velocity	d) Newton's first law of motion	c) $zz - bx = IY$	(b) m^2	C) Moment of acceleration	A) $(25) \text{M}^2 \text{S}^{-2}$	B) $(\text{M}^2 \text{S}^{-1})$	C) Energy	B) Angular impulse	C) $\text{Kg m}^2 \text{sec}^{-2}$	a) Isotropy of space	B) Translational Inertia	D) Both

Shweta
(Mr K.V. Shweta)

Shakke

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Mr. Kuldeep V. Bhongale

The screenshot shows a Google Forms interface for a 'Physics Quiz'. At the top, there is a header with the text 'Physics Quiz' and a diagram of a vector triangle with sides labeled 'a', 'b', and 'c', and an angle 'θ'. Below the header, it says 'Section 1 of 2' and 'Unit Test (B.Sc-3,Sem-V)'. The topic is 'Nucleus and its stability'. There is a text input field for 'Full Name of Students' and a 'Short answer text' field. At the bottom, it says 'After section 1 Continue to next section'.

The screenshot shows a Google Forms interface for a quiz titled 'All The Best'. It is 'Section 2 of 2'. The 'Max Mark' is 30 and the 'Time' is 40 min. The first question is '1. The atomic number is equivalent to which of the following?'. The options are: A) The number of neutrons in the atom, B) The number of protons in the atom, C) The number of nucleons in the atom, and D) None of the above. The second question is '2. Which of the following is correct for the number of neutrons in the nucleus?'. The options are: A) $N = A - Z$ and B) $N = Z - A$.

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3/5/2021 23:14:20	Karan Ratan Durgu	D) None of the above	A) It is an attractive force between electrons and protons in an atom. B) $N = Z$ C) $N = 2Z$ D) $N = A$	A) have the same number of protons and electrons, but a different number of neutrons. B) the amount of energy required to break a nucleus apart into protons and neutrons. C) transformed into visible light. D) absorbed by the nucleus which then breaks it apart.	Increased by 1	B) increase thermal velocities C) β^+ D) neutrons to velocities	b) slows down neutrons to thermal velocities c) β^+ d) absorbs neutrons to thermal velocities	c) Spontaneous fission	b) Nuclear fission	c) Materials with high nuclear number	14/30
3/6/2021 8:26:59	Ankita Reju Jewarte	B) The number of protons in the atom.	A) It is a strong, short-range, attractive force between electrons and protons in an atom. B) $N = A$ C) $N = Z$ D) $N = 2Z$	A) have the same number of protons and electrons, but a different number of neutrons. B) the amount of energy required to break a nucleus apart into protons and neutrons. C) transformed into visible light. D) absorbed by the nucleus which then breaks it apart.	Doesn't change	D) decrease thermal velocities C) β^+ D) neutrons to velocities	b) slows down neutrons to thermal velocities c) β^+ d) absorbs neutrons to thermal velocities	a) Nuclear fission	a) Nuclear fission	c) Materials with high nuclear number	28/30
3/6/2021 9:52:18	Dipak Vasant Walikar	B) The number of protons in the atom.	A) It is an attractive force between electrons and protons in an atom. B) $N = A$ C) $N = Z$ D) $N = 2Z$	A) have the same number of protons and electrons, but a different number of neutrons. B) the amount of energy required to break a nucleus apart into protons and neutrons. C) transformed into visible light. D) absorbed by the nucleus which then breaks it apart.	Increased by 1	B) increase thermal velocities C) β^+ D) neutrons to velocities	b) slows down neutrons to thermal velocities c) β^+ d) absorbs neutrons to thermal velocities	a) Nuclear fission	a) Nuclear fission	c) Materials with low atomic number	16/30
3/6/2021 17:37:58	Hemlata bhaskar patil	B) The number of protons in the atom.	A) It is an attractive force between electrons and protons in an atom. B) $N = A$ C) $N = Z$ D) $N = 2Z$	A) have the same number of protons and electrons, but a different number of neutrons. B) the amount of energy required to break a nucleus apart into protons and neutrons. C) transformed into visible light. D) absorbed by the nucleus which then breaks it apart.	Doesn't change	D) decrease thermal velocities C) β^+ D) neutrons to velocities	b) slows down neutrons to thermal velocities c) β^+ d) absorbs neutrons to thermal velocities	a) Nuclear fission	a) Nuclear fission	c) Materials with low atomic number	14/30
3/6/2021 20:19:57	Prajakta Rajkumar Patil	B) The number of protons in the atom.	A) It is a strong, short-range, attractive force between electrons and protons in an atom. B) $N = A$ C) $N = Z$ D) $N = 2Z$	A) have the same number of protons and electrons, but a different number of neutrons. B) the amount of energy required to break a nucleus apart into protons and neutrons. C) transformed into visible light. D) absorbed by the nucleus which then breaks it apart.	Increased by 2	D) decrease thermal velocities C) β^+ D) neutrons to velocities	b) slows down neutrons to thermal velocities c) β^+ d) absorbs neutrons to thermal velocities	a) Nuclear fission	a) Nuclear fission	c) Materials with low atomic number	14/30
3/6/2021 23:08:40	Ankita Kalita Utkar	B) The number of protons in the atom.	A) It is a strong, short-range, attractive force between electrons and protons in an atom. B) $N = A$ C) $N = Z$ D) $N = 2Z$	A) have the same number of protons and electrons, but a different number of neutrons. B) the amount of energy required to break a nucleus apart into protons and neutrons. C) transformed into visible light. D) absorbed by the nucleus which then breaks it apart.	Decreased by 1	D) decrease thermal velocities C) β^+ D) neutrons to velocities	b) slows down neutrons to thermal velocities c) β^+ d) absorbs neutrons to thermal velocities	a) Nuclear fission	a) Nuclear fission	c) Materials with low atomic mass	24/30
3/7/2021 18:10:44	Vishwaji Vithoba Tonge	B) The number of protons in the atom.	A) It is a strong, short-range, attractive force between electrons and protons in an atom. B) $N = Z$ C) $N = A$ D) $N = 2Z$	A) have the same number of protons and electrons, but a different number of neutrons. B) the amount of energy required to break a nucleus apart into protons and neutrons. C) transformed into visible light. D) absorbed by the nucleus which then breaks it apart.	Increased by 1	D) decrease thermal velocities C) β^+ D) neutrons to velocities	b) slows down neutrons to thermal velocities c) β^+ d) absorbs neutrons to thermal velocities	a) Nuclear fission	a) Nuclear fission	c) Materials with low atomic number	20/30
3/7/2021 18:14:35	Bhushan ganpat baware	B) The number of protons in the atom.	A) It is an attractive force between electrons and protons in an atom. B) $N = Z$ C) $N = A$ D) $N = 2Z$	A) have the same number of protons and electrons, but a different number of neutrons. B) the amount of energy required to break a nucleus apart into protons and neutrons. C) transformed into visible light. D) absorbed by the nucleus which then breaks it apart.	Increased by 2	D) decrease thermal velocities C) β^+ D) neutrons to velocities	b) slows down neutrons to thermal velocities c) β^+ d) absorbs neutrons to thermal velocities	a) Nuclear fission	a) Nuclear fission	c) Materials with low atomic number	16/30
3/10/2021 11:45:58	Ponina Sambhaj Thare	D) None of the above	A) It is a much weaker than the electromagnetic force. B) $N = Z$ C) $N = 2Z$ D) $N = A$	A) have the same number of protons and electrons, but a different number of neutrons. B) the amount of energy required to break a nucleus apart into protons and neutrons. C) transformed into visible light. D) absorbed by the nucleus which then breaks it apart.	Increased by 2	D) decrease thermal velocities C) β^+ D) neutrons to velocities	b) slows down neutrons to thermal velocities c) β^+ d) absorbs neutrons to thermal velocities	a) Nuclear fission	a) Nuclear fission	c) Materials with low atomic number	14/30

