# Nilkanthrao Shinde Science & Arts College, Bhadrawati, Dist. Chandrapur (MS)

# 2.5.1

# Mechanism of internal assessment is transparent and robust

# Unit Test BSC. II SemIII, 2/19/2000

Timestamp	email of the student	What is mixture homogeneous		What is the difference between molality & molarity?	What Is normality?	Define colligative properties.	May
2020/09/21 10:09:42 AM GMT+5:30	Pratikshazade06 @gmail.com	component throughout any given sample. A solution is a	A solution is a homogeneous mixture of two	Molality:- It is define as the number of moles of solute dissolve in 1000g of the solvent:- Molarity:- It is define as the number of moles of solute dissolve imper litre of the solution.	The normality of the solution is define as the number of gram equivalent of the solute	Colligative properties define as those which depend entirely upon the number of a particle of the solute dissolve in a known volume of a given solvent and not a all upon the nature of a solute is called colligative properties.	10
2020/09/21 10:12:43 AM ĜMT+5:30	thakurtushar17 1@gmail.com	The reaction in which the catalyst is present in the same phase as the reactants, is called as homogeneous mixture	A solution is a homogenous mixture of two or more substances having uniform properties.	1) Molality: It is defined as the number of moles of the solute dissolve in 100kg of the solvent . 2) Molarity: It is defined as the number of moles of solute dissolved in per litre of solution .	Normality of the solution is defines as the number of the gram equivalent of solute dissolved per litre of solution.	Colligative property are those who depends entirely upon the no. of particles of solute dissolved in a known volume of a	10
		A mixture which	A homogeneous mixture containing of one phases and			The change depends on the number of the solute particles	
Dr.L. S. Ladi 2015/09/2019AL 2015/09/34 Arts 24/41,50jst-Cha	se	is evenly distributed. e.g salt water is homogeneous mixture i.e. salt gets dissolved in	containing only two components is called as solutions. E.g < solution of NaCl	Molality (M) = moles of solute / kg of solvent. And. Miclarity = moles or solute / litre solute / litre	Normality (N) = No. gram equivalent of solute / litre of	in the solution, not the nature and size of solute particles is called as colligative properties.	0)

Dr. A. B. Sloll

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2020/09/21		Homogeneous mixtures are also referred to as solutions. While we normally think of solutions as liquids, such as soft drinks and lemonade, they can actually be in the form of solids, liquids, and gases. Homogeneous mixtures can also be a combination of these forms, as	a solution is a special type of homogeneous mixture composed of two or more substances. In such a mixture, a solute is a substance dissolved in another substance,	Molarity is the ratio of the moles of a solute to the total liters of a solution. The solution includes both the solute and the solvent. Molality, on the other hand, is the ratio of the moles of a solute to the	The normality of a solution is the gram equivalent weight of a solution For example, the concentration of a hydrochloric acid solution might be expressed as 0.1 N HCl. A gram equivalent weight or equivalent is a measure of the reactive capacity of a given chemical	colligative properties are those properties of solutions that depend on the ratio of the number of solute particles to the number of solvent molecules in a solution, and not on the nature of the chemical species present The word colligative is derived from the Latin colligatus	100
	hand /		1.02501			The state of the s	1
10:25:35 AM GMT+5:30	harshadaskar00 1@gmail.com	in a liquid-gas mixture.	known as a solvent.	kilograms of a	species (ion, molecule, etc.).	meaning bound together.	
2020/09/21 10:26:04 AM GMT+5:30	mayuriwatkar06 2000@gmail.co m	A reaction in which the catalyst is present in the same reaction as a reactant is called as homogeneous mixture.	A solution is a homogeneous mixture of two or more substance on molecular level.	Molarity-no.of moles of solute per litre of a solutions.Molali ty-no.of moles of solute per kilogram of solvent.	It is difine as no.of equivalent of solute per litre of solutions.	Property which depends on particles of solute dissolved in give n solvent and not on chemical composition of solute.	10
Dr. L.S. PRINC V.S. Science & hadrawati, Dis 2020/09/21 10:35:21 AM GMT+5:30	PAL Arts College	A homogeneous mixture is a mixture that has a same properties of its component throughout any given sample.	mixture of two	Molality:- a)It is defined as the number of solute dissol in 100g of the solvent. B) It is doneted by m. Molarity:- a) number of mole of solute dissolved in per litter of the solution b) It is denoted by m.	The number of gram equivalent of the solute dissolved per liter of the solution .	Colligative properties defined as those which depends entirely upon the number of particles of the solute dissolved in a known volume of a given solvent and not a all upon the nature of the solute.	رژ

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2020/09/21 10:13:39 AM	achalkarmaujiya		A solution is a homogenous mixture of two or more substances having uniform	Morality-It is defined as the no. Of moles of the solute dissolve in 1000g of the solvents, and Molarity- It is defined as the no. Of moles of the solute dissolve per litre		Colligative Properties are those which depend entirely upon the no. Of particles of the solute dissolved in a know volume of a given solvent and not at all upon the nature
GMT+5:30	111@gmail.com	catalysis.	properties.	of the solution.	solution.	of the solute.
	Egen		A solution is a homogenous mixture of two	The difference between molality and molarity is , in molality the	The normality of the solution defined as the number of gram	Colligative properties are those which depend entirely upon the number of particles of the solute dissolved in a known volume of a given solvent and not at all upon the nature. The various
2020/09/21 10:17:24 AM GMT+5:30	guddikakde@ gmail.com	Solution is homogenous if its composition is uniform through the body of solution.	or more substance having uniform properties throughout.Solu tions are contain two component solute and	number of moles of the solute dissloved in 100g of the	equivalent of the dissolved per litre of the solution.Normal ity =Number of gram equivalent of the soluteĀ-Volume of the solution	colligative properties are- 1. Osmotic pressure 2.Lowering of vapour pressure 3. Elevation in