Pradyumn
(M.V. K. V. Shivapule)

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Nilkanth Rao Shinde Sci And Arts college
Bhadrawati (2020-21)

Unit Test

Bsc - I (sem - I)

Sub: Mathematics (Paper - II)

Mark - 20

Time: 30 min

Que 1. Prove that if limit of a function $f(x, y)$
 $(x, y) \rightarrow (x_0, y_0)$ exists, then it is unique $\rightarrow 2$

Que 2. Prove that $\lim_{(x, y) \rightarrow (x_0, y_0)} (3x - 2y) = 14$ $\rightarrow 2$

Que 3. If $u = F(x - y, y - z, z - x)$ then

prove that $\frac{\delta u}{\delta x} + \frac{\delta u}{\delta y} + \frac{\delta u}{\delta z} = 0$ $\rightarrow 2$

Que 4. If a real valued function $f(x, y)$ is
continuous at $P_0(x_0, y_0)$ then there is
a neighbourhood of P_0 in which $f(x, y)$ is
(a) Unbounded (b) bounded
(c) convergent (d) divergent $\rightarrow 2$

Que 5. If $z = f(x, y)$ has a differential
 $dz = a \Delta x + b \Delta y$ at (x_0, y_0) then
 $f_x(x_0, y_0)$, $f_y(x_0, y_0)$ exists $\rightarrow 2$
(a) is unique (b) is ~~not~~ continuous
(c) is unique (d) None of these $\rightarrow 2$

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B.Sc.-I (sem-I)

UNIT TEST *Result*
Paper-II(Differential Calculus and Trigonometry)
Sub:-MATHEMATICS

Sr.no	Name of Students	Marks(20)
1	Gitika Durge	14
2	Divya v. vyvahare	13
3	prasad s vidhate	11
4	Atik B. Kshhirsagar	11
5	Supriya v. Balpane	13
6	Pallavi S. Gowardipe	12
7	Priti G. yelane	15
8	Saniya s. sheikh	16
9	Darshana D. Nagore	15
10	Achal R. Jidgalwar	16
11	Prachi S. Bagde	12
12	Prachi H. Kakde	14
13	Dipali S. Boinwar	13
14	Alishanaaz A. sheikh	16
15	Janvi P. Nimbalkar	15
16	Tumeshwari S Jiwtode	12
17	Kunal S. donge	14
18	Achal a. Awale	16
19	Sahil V. Nakshine	15
20	Khushi m.Kharwade	17
21	Saifaddin sayyad	15
22	Krupali A. Tonge	16
23	Komal D. Shivarkar	14
24	Ajay C. doppala	12
25	Priti V. Karekar	13
26	Vibha s. Tated	18
27	Hitashi M. Thegane	15
28	Tejas d. Telang	14
29	Ashwariya R. Tiwari	13
30	Komal c. thakare	14
31	Aishwariy R. Pijdurkar	15
32	Prajot N. Tangapendi	12
33	Anushka R. Jawade	18
34	Nikita V. Jambhulkar	16
35	Arti S. Kakade	14
36	Pranjali G. Raipure	13
37	Iramnaj S. Ali	12
38	Mahesh C. bawane	14
39	Prajakta N. bodhane	13
40	Mahek M. Sayyad	14
41	Mayuri D.Bodhe	15
42	Rahul S. Masalkar	12
43	Pratik J. Daogaokar	13
44	Rajeshwari S. Sundargiri	12

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45	Sairamshi Y. Allakatla	10
46	Keshav anil Kumar	10
47	Nikanksha P. Patil	12
48	Hitesh A. Dahule	13
49	Vipul . Nande	15
50	Pallavi S. Kale	13
51	Rajeshri S. Dixit	12
52	Swati S. Kamble	12
53	Dipak R. Soyam	11

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Bhadrawati (2020-21)

Unit Test

BSc-I (Sem-II)

Sub: Mathematics (Paper-I)

Mark - 20

Time: 30 min

Que 1:- solve any two. → ⑧

(A) solve $(1-x^2)(1-y)dx = xy(1-y)dy$

(B) Find the orthogonal trajectory of
 $e^n = a^n \cos n\theta$.

(C) solve $(x^2+y^2+2x)dx + 2ydy = 0$

Que 2:-

(A) If $\frac{d^2y}{dx^2} + x^2 \frac{dy}{dx} - y \sin x = 0$ is

(a) partial differential equation

(b) Ordinary differential equation

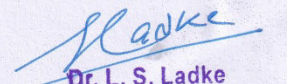
(c) None of these → ②

(B) If $\frac{dy}{dx} + \frac{x^2+3y^2}{3x^2+y^2} = 0$ is

(a) Homogeneous DE

(b) Nonhomogeneous DE

(c) None of these → ②


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① If $\mu(x, y)$ is an integrating factor of $Mdx + Ndy$, then μ satisfies the partial DE

① $N\mu_x + M\mu_y - \mu(N_x + M_y) = 0$

② $N\mu_x - M\mu_y - \mu(N_x - M_y) \neq 0$

③ $N\mu_x - M\mu_y + \mu(N_x - M_y) = 0$

④ $N\mu_x + M\mu_y - \mu(N_x + M_y) \neq 0 \rightarrow \textcircled{2}$

② If the DE $Mdx + Ndy = 0$ is homogeneous and $Mx + Ny \neq 0$ then integrating factor is

① $\frac{1}{Nx - My}$ ② $\frac{1}{Mx - Ny}$ ③ $\frac{1}{Mx + 1}$ ④ $\frac{1}{Mx + Ny} \rightarrow \textcircled{E}$

③ If $P = I_n(px - y)$ its primitive is

① $y = cx - e^c$ ② $y = cx^2 - e^{-c}$

③ $y = -cx - e^c$ ④ $y = -cx^2 + e^c \rightarrow \textcircled{2}$

④ Bernoulli's equation is

① $y' + Py = Q$

② $y' + Py = y^n Q$

③ $y' + Py = 0$

④ none of these $\rightarrow \textcircled{2}$

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B.Sc.-I (sem-II)

UNIT TEST *Result*

Paper-I(Ordinary DE and Difference Equations)

Sub:-MATHEMATICS

Sr.no	Name of Students	Marks(20)
1	Achal R. Jidgalwar	14
2	Aishwariy R. Pijdurkar	13
3	Ajay C. Doppala	12
4	Alishanaaz A. sheikh	12
5	Anushka R. Jawade	18
6	Arti S. Kakade	12
7	Ashwariya R. Tiwari	15
8	Atik B. Kshhirsagar	16
9	Darshana D. Nagpure	15
10	Dipak R. Soyam	16
11	Dipali S. Boinwar	12
12	Divya v. vyvahare	14
13	Gitika Durge	13
14	Hitashi M. Thegane	16
15	Hitesh A. Dahule	15
16	Janvi P. Nimbalkar	12
17	Keshav anil Kumar	14
18	Khushi m.Kharwade	18
19	Komal c. Thakur	15
20	Komal D. Shivarkar	11
21	Krupali A. Tonge	15
22	Kunal S. donge	16
23	Mahek M. Sayyad	14
24	Mayuri D.Bodhe	12
25	Nikanksha P. Patil	13
26	Pallavi S. Gowardipe	14
27	Pallavi S. Kale	15
28	Prachi H. Kakde	14
29	Prachi S. Bagde	13
30	Prajakta N. bodhane	14
31	Pranjali G. Raipure	15
32	prasad S. vidhate	12
33	Priti G. yelane	14
34	Priti V. Karekar	16
35	Priya R. Ambilkar	14
36	Rahul S. Masalkar	13

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37	Rajeshri S. Dixit	12
38	Rajeshwari S. Sundargiri	14
39	Sahil V. Nakshine	13
40	Saifaddin sayyad	16
41	Saniya s. sheikh	15
42	Supriya v. Balpane	12
43	Tejas d. Telang	13
44	Tumeshwari S Jiwtode	12
45	Vibha s. Tated	18
46	Vipul . Nande	12
47	Sameer R. Yadav	12
48	Rachi Katker	13
49	Krunali A Ghugul	15
50	Keshav anil Kumar	12

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UNIT TEST
BSc - II (Sem - III)

Sub :- Mathematics (Paper - I)

mark - 20

Time 30 min

Que 1:- A sequence can have at most one limit
or if $\lim S_n$ exists, it must be unique

Que 2:- Find the limit of the sequence $\{S_n\}$

$$S_n = \frac{1}{n^2+1} + \frac{1}{n^2+2} + \dots + \frac{1}{n^2+n}$$

Que 3:- Show that the sequence of decimal digits is convergent.