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2020/09/21 10:37:47 AM GMT+5:30	Nutankorade02 @gmail.com	A homogeneous mixture is a solid, liquid or gaseous mixture that has the same proportions of its components throughout any given sample. Conversely, a heterogeneous mixture has components in which proportions vary throughout the sample.	or more substances in relative amounts that can be varied continuously up to what is called the limit of solubility. The term solution is commonly applied to the liquid state of matter, but solutions of		Normality is a measure of concentration equal to the gram equivalent weight per litre of solution. Gram equivalent weight is the measure of the reactive capacity of a molecule. The solute's role in the reaction determines the solution's normality. Normality is also known as the equivalent concentration of a solution.	In chemistry, colligative properties are those properties of solutions that depend on the ratio of the number of solute particles to the number of solvent molecules in a solution, and not on the nature of the chemical species present The word colligative is derived from the Latin colligatus meaning bound together.
2020/09/21 10:39:08 AM GMT+5:30	nikitakamatkar7 7@gmail.com	A homogeneous mexture is a mixture that has a same properties of its component throughout any given sample a solution is a homogeneous mixture are called as mixture homogeneous.	A solution is a homogeneous mixture of two or more than substance on molecule level	Molality 1. it is defined as the number of moles of solute dissolve in 1000g of the solvent 2.it is denoted by m molarity 1.it is defined as the number of moles of solute dissolve per litre of the solute 2.it is denoted by M	ber of gram equivalent of solute/volume of solution in	

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2020/09/21 10:39:09 AM GMT+5:30	Prasadamrita87 5@gmail.com	A solution is a mixture homogeneous of two or more substance having uniform properties.	mixture of two	Molarity: It is defined as the number of moles of the solute dissolved per litre of the solution. Molality: It is defined as the number of moles of the solute dissolved in 1000g of the solvent.	The normality of the solution is defined as the number of gram equivalent of the solute	Colligative properties are those which depend upon the number of particle of the solute dissolved in known volume of giver solvent and not at all upon the nature of the solute	1
2020/09/21 10:41:16 AM GMT+5:30	tusharasutkar20 00@gmail.com	Homogeneous mixture is solid, liquid or gaseous mixture that has the same properties of its components throughout any given sample	Solution is a homogeneous mixture of two	Molarity:- the number of moles of solute per litre solvent. It is represented by M Molality: Number of moles of solute per kilogram of solvent. It is represented by m.		The properties which depends on the number of moles of solute present in the solution not on the nature of the solute	\
PRINCIPAL Cience & Arts wati, Dist-Cha 2020/09/21 10:44:32 AM GMT+5:30	College ndrapur Pritisunilthawari	It is a mixture in which the component are evenly distributed each other.	It is a homogeneous mixture of two or more substances.	dissolve in per litre of solvent.It is denoted by M.unit:	Normality is defined as the gram equivalent weight of solute dissolve per litre	solute are called	10

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of solution. properties.

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2020/09/21 10:44:38 AM GMT+5:30	prasad.neha221 02001@gmail.c om		present in a smaller amount is called solute,	ratio of number of moles of solute per	Normality is defined as	The properties which depends entirely upon number of particles of solute dissolved in given solvent and not upon chemical composition of solute are called as colligative properties.	
2020/09/21 10:46:57 AM GMT+5:30	tiwarikritesh92 @gmail.com	The mixture of solid, liquid and gas which have same proportion of its component throughout the sample is called homogeneous mixture	The homogeneous mixture of two or more substance is called solution	Molarity is the ratio of the mole of a solute to the total liters of a solution. Morality is the ratio of the moles of a solute to the kilograms of a solvent.	The number of gram equivalent of the solute dissolve per litter of the solution	These are those properties of solution that depend on the ratio of the number of solute particles to the number of solvent molecules in a solution, and not on the nature of chemical species	
2020/09/21 10:52:44 AM GMT+5:30	rupeshbhusari8	same proportions of its components.	homogeneous mixture of two or more	Morality 1.it is the number of moles of solute present in 1000g of the solvent. 2. It is denoted by m. 2.molarity 1.it is defined as the number of moles of solute dissolved per litre of the solution. 2.it is denoted by M.	It is the number of gram equivalent of solute per litre of the solution . It os denoted by	Colligative properties are those which depends upon the number of solute particles in solution and pet on the nature of the solute particles are called colligative properties.	1"

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2020/09/21 11:10:04 AM		A homogeneous mixture is a solid, liquid or gaseous mixture that has the same proportions of its components throughout any	mixture of two or more substances. It is mixture of solute and	Molarity is defined as number of moles of solute dissolved in one litre volume of solution. It's SI unit is mol/ltr. Molality is defined as number of moles of solute dissolved in one kg of solvent. It's SI unit is mol/kg.	Normality is defined as number of grams equivalent of solute dissolved in one litre volume of solution.	The property which depends upon number of solute dissolved in given solvent and does not depend upon nature of solute is called as Colligative property.	100
2020/09/21 11:10:41 AM GMT+5:30	@gmail.com  Pratikshachama tkar@gmail.com	a homogeneous mixture is a mixture that has a same properties of its component throughout any given sample a solution is a homogeneous mixture	A solution is a himogeneous mixture of two or more substance having uniform properties throughout	Molality it ia defined as the nomber of moles of solute dissolve in 1000g of the solvent 2. It is denoted by m molarity it is defined as the number of moles of solute dissolve per litre of the solution 2.it is denoted by M	The normality of the solution is defined as the number of gram equivalent of	Colligative properties defined as those which depends entirely upon the number of particles of the solute dissolved in a known volume of a given solvent and not all upon the nature of the solute these is called as colligative properties	

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2020/09/21 11:22:51 AM GMT+5:30	Lilharekhushbu 7@gmail.com	The reaction in which the catalyst is present in the same phase as the reactant, is called as homogeneous catalysis.	Solution is a mixture of solute and solvent the component which is present in larger quantity is known as called solvent and other one is called solute.	defined as the no. of moles of	The normality of the solution is defined as the no. of gram	Colligative properties are those which depend upon the no. of particle of the solute dissolved in a volume of given solute and not at all upon the nature of the solute.
2020/09/21 11:40:14 AM GMT+5:30	Sapanamandal0 02@gmail.com	A mixture in which it's constituents are distributed uniformly is called as homogeneous mixture	A solution is a homogeneous mixture of two or more substance having uniform properties	Solution and a solvent	The number of gram equivalents of the solute dissolved per litre of the solution	Colligative properties are those which depend entirely upon the number of particles of the solute dissolved in a known volume of a given solvent and not at all upon the nature of the solute
2020/09/21 11:41:06 AM GMT+5:30	runal@amate1@ gmail.com	in which the component mixed are uniformely distributed throughout the	homogeneous mixture of two or more substance the component not mixed and may be varied	molarity is defin as the ratio of no. of mole of solute to the volume of solution in dm3. it is temperature dependance quantity.and molality is defin as the ratio of mole of solute in kg to the mass of solvent in kg		colligative properties means collective properties. these properties of solution depend only no. of solute particle not on the nature of particle

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(Or A. B. Blott)

2020/09/21 11:50:32 AM GMT+5:30	sidhanti.sonewa ne890@gmail.c om	It is type of mixture in which the component mixed are uniformly distributed throughout the mixture. And generally homogeneous mixture means composition of same type of components.	Solution is composition of solute and solvent. Or it is homogeneous mixture of two or more pure substance whose composition can be varied.	is number of moles of the solute dissolved per 1kg of the	Normality: it is the number of gram equivalent of the solute dissolved per litre of solution. It is denoted by N.	Colligative properties are used to define vapour pressure osmosis freezing point qan boiling points	
2020/09/21 12:09:03 PM GMT+5:30	ure711@gmail.c	A homogeneous mixture is a solid, liquid or gaseous mixture that has the same propertion of its components Throughout any given sample.	special type of homogeneous mixture composed of two or more	the ratio of the moles of a solute to the kilograms of a	weight per liter of solutions. Gram equivalents weight is the measure of the reactive capacity of a	Colligative properties are those properties of solutions that depends on the ratio of the number of solute particals to the number of the solvent molecules in solutions, and not are a nature Of the chemical species present.	

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2020/09/21 12:30:51 PM GMT+5:30	Diptibhajanjar2 019@gmail.com	A homogeneous mixture is a solid liquid and gaseous mixture	mixture of solute and	Molarity:1. it is no. Of moles of solute dissolved per litter of solution 2.it is denoted by symbol M. molality:1. it is no. Of mole of solute dissolve in 1000g of solvent 2.it is denoted by m.	equivalent to the solute dissolved per	Colligative properties depend number of particles of the solute it is not depend on nature
2020/09/21 12:10:17 PM GMT+5:30		homogeneous mixture is a solid, liquid or gaseous mixture that has the same proportions of its components throughout any given sample.	The homogeneous mixture of two or more components or solution are known as solution	difference between molality and molarity Image result for difference between molality and molarity View all Molarity is the ratio of the moles of a	normality is described as the number of gram or mole equivalents of solute present in one litre of a solution.	

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the solute it is not depend on nature