Que 4:- Every convergent sequence of real numbers is

- (a) Cauchy's sequence
- (b) divergent sequence
- (c) unbounded sequence
- (d) convergent sequence

Que 5:- If $\lim S_n = +\infty$, the sequence $\{S_n\}$ is

- (a) convergent
- (b) divergent
- (c) limit exist
- (d) None of above

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Que 6: If monotone sequence of real numbers is bounded it and only if

(a) Unique

(b) Divergent

(c) Convergent

(d) None of these

Que 7: - select the correct answer

(i) $\langle \frac{1}{n} \rangle$ is a bounded sequence

(ii) $\langle \frac{1}{n} \rangle$ is a Cauchy sequence

(a) (i) is false

(b) (ii) is false

(c) (i) & (ii) are false

(d) (i) & (ii) are true

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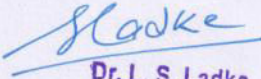
B.Sc.-II (sem-III)

UNIT TEST RESULT

Paper-I(Real Analysis)

Sub:-MATHEMATICS

Sr.no	Name of Students	Marks(20)
1	Hemant M. Farkade	14
2	Tushar D. Asutkar	13
3	Sidhanti Y. Sonewane	13
4	Pooja D. Modot	15
5	Samir A. Nasnurkar	13
6	Nandak R. Kamde	12
7	Pranjal M. Bobade	15
8	Suraj G. Bobade	16
9	Prem S. Ramteke	15
10	Amrita K. Prasad	16
11	Priya U. Shah	12
12	Gayatri B. Donge	14
13	Shruti B. Raipure	13
14	Shreya S. Junarkar	16
15	Prachi v. Patil	15
16	Ritul S. Sharma	12
17	Madhuri P. Kandale	14
18	Sejal S. Petkar	16
19	Kajal A. Varbhe	15
20	sejal S. Deogade	15
21	Shruti M. Ramteke	15
22	Bhavana W. Thakare	16
23	Chirag M. Patil	14
24	shruti A. Deogade	12
25	Dipti V. Bhajankar	13
26	Aman S. Patil	16
27	Payal S. Bavane	15
28	Ritu D. Ingade	14
29	Sapna M. Mandal	13
30	Kiran D. Satpute	14
31	Aniket P. Walke	15
32	Amisha S. Kakde	12
33	Dhanashree k nagraade	14
34	Nikita v kurai	16
35	Vaishnvi M Dhambare	14
36	dhanshree r chende	13
37	snehal s sukhdeve	12
38	vaishnvi p dhage	14
39	Kiran D. Ghosare	13
40	Smita N. Wadhai	14
41	Neha S. Manthanwar	15
42	Mrunali M. Donge	12
43	Sejal R. Nandurkar	13
44	Achal D. Doye	12


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45	Prajakta M. Gedam	15
46	Prashant B. Narwade	14
47	Santosh R. More	12
48	Sushma P. Kokude	13

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Bhadrawati (2020-21)

Unit Test

BSc-II (Sem-IV)

Sub: Mathematics (Paper-I)

Time 30 min

mark-20

Que 1 :- solve any two

(A) Group definitions

(B) Inverse element definitions

(C) Identity element definitions

Que 2 :- show that the binary operation,
defined on \mathbb{R} by $a \cdot b = a + b + ab$ is
commutative and associative

Que 3 :- If G is a group such that
 $(ab)^2 = a^2 b^2$, $\forall a, b \in G$ show that
 G must be abelian

Que 4 :- A group G is finite if $o(G)$ is

(a) Infinite (b) Finite (c) Commutative

(d) None of these

Que 5 :- The identity of a group G is

(a) Unique

(b) Not Unique

(c) Inverse

(d) None of these

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Que 6 :- If a nonempty subset H of the group G is a subgroup of G if and only if.

(a) $a, b \in G \Rightarrow ab^{-1} \in G$

(b) $a, b \in H \Rightarrow ab^{-1} \in H$

(c) $a, b \in H \Rightarrow a^{-1}b^{-1} \in H$

(d) $a, b \in G \Rightarrow a^{-1}b^{-1} \in G$

Que 7 :- In a group G , $a^2 = a, \forall a \in G$ then

(a) $o(G) = 1$

(b) $o(G) = 2$

(c) $a^{-1} \neq e$

(d) G is non-abelian

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B.Sc.-II (sem-IV)

UNIT TEST *Result*

Paper-I(Algebra)

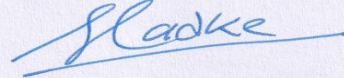
Sub:-MATHEMATICS

Sr.no	Name of Students	Marks(20)
1	Achal D. Doye	14
2	Amisha S. Kakde	13
3	Amrita K. Prasad	14
4	Aniket P. Walke	15
5	Bhavana W. Thakare	13
6	Chirag M. Patil	12
7	Dhanashree k Nagraade	15
8	Dhanshree R Chende	16
9	Gayatri B. Donge	15
10	Hemant M. Farkade	16
11	Kajal A. Warbhe	12
12	Kiran D. Ghosare	14
13	Kiran D. Satpute	13
14	Madhuri P. Khandale	16
15	Nandak R. Kamle	15
16	Nikita v Kurai	12
17	Pooja D. Modot	14
18	Prachi v. Patil	16
19	Prajakta M. Gedam	15
20	Pranjal M. Bobade	15
21	Prashant B. Narwade	15
22	Prem S. Ramteke	16
23	Priya U. Shah	14
24	Ritul S. Sharma	12
25	Samir A. Nasnurkar	13
26	Santosh R. More	16
27	Sapna M. Mandal	15
28	Sejal R. Nandurkar	14
29	sejal S. Deogade	13
30	Sejal S. Petkar	14
31	Shreya S. Junarkar	15
32	shruti A. Deogade	12
33	Shruti B. Raipure	15
34	Shruti M. Ramteke	16
35	Sidhanti Y. Sonewane	14
36	Smita N. Wadhai	13
37	Snehal S. Sukhadeve	12
38	Suraj G. Bobade	14
39	Sushma P. Kokude	13
40	Tushar D. Asutkar	14
41	Vaishnvi P Dhage	15

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42	Runail S Amate	12
43	Palash D. Ghate	13
44	Pratiksha D Parchake	13
45	Payal S. Bavane	15
46	Ritu D. Ingade	14
47	Vaishnvi M Dhambare	13
48	Aman S. Patil	15



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Nilkanth Rao Shinde SA and Art's college
Bhadrawati (2020-21)

Unit Test -

BSc - III (sem-V)

Sub: Mathematics (Paper-1)

mark-20

Time 30 min

Que 1:- In \mathbb{C}^2 define, for $u = (\alpha_1, \alpha_2)$ and $v = (\beta_1, \beta_2)$, $(u, v) = 2\alpha_1\bar{\beta}_1 + \alpha_1\bar{\beta}_2 + \alpha_2\bar{\beta}_1 + \alpha_2\bar{\beta}_2$
show that this defines an Inner product
on \mathbb{C}^2

Que 2:- Let V be a set of all continuous
complex-valued functions on the closed
interval $[0, 1]$, If $f(t), g(t) \in V$, define

$$(f(t), g(t)) = \int_0^1 f(t) \cdot \bar{g}(t) dt$$

Show that this defines an Inner
product on V .