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2020/09/21 12:33:08 PM GMT+5:30	Potepriti8@gma il.com	A solution is a homogeneous mixture	Two or more substance having uniform properties is called solution	Molality: it is defined as the number of moles of the solute dissolved in 1000g of the solvent. And molarity: it dissolved per litre of the solution	The normality of the solution is defined as the number of gram equivalent of the solut dissolved per litre of the solution	properties are those which depend entirely upon the number of particles of the solute dissolved in a known volume of a given solvent and not at all upon the nature of the solute. These properties depends upon the nature of the solvent. The various colligative properties are:  1. Lowering of vapour pressure of the solvent.  2. osmotic pressure of the solution.
2020/09/21 12:41:28 PM GMT+5:30	madkechueli@g mail.com	A homogeneous mixture is a mixture which solute substance are evenly spread throughout the entire the solution.	solvent . Solute is a substance	Molality:- molality diffine as the number moles of solute dissolve in 1 KILOGRAM (kg) of solvent. Molarity:- molarity is define as number of moles of solute dossolve in 1 LITER of solvent.	Normality define as equivivalent number of solute dissolve in per liter of	The properties which entirely depends upon number of particles of solutes dissolved in given particular solvent but not upon chemical nature of solute

upon chemical nature of solute

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2020/09/21 2:07:07 PM GMT+5:30	shreyajunarkar @gmail.com	The reaction in which the catalyst in present in the same phase as the reactant is called homhgeneous catalyst.	A solution is a homogeneous mixture of two or more substance on molecular level the longtituent of the mixture present in a smallen amout is called solute.	1)molality=molality is defined as the number of solute per kilogram of solvent. 2)molarity=mol arity is defined as the number of moles of solute per litre of solution.	The normality of the solution	The properties which depend upon the numbers of particles of solute dissolved in given solvent and note upon chemical composition of solute are called colligative properties.	ia
2020/09/21 2:38:33 PM GMT+5:30	poojamadot@g mail.com	A homogeneous mixture is a solide, liquid or gaseous mixture that has the same proportions of its components throughout any given sample.		An important distinction between a solution and a solvent. Molarity is the ratio of the mole of a solute to the total liters of solution Molarity on the other hand is the ratio of the moles of a solute to the kilogram of a solvent	Normality is a measure of concentration equal to the gram equivalent weight per liter of solution. Gram equivalent weight is the measure of the reactive capacity of a molecule.	Colligative properties are those properties of solutions that depend on the ratio of the number of solvent molecules in a solution and not an the nature of the chemical species present.	fo
2020/09/22 10:00:47 AM GMT+5:30	Pranaiidate3357 0@gamil.com	with solvent is called	The mixture of solute and solvent is called solution		of the solution is called	The properties wich depend entirely upon of particles of solute dissolved in solvent and chemical composition of soltute are called colligative properties	10

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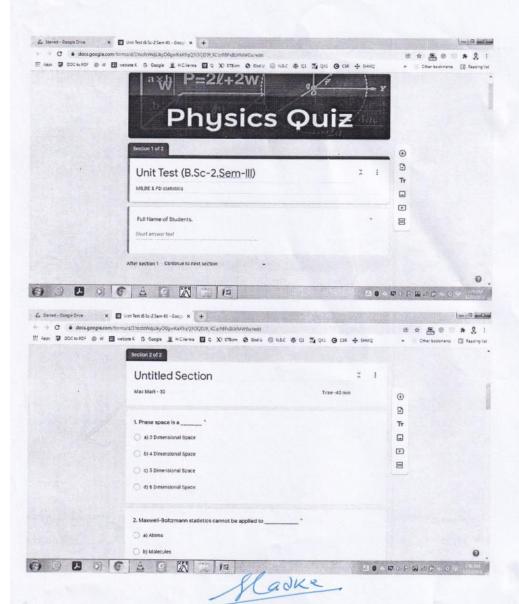
# Unit fest B. Sc. II, Sem III, 01/10/2020

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of the	2	19	3	3	3	8 7	35
Which of the following observations is incorrect about the order of a reaction?	(a) Order of a reaction is always	(c) Order of reaction is the sum of power to express the rate of reaction to the concentration reterns of the reactants.	(b) The stoichiometric coefficient of the reactants the order the order the order the order the order	(c) Order of reaction is the sum of power to express the rate of reaction to the concentration reactions of the reactions.	(d) Order can only be assessed experimentally	ect se	A SHA MA
The rate constant of zero order reactions has the unit	(a) s-1	(a) 5-1	(d) L mol-1 s-1	(b) mol L-1 s-1	(5) mol L-1 s-1	(c) L2 mol-2 s-1	
In the reaction 2A + B ât' A2B, if the concentration of A is doubled and that of B is halved, then the rate of the	(a) increase 2 times	(a) increase 2 times	(a) increase 2 times	(d) remain the	(a) increase 2 times	(a) increase 2 times	
The reactant Vs time concentration diagram for a reaction is a straight line with a negative pendulum. The reaction follows an equation for the intensity.	(b) first order	(a) zero order	(a) Zero order	Option 2	(c) second order	(a) zero order	ke College
The rate of a chemical reaction tells us about	(c) how slow or fast the reaction is taking place	(d) none of the above	(c) how slow or fast the reaction is taking place	a) the reactants taking payr in the reaction	(d) none of the above	(c) how slow or fast the reaction is taking place	Prigs. Ladke PRINCIPAL N.S. Science & Arts College Bhadrawati Diet Change
For a chemical reaction Ag+18, it is found that the rate of reaction doubles when the concentration of A is increased four times. The order of reaction is	(a) Two	(c) Half	(c) Half	(b) One	(c) Half	(c) Half	N.S. Scie
For Property of Pr	(a) Second order	(a) Second order	c) Pseudo- unimolecular	(a) Second order	c) Pseudo- unimolecular	c) Pseudo- unimolecyjar	
. Which among the following is a false statement?	c) Molecularity of a reaction may be zero or fraction.	c) Molecularity of a reaction may be zero or fraction.	(b) Half life of a third order reaction is inversely proportional to square of initial concentration of the reactant.	(a) Rate of zero order reaction is independent of initial concentration of reactant.		(b) Half life of a third order reaction is inversely proportional to square of initial concentration of the reactant.	
The half life of the first order reaction having rate constant K = 1.7 x 10-5s-1 is	(d) 1.8 h	(d) 1.8 h	c) 11.3 h	(d) 1.8 h	c) 11.3 h	c) 11.3 h	
. What will be the fraction of molecules having energy equal to or greater than activation energy, Ea?	(a) K	(d) e-Ea/Rt	(c) Ae-Ea/Br	(b) A	(c) Ae-Ea/Rť	(c) Ae-Ea/Rt	
. Which of the following statements about the catalyst is true?y.	(a) A catalyst accelerates the rate of reaction by bringing down the Sameer Nasnurkar activation energy	(d) A catalyst makes equilibrium constant more favourable for forward reaction.	(b) A catalyst does not participate in reaction mechanism.	(a) A catalyst accelerates the rate of reaction by bringing down the activation energy	(b) A catalyst does not participate in reaction mechanism.	(b) A catalyst does not participate in reaction mechanism.	
Email of student	Sameer Nasnurkar		Pratiksha zade	Achal raut	Parikshita Chandrapal Kakde	Tushar Thakur	
Timestamp	2020/10/01 11:58:07 AM GMT+5:30	2020/10/01 12:01:07 PM GMT+5:30	2020/10/01 12:01:54 PM GMT+5:30	2020/10/01 12:02:54 PM GMT+5:30	2020/10/01 12:03:55 PM GMT+5:30	2020/10/01 12:06:37 PM GMT+5:30	