Que 3:- Let V be an Inner product space over
 F . In V define the distance $d(u, v)$ from u
to v by $d(u, v) = \|u - v\|$.

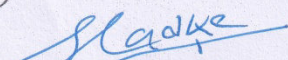
prove that

① $d(u, v) \geq 0$ and $d(u, v) = 0 \Leftrightarrow u = v$

② $d(u, v) = d(v, u)$

Que 4:- Every orthogonal set is

- (a) linearly dependent
- (b) linearly independent
- (c) linear space
- (d) none of these


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Que: ⑤ select the correct answer

- ① $W \cap W^\perp = I$
- ② $W \cap W^\perp = \{0\}$
- ③ $W^\perp \cap W^\perp = \{0\}$
- ④ $W \cap W^\perp = \{0, I\}$

Que 6:- Let V be an inner product space.

then $u, v \in V$ and $\alpha \in F$

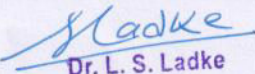
- ① $\|\alpha u\| = |\alpha| \|u\|$
- ② $\|u+v\| \leq \|u\| + \|v\|$
- ③ $\|u\| = 0 \Leftrightarrow u = 0$
- ④ None of these

Que 7: If u and v be vectors in an inner product space such that $\|u+v\| = 8$, $\|u-v\| = 6$ and $\|u\| = 7$ then the value of $\|v\|$ is

- ① $\sqrt{5}$
- ② 1
- ③ $\sqrt{2}$
- ④ 2

Nilkanthrao Shinde Science and Arts College, Bhadrawati dist. Chandrapur
B.Sc.-III (sem-V)
UNIT TEST Result
Paper-I(Linear Algebra)
Sub:-MATHEMATICS

Sr.no	Name of Students	Marks(20)
1	Tejaswini M. Pise	15
2	Priyanka N. Lingayat	14
3	Hemlata B. Gathe	16
4	Kanchan K. Nanne	15
5	Jayesh J. Bodhe	15
6	Saurabh S. Vidhate	16
7	Shivam S. Saxena	16
8	Prajwal V.Pandhare	14
9	Pallavi V. Khamankar	15
10	Gayatri K. Bonde	16
11	Somya C. Singh	19
12	Rahul A. Shah	14
13	Gandhi M. Pothagani	13
14	Sakshi D. Masirkar	16
15	Priyanka S. Tarale	15
16	Atul L. Mohurle	15
17	Jai S. Pandhare	14
18	Ankita K. Urkude	16
19	Pornima S. There	15
20	Shivani Y. Jogi	16
21	Neha P. Ambekar	18
22	Prajakta R. Hatzade	17
23	Akshay M. Bhusari	14
24	Sakshi S. Awari	14
25	Pritija A. Kodape	13
26	Samiksha S. Pawar	16
27	Karishma T. Ghugal	15
28	Priya R. Ghorude	14
29	Pooja A. Pal	13
30	Sahil P. Deogadde	14
31	Kunal P. Shindew	15
32	Sarika S. Nagpure	16
33	Tanmay A. Kolhe	16
34	Prashant S. Gupta	16
35	Nikhil B. Atram	14
36	Subodh S. Raut	13
37	Jayshree R. Durve	15
38	Nishas Mendhe	14
39	Ankita R. Jawade	16
40	Ashitosh K. Prasad	14
41	Pritam S. Harshe	15
42	Gauri I. Sharma	15
43	Akshay A. Aaut	13
44	Harshal B. Khangar	12


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45	Deepak V. Wailkar	15
46	Jitesh J. Mujumdar	16
47	Karan R. Durge	14
48	Chetan S. Datey	16
49	Bhagyashree Y. Parkhi	15
50	Pranali R. Thak	16
51	Shivani S. Ramteke	14
52	Shubhangi N. Dhavale	16
53	Pranay B. Zalwade	15
54	Bhushan G. Bavane	14
55	Pankaj P. Khilare	17
56	Gurudeo P. Kumbhare	16
57	Vishwjit V. Tonge	15
58	Shobit P. Bura	14
59	Simran S. Sheikh	16
60	Chitalee N. Darvankar	15
61	Yogita S. Bahure	16
62	Amit D. Sapat	15
63	Pallavi R. Deharkar	14
64	Sanisagar A. Nikhade	16
65	Suraj B. Mankar	16
66	Esha V. Choudhary	14
67	Vaishali M. Khade	15
68	Pallavi M. Adate	16
69	Ranjana P. Paswan	15
70	Vanshree S. Kamble	15
71	Komal C. Ranvir	14
72	Pronoti C. Salame	15
73	Priti G. Punvatkar	16
74	Shital T. Shiwarkar	14
75	Samiksha B. Dongare	16
76	Ganesh U. Masharkar	13
77	Durgeswari P. Mallelwar	15
78	Aman R. Rajbhar	14
79	Pallavi V. Vidhate	16

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Milkanth Rao Shinde Sci. and Arts College
Bhadrawati (2020-21)

Unit Test

BSc-III (Sem-VI)

Sub:- Mathematics (Paper-I)

Time 30 min

mark-20

Que 1:- The derivative of any constant vector function is

(a) one

(b) n

(c) zero

(d) None

→ (2)

Que 2:- If f and g are two differentiable vector functions of t then

(a) $\frac{d}{dt}(f+g) = \frac{df}{dt}g + f \cdot \frac{dg}{dt}$

(b) $\frac{d}{dt}(f+g) = \frac{df}{dt}g - f \frac{dg}{dt}$

(c) $\frac{d}{dt}(f+g) = \frac{df}{dt} + \frac{dg}{dt}$

(d) $\frac{d}{dt}(f+g) = \frac{df}{dt} \cdot \frac{dg}{dt} + \frac{dg}{dt} \cdot \frac{df}{dt}$ → (2)

Que 3:- If $\oint_S \vec{f} \cdot \vec{n} ds = 0$, \vec{f} is called

(a) Irrotational (b) rotational

(c) solenoidal (d) None of these

→ (2)

Que 4:- A vector field f is conservative

(a) $\nabla \cdot f = 1$

(b) $\nabla \cdot f = 0$

(c) $\nabla \times f = 0$

(d) $\nabla \times f = 1 \rightarrow$ (2)

Que 5:- If $v = \omega \times z$. Prove that

$\omega = \frac{1}{2} \text{curl } v$, where ω is a constant vector \rightarrow (4)

Que 6:- If $f = t^3 i + \frac{1}{2} t^2 j + \frac{1}{3} t k$

find the value of

$\int f \cdot dz$ from $(0,0,0)$ to $(2,4,8)$

along the path $C: x=t, y=t^2, z=t^3 \rightarrow$ (4)

Que 7:- Evaluate the surface integral

$\iint_S (yz i + zx j + xy k) \cdot ds,$

where S is the surface of the sphere

$x^2 + y^2 + z^2 = a^2$ in the first octant. \rightarrow (4)

Nilkanthrao Shinde Science and Arts College, Bhadrawati dist. Chandrapur
B.Sc.-III (sem-VI)

UNIT TEST *Result*

Paper-I(Complex Analysis and Vector Calculus)

Sub:-MATHEMATICS

Sr.no	Name of Students	Marks(20)
1	Akshay A. Raut	14
2	Akshay M. Bhusari	13
3	Aman R. Rajbhar	15
4	Amit D. Sapat	14
5	Ankita K. Urkude	13
6	Ankita R. Jawade	14
7	Ashitosh K. Prasad	15
8	Atul L. Mohurle	16
9	Bhagyashree Y. Parkhi	15
10	Bhushan G. Bavane	16
11	Chetan S. Datey	12
12	Chitalee N. Darvankar	14
13	Dipak V. Wailkar	13
14	Durgeswari P. Mallelwar	16
15	Isha V. Choudhary	15
16	Gandhi M. Pothagani	12
17	Ganesh U. Masharkar	14
18	Gauri I. Sharma	16
19	Gayatri K. Bonde	15
20	Gurudeo P. Kumbhare	16
21	Harshal B. Khangar	15
22	Hemlata B. Gathe	16
23	Jai S. Pandhare	14
24	Jayesh J. Bodhe	12
25	Jayshree R. Durve	13
26	Jitesh J. Mujumdar	18
27	Karan R. Durge	15
28	Karishma T. Ghugal	14
29	Komal C. Ranvir	13
30	Kunal P. Shinde	14
31	Neha P. Ambekar	19
32	Nikhil B. Atram	12
33	Nishas Mendhe	18
34	Pallavi M. Adate	16
35	Pallavi R. Deharkar	14
36	Pallavi V. Khamankar	13
37	Pankaj P. Khilare	16
38	Pooja A. Pal	14
39	Pornima S. There	13
40	Prajakta R. Hatzade	14
41	Prajwal V. Pandhare	15

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42	Pranali R. Thak	12
43	Pranay B. Zalwade	13
44	Prashant S. Gupta	12
45	Pritam S. Harshe	14
46	Priti G. Punwatkar	15
47	Pritija A. Kodape	12
48	Priya R. Ghorude	13
49	Priyanka S. Tarale	15
50	Pronoti C. Salame	13
51	Rahul A. Shah	12
52	Ranjana P. Paswan	14
53	Sakshi D. Masirkar	16
54	Sakshi S. Awari	15
55	Samiksha S. Pawar	14
56	Shanisagar A. Nikhade	13
57	Sarika S. Nagpure	15
58	Saurabh S. Vidhate	14
59	Shital T. Shiwarkar	16
60	Shivam S. Saxena	16
61	Shivani S. Ramteke	14
62	Shivani Y. Jogi	14
63	Shobit P. Bura	15
64	Shubhangi N. Dhavale	16
65	Somya C. Singh	20
66	Subodh S. Raut	14
67	Suraj B. Mankar	13
68	Tanmay A. Kolhe	14
69	Vaishali M. Khade	15
70	Vishwjit V. Tonge	16
71	Yogita S. Bahure	15

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B.Sc sem - II 28/10/2020

Unit test III Marks 15

Sub - Botany Question paper

- Q1 What is auxin? Give its role - 5
- Q2 What is Gibberellic Acid? Explain it - 5
- Q3 What is abscisic acid? Explain it - 5

Name of students	Obtained mark	Total marks
(i) Chuli p. Madke	12	15
(ii) Kamal Bhathe	09	15
(iii) Neha Prasad	15	15
(iv) Ritesh Tiwari	15	15
(v) Karishma Ghandarkar	15	15
(vi) Harshad Asker	15	15
(vii) Nutan Korde	14	15
(viii) Mayur V. Gawarekar	10	15

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PLM 15 BSC Sem V
Unit test - II 29/10/2020

Question paper Sub - Botany

- Q1 Describe the sex determination in *Scoropachya*. 5
- Q2 Describe the sex-determination in plant (*Melastomium*) 5
- Q3 What is barr bodies 2 1/2
- Q4 What is Lyon's hypothesis 2 1/2

Name of students	obtained marks	Total marks
1) Nishi S. Kherkar	08	15
2) Anuraj N. Shaikh	09	15
3) Manika Korde	07	15
4) Nidhi A. Balki	10	15

L. S. Ladke
29/10/2020
Asst. Prof in Botany
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Total
marks - 15

BSC. SEM I
Unit test-I on unit II 22/10/2020
Sub: Botany - Paper 2
Question paper

- Q1 Describe asexual reproduction in *Noctua* - 5
- Q2 Give the economic importance of Algae - 2
- Q3 Give the economic importance of Fungi - 2
- Q4 Write about classification of plants - 2
- Q5 Explain sexual reproduction types in algae - 2

Name of students	Obtained marks	Total marks
(1) Dolly Singh	08	15
(2) Nilima Yedwar	10	15
(3) Gudiya Singh	07	15
(4) Tanvi Patil	12	15
(5) Samiksha Swan	11	15
(6) Sakshi Mathe	13	15

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23/10/2020
Asst prof in botany
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09/12/2020 Unit Test - Time - 3:50 - 10:30 am B.Sc. Sem IV
UPIC - I P-II

Q-1 :- Write on — 10 Marks

- a] tRNA
- b] mRNA

Q-2 :- Write short notes on — 10 Marks
 $\frac{2}{2} \times 4$

- a] srRNA
- b] Antisense RNA
- c] miRNA
- d] rRNA

Q-3 :- Write answers in one or two
lines only. — 5 Marks each

- a] siRNA
- b] Types of RNA
- c] RNA
- d] snRNA
- e] Function of mRNA

ARNNNSR
09/12/2020

Participants —

1. Likita Suresh
2. Prayanka Samal
3. Shubhangi Kshirsagar
4. Rashmi Patel
5. Aswini Shetty
6. Praveen Kumar Nasa

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ARNNNSR

Nikantaraso Shinde Science and Arts College, Bhadrawati

B.Sc. Sem-V; Time - 10:40 - 11:20 pm

Paper - II
Test, Unit II date - 01/21/2021 online exam

Topic - Karyotype :-

Q-1 :- Write on - 10 marks

- Aneuploidy.
- Inversion

Q-2 :- Write short notes on - 10 marks - $2\frac{1}{2}$ marks each

- Deletion or deficiency
- Duplication
- Translocation complex
- Nullisomics

Q-3 :- Write answers in two or three lines only. 5 Marks 1 Mark each

- Karyotype
- Haploid
- Diploid
- Triploid
- Tetraploid

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01/21/2021

XXXXXXXXXX

Nikonthao Shinde Science and Arts College,
Bhadrawati, Dist- Chandrapur, Maharashtra
Sem. III
Unit I
P-II Unit I
Time- 9:50 to 10:30 am.

Q-1 :- Write on - 10 Marks

- Physical and chemical properties of carbohydrates.
- Structure of Glucose

Q-2 :- Write short notes on - 10 Marks

- Structure of Sucrose
- Cholesterol
- Sphingolipids
- Physical properties of lipids

Q-3 :- Write answers in one or two lines only. - 5 Marks

- carbohydrates
- protein
- lipid
- sterols
- Lecithine

~~NRPNASAR~~
08/12/2020

~~NRPNASAR~~

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Nilkanthrao Shinde, Science and
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Test P-II Unit-II Sem-III Date: 12/11/2020

Time - 10-40 to 11-20 am

Q-1:- Write on - 10 Marks

- IUB system of Nomenclature of Enzymes.
- Characteristics of the Enzyme.

Q-2:- Write short notes on - 10 Marks

- Holoenzyme
- Apoenzyme
- Co-enzyme
- Co-factor

Q-3:- Write answers in two or three lines only. - 5 Marks

- Enzymology
- Enzyme
- Enzyme-substrate complex
- Lock & Key Model of Enzyme
- Any two properties of Enzyme

MRPNaloo

12/11/2020

Participants -

- Pravankumar Naware
- Nutan Kardo
- Mayur, Vishal, Ganesh
- Neha Poojari
- Komal Badhe
- Ritesh Tiwari

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MRPNaloo

Nitkeshrao sheide science and Arts college;
Bhadrawati; Dist-Chandrapur.

Test - 07/01/2021 P-II Time - 9.00am-9.45
Unit-III am

Topic - Ascent of Sap

Que-1 :- Write on — 10 marks

- a] Ascent of sap
- b] Root Pressure Theory.