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N.	(b) The stoichiometric coefficient of the reactants doesnât** taffect	the order  (b) The stoichiometric coefficient of the reactants doesn't affect	the order  (a) Order of a reaction is always	, amore manage	(a) Order of a reaction is always	a whose number (d) Order can only be assessed experimentally	(d) Order can only be assessed experimentally
	(A)	L-5 (5) 1011 (6)	(f) L2 mol-2 s-1	(h) mol 1.1 c.1	1.00 Jones (16)	(b) mol (-1.5-1	(b) moi t-1 s-1
	(d) remain the	(c) decrease 2	times  (a) increase 2.	(a) increase 2	(a) increase 2	(a) Increase 2 times	(a) increase 2 times
	(b) first order	The state of the s	it) second order	(a) zero order	(a) zero order	(a) zero order	(a) Zero order
	a) the reactants taking part in the reaction	(d) none of the	(b) the products formed in the reaction	(b) the products formed in the reaction	(c) how slow or fast the reaction is taking place	(b) the products formed in the reaction	(b) the products formed in the reaction
	(b) One	ener(I)		(c) Half	(b) One	(c) Half	(c) Half
	c) Pseudo- unimolecular	(a) Second order	c) Pseudo- unimolecular	c) Pseudo- unimolecular	c) Pseudo- unimolecular	c) Pseudo- unimolecular	lar
	c) Molecularity of a reaction may be c) Pseudo-zero or fraction. unimolecu	(b) Half life of a third order reaction is inversely proportional to square of initial concentration of the reactant.	a to	c) Molecularity of a reaction may be c) Pseudo- zero or fraçtion. unimolecu	c) Molecularity of a reaction may be c) Pseudo- zero or fraction. unimolecu	c) Molecularity of a reaction may be c) Pseudo- zero or fraction. unimolecu	c) Molecularity of a reaction may be (c) Pseudo- zero or fraction. Iunimolecu
	b) 9.7 h	(a) 12.1h	c) 11.3 h	c) 11.3 h	c) 11.3 h	c) 11.3 h	0,11.37h
	( <u>e</u> )	(c) Ae-Ea/Rt	(d) & Ea/Rt	(d) e-Ea/Rt	(d) e-Ea/Rt	(d) e-Ea/Rt	(d) e-Ea/Rt
	(a) A catalyst accelerates the rate of reaction by snehakamble1920 bringing down the 01@gmail.com	(b) A catalyst does not participate in reaction mechanism.	(a) A catalyst accelerates the rate of reaction by bringing down the activation energy	(a) A catalyst accelerates the rate of reaction by bringing down the activation energy	(a) A catalyst accelerates the rate of reaction by bringing down the activation energy		(a) A catalyst accelerates the rate of reaction by bringing down the activation energy (
	snehakamble1920 01@gmail.com	Niima askar	(a) A catalyst accelerates the rate of reaction by Aabeda Fatema bringing down the Israril Khan Pathan activation energy	(a) A catalyst accelerates the rate of reaction by Pritisunlithawani 2 bringing down the G@gmail.com activation_efeegy	mayuriwatkar062   000@gmail.com	(a) A catalyst accelerates the rate of creation by Patikshachamatk bringing down the ar@gmail.com activation energy	(a) A catalyst accelerates the accelerates the rate of reaction by Nikitakamatka77 bringing down the @gmail.com activation energy
	2020/10/01 12:06:52 PM GMT+5:30	2020/10/01 12:10:19 PM GMT+5:30	2020/10/01 12:10:34 PM GMT+5:30	2020/10/01 12:17:06 PM GMT+5:30	2020/10/01 12:23:48 PM GMT+5:30	2020/10/01 12:24:35 PM GMT+5:30	2020/10/01 12:26:09 PM GMT+5:30