Que-2 :- Write short notes on — 10 Marks

- a] water conduction through xylem; 2½ marks each
- b] cohesion Adhesion - Theory.
- c] Salicaceae plant
- d] Galotropis process a plant - as an example for Ascent of sap.

Que-3 :- Write answers in two or three 5 Marks
lines only.

- a] Transportation.
- b] xylem
- c] Ascent of sap
- d] Root pressure
- e] phloem

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Bhadrawati, Dist - Chandrapur - 442 902

Visit Test

Max Marks - 25

Date - 27/01/2021; class - B.Sc. Sem - III

Time - 9-50 - to 10-30 am P-II, Unit - IV

Topic - Photosynthesis

Question Paper -

Q-1 :- Write on - 10 marks

- Photosynthetic pigments
- Action spectra, Red drop and Emission enhancement effects

Q-2 :- Write short notes on - 10 marks
(2½ each)

- Cyclic photophosphorylation
- Non-cyclic photophosphorylation
- C₃ pathway
- C₄ pathway

Q-3 :- Write answers in two or three lines only. 5 marks (1 mark each)

- Photosynthesis
- Examples of C₃ plants
- Examples of C₄ plants
- Examples of CAM plants
- Contribution of Emerson.

~~PR~~
27/01/2021

Participants -

- Ritesh Divali
- Neha Prasad
- Dr. Praveenkumar N. Narsare.

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Nilkanteshwar Shinde Science and Arts College, Bhadrawati
Dist-Chandrapur, Maharashtra State

B.Sc. Part - I, Sem-II, BOTANY online exam

Paper-II; Unit-I; Unit Test-I

Question Paper -

Max. Marks - 25

Que-1; - Write on -

10 Marks

a] Bennettian Theory - 5 Marks

b] Taxonomic ranks - 5 Marks

Que-2; - Write short notes on - 10 Marks

a] Saharianthus flower $2\frac{1}{2}$ each

b] Theophrastus

c] Magnolia

d] Lectotype

Que-3; - Write answers in two or three lines
only

- 5 Marks
01 Mark each

a] Linnaeus

b] Eryngocarpon

c] Isotype

d] Binomial Nomenclature

e] Neotype

ARNNASASE

Dr. P.N. Nasale

Ladke

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Kantheao Shiinde Science and Arts College, Bhadravati
Chandrapur, Maharashtra State.

Se. Part - I, Sem-II, BOTANY online exam

Part-II; Unit-I; Unit Test-I

Question Paper -

Max. Marks - 25

Que-1 :- Write on -

10 Marks

- a] Bennettian Theory - 5 Marks
- b] Taxonomic Ranks - 5 Marks

Que-2 :- Write short notes on - 10 Marks

- a] Saharianthus flower $2\frac{1}{2}$ each
- b] Theophrastus
- c] Magnolia
- d] Lectotype

Que-3 :- Write answers in two or three lines
only

- 5 Marks
01 Mark each

- a] Linnaeus
- b] Eryngium
- c] Isotype
- d] Binomial Nomenclature
- e] Neotype

(MRPNASARE)
Dr. P.N. Nasare

Vilkantshree Shri Sai Science and Arts College,
Shadrawati, Dist - Chandrapur.

B. Sc. Part-I, Sem-II BOTANY Online Exam

Paper-II; Unit-II; Unit Test-II

Question Paper - Max. Marks-25

Que-1:- Write on- 10 Marks

1) Virtual (Digital) Herbarium - 5M

2) Types of classification of Angiosperms - 5M

Que-2:- Write short notes on- 10 Marks
2/5 Marks each

a) Merits of Bentham & Hooker System

b) Demerits of Bentham and Hooker System

c) Classical Herbarium

d) Disfigurement of Bentham and Hooker System.

Que-3:- Write answers in two or three lines only

a) Artificial classification of Angiosperms - 5 Marks
or 1 mark each

b) Herbarium

c) Genera Plantarum

d) Gymnospermae

e) Angiospermae

ARANNALARE

(Dr. P. N. Nalare)

Shacke
Dr. L. S. Ladka
PRINCIPAL
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Shadrawati, Dist-Chandrapur

ARANNALARE

Nilkanteshwar Shinde Science and Arts College;
Bhadrawati; Dist - Chandrapur.

B.Sc. Part-I, Sem-II BOTANY Online Exam

Paper-II; Unit-III, Unit Test-III

Question Paper

Max. Marks-25

Que-1; - Write on - 10 Marks

- a] Reproductive characters of Brassicaceae
- b] Morphological characters of Malvaceae

Que-2; - Write short notes on - 10 Marks

- a] Morphological characters of Brassicaceae
- b] Reproductive characters of Malvaceae
- c] Reproductive characters of Fabaceae
- d] Morphological characters of Solanaceae

Que-3; - Write answers in two or three lines
only. 05 Marks

- a] Cruciform corolla.
- b] Vexillary aestivation
- c] Capsule
- d] Monadelphous condition
- e] 02 Members of family Solanaceae

Ladke

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MRPN Ladke

Nilkarthrao Shinde Science and Arts College;
Bhadrawati; Dist - Chandrapur.

B.Sc. Part-I, Sem-II BOTANY Online Exam

Paper-II; Unit-IV; Unit-Test-IV

Question Paper - Max. Marks-25

Que-1 :- Write on :- 10 Marks

- Reproductive characters of Asclepiadaceae
- Morphology of Asteraceae
- ~~Morpho~~

Que-2 :- Write short notes on - 10 Marks

- Ray floret
- Disc floret
- Morphological characters of ^{Liliaceae} ~~Ray floret~~
- Reproductive characters of ~~Asclepiadaceae~~ ^{Poaceae}

Que-3 :- Write answers in two or three lines
only. 5 Marks
01 Mark each.

- ~~Coenoc~~ Gynostegium
- Pappus
- Palea
- Inflorescence of Liliaceae
- Economic importance of Poaceae NRDNNR

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NRDNNR

Nilkantharao Shinde Science and Arts College, Bhadrawati

Unit Test

Session- 2020-2021

Class- B.Sc-II Sem-III

Subject- Inorganic Chemistry (Unit-I)

Time-45 Minutes

Mark-12

Q-1 what are polyhalide? Give the classification of polyhalides. (5Mark)

Or

What is Carbides? Discuss classification of carbides with suitable example.

Q-2 (Any Two)

A) Explain Chain silicates with examples (2.5 marks)

B) Explain basic nature of Iodine. (2.5 mark)

C) Describe the structure of Borazine. (2.5 mark)

Q-3. Any Two

1. What are interhalogen compounds? (1 mark)

2. Draw the structure of borazine. (1 mark)

3. What are Silicates? (1 mark)

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Bhadrawati, Dist-Chandrapur

(Ladke)
(Pr. S. Ladke)

Nilkantharao Shinde Science and Arts College, Bhadrawati

Unit Test

Session- 2020-2021

Class- B.Sc-I Sem-I

Subject- Inorganic Chemistry (Unit-I)

Time-45 Minutes

Mark-12

Q-1. Define Ionization potential. Describe the factors affecting on it. Explain the Trend of Ionization potential in group and period. (5Mark)

Or

Define Quantum number. Discuss the orbital quantum number and magnetic quantum number in detail

Q-2 (Any Two)

A) Define electronegativity. Why is electron affinity of fluorine less than chlorine? (2.5 mark)

B) State and explain Hund's rule of maximum multiplicity. (2.5 mark)

C) Calculate the effective nuclear charge for 3d electron in Chromium. (2.5 mark)

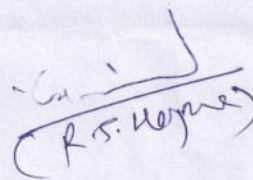
Q-3. Any Two

1. Define Carrying Constant. (1 mark)

2. Al is good reducing agent Explain. (1 mark)

3. Write Schrodinger equation. (1 mark)


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(R. S. Nerurkar)

Nilkantharao Shinde Science and Arts College, Bhadrawati

Unit Test

Session- 2020-2021

Class- B.Sc-III Sem-IV

Subject- Inorganic Chemistry (Unit-I)

Time-45 Minutes

Mark-12

Q-1 What is error? Explain determinate and indeterminate error in detail.
(5Mark)

Or

Explain principle and various types of interferences in Flame Photometry.

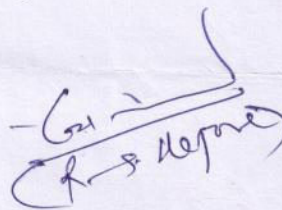
Q-2 (Any Two)

- A) Explain i) Significant figure ii) Accuracy (2.5 marks)
- B) Explain F-Test and T-Test. (2.5 mark)
- C) Distinguished between Accuracy and Precision. (2.5 mark)

Q-3. Any Two

- 1. Define i) Mean ii) Median (1 mark)
- 2. Calculate Significant figure i) 6.0213 ii) 0.00215 (1 mark)
- 3. Write two advantages of Flame Photometry (1 mark)


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Nilkanthrao Shinde Science and Arts College, Bhadrawati
Dist. Chandrapur

Department of Chemistry

Unit Test

Topic:- NMR Spectroscopy

M 20

a. An organic compound having molecular formula C_3H_6O shows following NMR data

i) 3H – Triplet δ -1.5

ii) 2H- Quartet δ -2.6

iii) 1H- Singlet δ -7.2

Deduce the structure.

5

b. Write brief note on-

i) The Role of TMS in NMR Spectroscopy.

ii) Equivalent and Non-equivalent protons.

5

c. What is chemical Shift ?

$2\frac{1}{2}$

d. Define coupling constant 'J'.

$2\frac{1}{2}$

e. What is shielding and deshielding of protons in NMR spectroscopy?

$2\frac{1}{2}$

f. What is spin – spin coupling.

$2\frac{1}{2}$

Sachin H. Shirsane
Lt. Sachin H. Shirsane

Ladke
Dr. L. S. Ladke
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N.S. Science & Arts College
Bhadrawati, Dist-Chandrapur

Nilkanthrao Shinde Science and Arts College, Bhadrawati
Dist. Chandrapur

Department of Chemistry

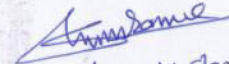
Unit Test

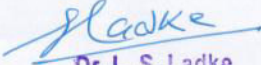
Topic:- Organic Synthesis Via Enolate

M 20

Date 10/11/2020

-
- Q.1 What are enolates? Give two examples with their structures. 5 M
- Q.2 How is acetoacetic ester prepared by Claisen condensation ? Starting from acetoacetic ester, how will you prepare the following?
- (I) 4-methyl uracil & (II) Ethyl methyl ketone 5M
- Q.3 Give short account of Keto-Enol Tautomerism. 2.5 M
- Q.4 How will you obtain following compounds from diethyl malonate
- (I) n-valeric acid & (II) adipic acid 2.5 M
- Q.5 How is diethyl malonate prepared from acetic acid 2.5 M
- Q.6 Write nte on Acidity of α -hydrogen atom 2.5M

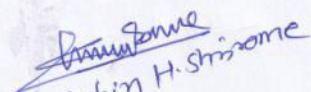

Lt. Sachin H. Shirname

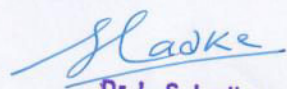

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Nilkanthrao Shinde Science and Arts College, Bhadrawati
Dist. Chandrapur
Department of Chemistry
Unit Test
Topic:- Green Chemistry

M 20

-
- | | |
|---|-------------------------------|
| a. Define Green Chemistry. Discuss the main purpose of Green Chemistry. | 5 |
| b. Explain the alternative solvent or replacement of solvents in green chemistry. | 5 |
| c. What do you mean by cradle to cradle design ? | 2 ¹ / ₂ |
| d. Discuss scientific areas for practical applications of green chemistry. | 2 ¹ / ₂ |
| e. Write short note on Reduction of solvent toxicity. | 2 ¹ / ₂ |
| f. Write a note on Phase transfer catalysis. | 2 ¹ / ₂ |


Lt. Sachin H. Shirsone

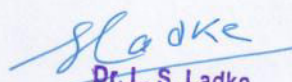

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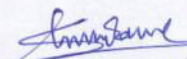
Nilkanthrao Shinde Science & Arts college, Bhadrawati
Department of Chemistry
Unit Test(online)
Topic :- Polymer

Marks 20

Date:- 17/08/2020

-
- Q.1 How do you Classify polymer in various way? (5M)
- Q.2 What are cross-linking process? Give any two example. (5M)
- Q.3 What is phenol formaldehyde resin? Give it's properties and application (2.5M)
- Q.4 What is ruber? Explain neoprene & Buna-S with application. (2.5M)
- Q.5 What is vulcanization process? (2.5M)
- Q.6 What is conductiong polymer? (2.5M)


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Lt. Sachin H. Shrivastava

Nilkanthrao Shinde Science & Arts college, Bhadrawati

Department of Chemistry .

Unit Test(online)

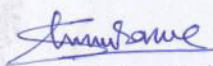
Popic :- Polymer


Marks 20

Date:- 17/08/2020

Test result

Sr.no	Name	Mark obtained	Out of	date	remark
1	Aramnaz Sheik	17	20	17/08/2020	
2	Atul Mohurle	14	20	17/08/2020	
3	diksha deogade	15	20	17/08/2020	
4	Durgeshwari Mallelwar	13	20	17/08/2020	
5	gayatri Bonde	18	20	17/08/2020	
6	Mrunal Dadmal	17	20	17/08/2020	
7	Neha Ambekar	19	20	17/08/2020	
8	Nidhi Balki	13	20	17/08/2020	
9	nikita dhengale	14	20	17/08/2020	
10	pooja vidhate	13	20	17/08/2020	
11	Priya ghorude	12	20	17/08/2020	
12	Saujanya Chukkwar	16	20	17/08/2020	
13	shivani patarange	15	20	17/08/2020	
14	somya singh	19	20	17/08/2020	
15	Swati Giradkar	15	20	17/08/2020	


Lt. Sachin H. Shrivastava


Dr. L. S. Ladke
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Bhadrawati, Dist-Chandrapur

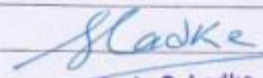
16	Vaishvi dulhare	16	20	17/08/2020	
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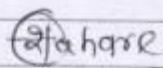
Sachin
H. Sachin H. Shrivastava

Ladke
Dr. L. S. Ladke
PRINCIPAL
N.S. Science & Arts College
Bhatrawati, Dist-Chandrapur

UNIT TEST Result.
B.A III - sem V
Sub - History.

Sr.	Name of students	Mark out of 20
①	Bobade Pranali	12
②	Bodhe Kiran	13
③	Dehaskar Utkarsha.	14
④	Dongare Pratiksha.	13
⑤	Giripurje Divya.	12
⑥	Kakade Bharati	09
⑦	Matte Mayuri	08.
⑧	Nibude papita	16
⑨	Klabh Rashmi	08.
⑩	Kalawati Khargas,	09
⑪	Pimpalkar Vaishnavi	08.


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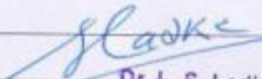

Prof. Kundan D. Shahare
Department of History
N.S. College Bhadravati

UNIT Test Result.

B.A II year. sem IIIrd.

Sub: => History.