8	B	8	3	8	13	3 -
(d) Order can only be assessed experimentally	(d) Order can only be assessed experimentally	(a) Order of a reaction is always a whole number	(a) Order of a reaction is always a whole number	(b) The stoichiometric coefficient of the reactants doesn't affect the order	(b) The stoichiometric coefficient of the reactants doesnât <sup>®</sup> t affect the order	(a) Order of a reaction is always a whole number
(b) mol L-1 s-1	(d) L mol-1 s-1	(a) s-1	(a) s-1	(b) mol t-1 s-1	(b) mol L-1 s-1	(b) mol L-1 s-1
(a) increase 2 times	(a) increase 2 times	(d) remain the same	(d) remain the same	(a) increase 2 times	(a) increase 2 times	(c) decrease 2 times
(a) zero order	(a) zero order	(a) zero order	(c) second order	(a)-zero order	(a) zero order	(d) third order
(b) the products formed in the reaction	(b) the products formed in the reaction	(c) how slow or fast the reaction is taking place	(c) how slow or fast the reaction is taking place	(d) none of the above	(d) none of the above	a) the reactants taking part in the reaction
(c) Half	(c) Half	(c) Half	(b) One	(d) Zero	(d) Zero	(c) Half
ar	c) Pseudo- unimolecular	(b) Unimolecular	(b) Unimofecular	(b) Unimolecular	(b) Unimolecular	<u></u>
c) Molecularity of a reaction may be c, Pseudo- zero or fraction. unimolecu	c) Molecularity of a reaction may be c) Pseudo- zero or fraction. unimolecu	c) Molecularity of a reaction may be zero or fraction.	c) Molecularity of a reaction may be zero or fraction.	(a) Rate of zero order reaction is independent of initial concentration of reactant.	(a) Rate of zero order reaction is independent of initial concentration of reactant.	c) Molecularity of a reaction may be c) Pseudo- zero or fraction. Unimolecular
c) 11.3 h	c) 11.3 h	c) 11.3 h	c) 11.3 h	c) 11.3 h	c) 11.3 h	(d) 1.8 h
(d) e-Ea/Rt	(d) e-Ea/Rt	(d) e-Ea/Rt	(d) e-Ea/Rt	(a) K	(a) K	(a) K
(a) A catalyst accelerates the rate of reaction by priyankayergude2 bringing down the 001@gmail.com activation energy.	(a) A catalyst accelerates the rate of reaction by bringing down the activation energy	(a) A catalyst accelerates the rate of reaction by bringing down the activation energy	> 0	> 0	(a) A catalyst accelerates the rate of reaction by bringing down the 7385286869 activation energy	(a) A catalyst accelerates the rate of reaction by bringing down the activation energy
priyankayergude2 001@gmail.com	Khushbu lilhare	Ritesh tiwari	(a) A catalyst accelerates the rate of reaction by prasad neha2210 bringing down th 2001@gmail.com activation energy.	mehandichalkhur e711@gmail.com	7385286869	rupeshbhusari84 @gmail.com
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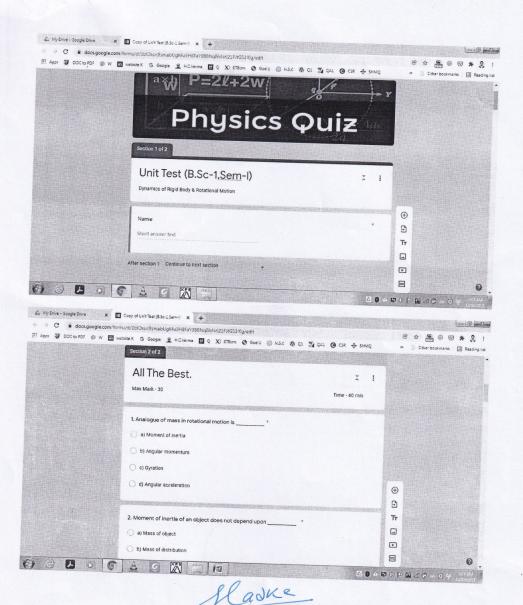
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(d) Order can only be assessed	(c) Order of reaction is the sum of power to express the rate of reaction to the concentration terms of the reactants of the reactants.	(b) The stoichiometric coefficient of the reactants doesnâe*** affect the order	(b) The stoichiometric coefficient of the reactants doesnâtent	the order (c) Order of reaction is the sum of power to express the rate of reaction to the concentration terms of the concentration reactions of the concentration reaction reactions of the concentration reactions of the concentration reaction	(a) Order of a reaction is always a whole number
(b) mol 1-1 s-1	(8)	(b) mol (-1 s-1	<u> </u>	(a) 5-1	(b) mol L-1 s-1
(a) increase 2	(a) increase 2 times	(a) increase 2 times	(b) increase 4	antes (a)-increase 2 times	(a) increase 2 times
(a) zero order	(d) third order	Option 2	approximation (a)	(d) third order	(a) zero order
(c) how slow or fast the reaction is taking place	(c) how slow or fast the reaction is taking place	(d) none of the above	(c) how slow or fast the reaction is taking place	(c) how slow or fast the reaction is taking place	(c) how slow or fast the reaction is taking place
(c) Haif	(c) Half	ego (q	ano (f)	c) Half	) Half
(a) Second order	(b) Unimolecular	c) Pseudo- unimolecular	a) Second order		c) Molecularity of a reaction may be second order (c) Half
c) Molecularity of a reaction may be zero or fraction.	c) Molecularity of a reaction may be zero or fraction.	(d) For a first order reaction, t1/2=0.693K	c) Molecularity of a reaction may be zero or fraction.	- e	c) Molecularity of a reaction may be zero or fraction.
c) 11.3 h	c) 11.3 h	(d) 1.8 h	93.7 h	c) 11.3 h	c) 11.3 h
(d) e-Ea/Rt	(d) e-Ea/Rt	(a) K	(a) K	(d) e-Ea/Rt	
(a) A catalyst accelerates the rate of reaction by bringing down the activation energy (d) e-Ea/Rt	(a) A catalyst accelerates the acted action by bringing down the activation energy (d) e-Ea/Rt	(c) A catalyst makes the reaction feasible by making â*fG more negative	(a) A catalyst accelerates the rate of reaction by bringing down the activation energy (a) K	(a) A catalyst accelerates the rate of reaction by bringing down the activation energy (d) e-Ea/Rt	(a) A catalyst accelerates the rate of reaction by bringing down the activation energy (a) K
achalkannaujiya1 11@gmail.com	Nutan sharad korade	(c) A catalyst makes the reaction feasible Priyasah652001@ by making â*16 gmail.com more negative	Chueli Madke	Harshad askar	
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							UNIT	UNIT TEST- B.Sc -2 (SEM-3), (MB,	-2 (SEM-3),	BE	& FD STATISTICS)	STICS)			Jemi	T