Sr.		mark out of 20.
①	Amane T. Pranalika.	14
②	Bhoyar M. Prajakta.	13
③	Dhabekar D. Karishama.	12
④	Ekade Mahesh	10
⑤	Go Gaikwad. Amol.	16
⑥	Kakde C. Damini	12
⑦	Khamankar I. Wachespati	12
⑧	Kosare A. Sweta.	13
⑨	Kutlermare Varsha B.	08
⑩	Milmile Karishama. B.	07


Dr. L. S. Ladke
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Bhadrawati, Dist-Chandrapur

शिंदे
प्रा. कुंदन दे. शारंग
इतिहास विभाग मुख्यालय
नि. शि. वि. क. महा. भद्रावती

UNIT Test = Sem V

B.A III = History

माकर्स => 20

वेळ: 9 लास

प्र१ला

अ) स्टॅम्प कायद्या विषयी आपणास काय माहिती आहे.

ब) अमेरिकन स्वतंत्र्य युद्धात 'लोयन टी पार्टी' वटना स्पष्ट करा.

क) फ्रेंच क्रांतीत लेखी कॅम्ब्रिज रूसी ये कार्य स्पष्ट करा.

द) माटेस्क्यू चा सत्ता विभाजनाचा सिद्धांत स्पष्ट करा.

Ladke

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निहतेराय सिदी विशाल व कुल महाविद्यालय भद्रावती.
एच. केंद्र.

Unit. Test.

Class - B.A. III Sem. VI

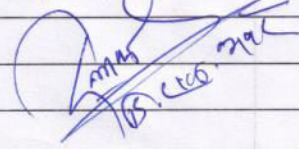
Subject - Sociology

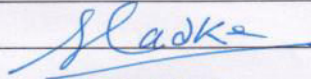
Marks - 20

प्र.1 राजकीय संघटनेचा अर्थ लिहा.

प्र.2 जातिवादी संघटना समाजातील जायदपाले कायप लिहा

प्र.3 जातिवादी समाजातील व्यायमी संघटना लिहा.


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Bhadrawati, Dist-Chandrapur

निहलुडरुडु शुरु वलशुनु व अलु नुडलुडुडुडु, शुरुडुडु
डु. डुडुडु

Unit Test.

class - B.M.II Sem. IV

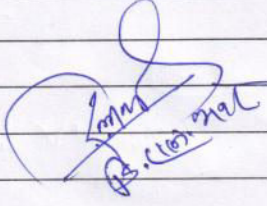
Subject - Sociology

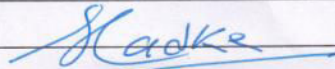
marks - 20

डु. 1. शुरुडुडुडुडु डुडुडुडुडु डुडुडुडु.

डु. 2. शुरुडुडुडुडु डुडु डुडुडु डुडु.

डु. 3. लुडुडुडुडु डुडुडु डुडुडुडु.


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Nilkanthrao Shinde Science and Arts College, Bhadrawati
Dist. Chandrapur.

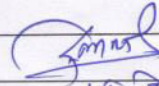
B.A. II, Sem. IV

Unit Test Result - Sociology

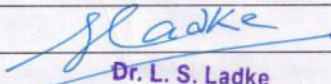
Total
Marks 20

Name of the Student

1	Ku. Karishma B. Milmile	19
2	Ku. Gita m. Thawari	18
3	Ku. Pallavi m. Balpane	18
4	Ku. Damini C. Kakde	17
5	Ku. Shrutika S. Detha	17
6	Ku. Priyanka D. Detha	16
7	Ku. Akshata S. Umare	19
8	Ku. Pranalika T. Amane	17
9	Ku. Vaishali R. Thakore	19
10	Amol U. Gaikwad	19



Dr. L. S. Ladke



Dr. L. S. Ladke
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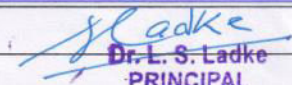
N.S. Science & Arts College
Bhadrawati, Dist. Chandrapur

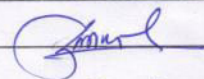
Nilkanthrao Shinde Science and Arts College,
Bhadrawati. Dist. Chandrapur.

Unit Test Result
B.A. III Sem. VI
Subject. - Sociology

Total marks - 20

	Name of the Student	
1	Ku. Prati Suryawansi	18
2	Ku. Pranali D. Bobade	17
3	Raishagri G. Tarale	18
4	Sonali D. Mondhe	16
5	Mahesh H. Kakde	19
6	Nishal B. Jivane	15
7	Mangala S. Tajme	16
8	Sanket B. Lode	18
9	Sanita D. Asekarr	17
10	Yogeshwar M. Shende	17


Dr. L. S. Ladke
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Dr. Pravin D. Patil

**NILKANTHRAO SHINDE SCIENCE AND
ARTS COLLEGE, CENTRE FOR HIGHER
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BHADRAWATI, DIST. CHANDRAPUR**



**DEPARTMENT OF SOCIOLOGY
CERTIFICATE**

This is to certify that Shri / Ku. / Smt. Akshata
Shalik Umrae is a bonafide student of
Class BA And year Roll No. _____ He / She has done
Unsatisfactory / Satisfactory / Good / Excellent Assignment work in
the subject Sociology in Academic Session 2020-2021

Incharge Teacher

Head of the Deptt.

N. S. Sci. & Arts College, Bhadrawati

Principal

N. S. Sci. & Arts College, Bhadrawati

Dr. L. S. Ladke
PRINCIPAL

**N.S. Science & Arts College
Bhadrawati, Dist-Chandrapur**

Anand Screen, Bhadrawati 9850372295

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ARTS COLLEGE, CENTRE FOR HIGHER
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BHADRAWATI, DIST. CHANDRAPUR**



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CERTIFICATE**

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सुर्यवंशी is a bonafide student of

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V Sem
Unsatisfactory / Satisfactory / Good / Excellent Assignment work in
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Incharge Teacher


Head of the Deptt.

Principal

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CERTIFICATE**

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Ramdeke is a bonafide student of
Class A.A. III year Roll No. _____ He / She has done
Unsatisfactory / Satisfactory / Good / Excellent Assignment work in
the subject History in Academic Session 2020-21.

Dr. K. D. Shahare.
Incharge Teacher

Shakur.

Head of the Deptt.

N. S. Sci. & Arts College, Bhadrawati

Principal

N. S. Sci. & Arts College, Bhadrawati

Ladke

Dr. L. S. Ladke
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CERTIFICATE**

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Gairwal is a bonafide student of
Class B.A. II Roll No. _____ He / She has done
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the subject History in Academic Session 2020-21.

Dr. K. D. Shahare
Incharge Teacher

[Signature]
Head of the Deptt.

N. S. Sci. & Arts College, Bhadrawati

Principal

N. S. Sci. & Arts College, Bhadrawati

[Signature]
Dr. L. S. Ladke
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Anand Screen, Bhadrawati 9850372295

**NILKANTHRAO SHINDE SCIENCE AND
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BHADRAWATI, DIST. CHANDRAPUR



**Assignment / Seminar /
Unit Test Book**

ACADEMIC SESSION

2020 - 2021

Name of the Student: Nitu Kumari Dilip Kumar Yadav

Class: Bsc-1st year 2nd Sem Section: _____

Subject: chemistry Roll No. _____

Univ. Enrolment No: _____

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Ladke

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Unit Test Book**

**ACADEMIC SESSION
2020 - 2021**

Name of the Student: Sameer A. Nasurkar
Class: B.Sc-II Sem-A Section: (PCM)
Subject: Chemistry (organic) Roll No. _____
Univ. Enrolment No: _____

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**Assignment / Seminar /
Unit Test Book**

ACADEMIC SESSION

20²⁰ - 20²¹

Name of the Student : Shivam. J. Senare
Class : Bsc - Final (Sem II) Section : _____
Subject : Chemistry Roll No. _____
Univ. Enrolment No : _____

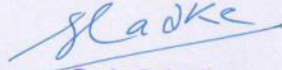
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Nilkanthrao Shinde Science and Arts College, Bhadrawati

Preliminary Examination-2021

Preliminary Exam was not taken during the session-2020-21 due to corona pandemic.



Dr. L. S. Ladke
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