American sacout carried associate carried with a Stiffing's earth of Stiffing's earth	average kinetic kinetic hinetic energy associat ed with goldzmann degree statistics cannot of treadom	average kinetic energy associat ed with each degree of	A Strill	g's natio	5. The equation for 6. Bosons the have evaluation symmetrics of \$in M.B. waven distribution functions functions (1 they do not not the part of the par	= =	7. Bose- Einstein statistics is for	8. The difference between fermions and Ebosons is that Ebosons, wave	9. Bose- Einstein statistics can be applied to	n ned	11. During Bose- Einstein condensati on all the atoms fall 1 back to the s	12. Fermi-Dirac statistics is for it	At T > he he ability state E > EF is	ic gy of sle is nden	e e a . cs	, (
is a		si				obey b) Pauli's	the			one particle	N First	c) Particles with	Gero.	ure only.		90
bhagawan Dimensional d) Lattice b) 2kT b) N/2InN+2	b) 2kT		b) N/2Inf		c) B = KT	Principle b) Pauli's		d) Differentiable	a) Electrons	b) False		200	b) False	b) False	а) Тгие	6/30
Sameer Dimensional Ophotons c) KT/2 a) NInN-N Nasnurkar Soace c) Photons c) KT/2 a) NInN-N	c) KT/2		a) NinN-I	7	a) ß = 1/kT	Exclusion	able	c) Symmetric	b) Photons	a) True	a) Ground C	Distinguishable	a) True	a) True	b) False	24 / 30
d) 6 Dimensional	b) 2kT	1	b) N/2Inf	4+2	a) B = 1/kT	b) Pauli's Exclusion Principle	d) Particles with integral spin	c) Symmetric	b) Photons	b) False	a) Ground c	c) Particles with half integral spin	b) False	a) True	b) False	22 / 30
a) b nnath Dimensional	C) KTD		-NoiN (e	z	a) B = 1/kT		o) Particles with integral spin	c) Symmetric	b) Photons	b) False	b) First exited state	c) Particles with half integral spin	b) False	a) True	b) False	24 / 30
ional contract in 2017	T-10-14		Nel		a) R = 1/kT		d) Particles with integral soin	c) Symmetric	c) Fermions	а) Тгие	a) Ground It	a) Distinguishable particles	a) True	a) True	b) False	18 / 30
5) 4 Dimensional	T4 (6		Null	Z	d) 8 = 2/kT	d) Heisenberg's Uncertainty		e e	b) Photons	a) True	b) First exited state	a) Distinguishable particles	a) True	a) True	a) True	8/30
ional on Lattice (9) kT	9 6		NoIN (e	2	a) 8 = 1/kT	b) Pauli s Exclusion Principle	o) Particles with integral spin	c) Symmetric	b) Photons	a) True	a) Ground state	c) Particles with half integral spin	b) False	а) Тгие	b) False	24/30
dilip Dimensional	OTA (2	1	Anicin id	5		b) Pauli s Exclusion Principle	o) Particles with integral spin	c) Symmetric	c) Fermions	b) False	a) Ground state	a) Distinguishable particles	b) False	a) True	b) False	18/30
c) KT/2	c) KT/2		d) NinN+ñ	7		Rule of Maximum Multiplicity	b) Symmetrical Particles	b) Single Valued d) Protons	d) Protons	b) False	d) Insufficient information	a) Distinguishable particles	a) True	b) False	b) False	6/30
ional c) Photons b) 2KT	b) 2KT		c) InNi		a) ß = 1/kT	b) Pauli s Exclusion Principle	c) Particles with half integral spin	c) Symmetric	c) Fermions	a) True	a) Ground state	c) Particles with half integral spin	a) True	a) True	a) True	18/30
andrao Dimensional Space c) Photons	c) KT/2			N	b) ß = -1/KT		a) Particles with integral spin	c) Symmetric	b) Photons	a) True	a) Ground state	c) Particles with half integral spin	b) False	b) False	b) False	24/30
Tushar a) 3 Tushar a) 3 Tushar a) 3 Tushar a) 3 Tushar armarethwar Dimensional a) Atoms a) MinN-N	a) KT		a) NinN-N		a) ß = 1/kT	a) Aufbau principle	a) Distinguishable particles	a) Continuous	a) Electrons	а) Тгие	a) Ground state	a) Distinguishable particles	a) True	a) True	а) Тгие	10 / 30
d) 6 bhay Dimensional e Space	c) KT/2		c) InN!		a) ß = 1/KT	b) Pauli s Exclusion Principle	a) Particles with integral spin	c) Symmetric	d) Protons	b) False	a) Ground state	c) Particles with half integral spin	b) False	a) True	a) True	22 / 30
anjiv	d) Lattice c) kT/2				a) ß = 1/kT	b) Pauli's Exclusion Principle	d) Particles with integral spin	c) Symmetric	b) Photons	a) True	a) Ground state	c) Particles with half integral spin	b) False	а) Тгие	b) False	28 / 30
d) 6 nna Dimensional o thakare Space c) Photons c) KT/2	c) KT/2			N+2	a) ß = 1/kT	b) Pauli's Exclusion Principle	o) Particles with integral spin	c) Symmetric	d) Protons	a) True	a) Ground state	c) Particles with half integral spin	b) False	b) False	а) Тгие	22/30
	c) KT/2			Z	a) ß = 1/kT	Exclusion Principle	u) rancies with integral spin	c) Symmetric	b) Photons	a) True	a) Ground state	c) Particles with half integral spin	b) False	a) True	b) False	30 / 30
a) 3 unii Dimensional Ci Photons	c) KT/2			Z+N-	a) B = 1/KT		b) Symmetrical Particles	b) Single Valued c) Fermions	c) Fermions	a) True	b) First exited state	Distinguishable particles	b) False	a) True	а) Тгие	14/30
0) b Dimensional	b) 2kT			Z			a) Distinguishable particles	a) Continuous	a) Electrons	b) False	a) Ground state	Distinguishable particles	b) False	а) Тгие	a) True	16 / 30
Space (c) Friotons (b) 2KI	c) Photons C) KT/2			*Nu			d) Particles with integral spin	c) Symmetric	b) Photons	a) True	a) Ground state	a) Distinguishable particles	b) False	a) True	b) False	26 / 30
nwar Space C) Filotolis C) NIZ.  a) 3 nesh Dimensional C) Photons a) KT	C) Photons a) kT			_			d) Particles with integral spin	c) Symmetric	d) Protons	a) True	a) Ground state	a) Distinguishable particles	a) True	a) True	а) Тгие	16/30
a) kT	c) Photons a) kT			Z	a) ß = 1/KT		a) Distinguishable particles	c) Symn	b) Photons	a) True	b) First exited state	Distinguishable	a) True	а) Тие	a) True	16/30
							N	8/10								

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	n b) False	100	a) True	a) True	b) False	a) True					b) False b	b) False	
	c) Particles with half integral spin b) False	c) Particles with	a) Distinguishable particles	c) Particles with half integral spin a) True	Distinguishable particles	Distinguishable particles	c) Particles with	c) Particles with	half integral spin b) False	c) Particles with half integral spin b) False	d) Particles with integral spin b	Distinguishable particles	Ondial
	a) Ground state	a) Ground state	a) Ground state	a) Ground state	b) First exited state	a) Ground state	a) Ground state	pun	State				a) Ground (a)
	a) True	a) True	a) True	a) True	a) True	a) True	b) False	a) True		а) Тпе	a) True	a) True	31
	b) Photons	b) Photons	b) Photons	a) Electrons	b) Photons	c) Fermions	b) Photons	b) Photons		d) Protons	b) Photons	c) Fermions	100
	c) Symmetric	c) Symmetric	c) Symmetric	c) Symmetric	a) Continuous	c) Symmetric	c) Symmetric	c) Symmetric	0.0	1			C) Symmetric
0) Particles with integral	spin spin	Particles d) Particles	spin a) Particles with integral	spin b) Symmetrical	cles		Distinguishable particles d) Particles		_	1_	ricies	o) Particles with integral	
				C) Hund's Rule of	Maximum b) Pauirs Exclusion	b) Pauli's	Principle b) Pauli's		_	Exclusion w		-	1
			b) N/2inN+2 a) R = 1/kT	a) R = 10.7	a) R = 40.7	1 b - 1 K	b) B = -1/kT		a) B = 1/kT	c) B = kT	c) B = KT		
ING (S	C) KT/2 a) NIPM N	N-Allian (a)	b) N/ZinN+2	a) NinN-N		1		c) kT/2 b) N/2InN+2 c) B = kT	a) NInN-N		c) InNi	c) kT/2 (a) NinN-N (a) B = 1/kT	
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c) Photons	d) Lattice	c) Photons	d) Lattice	c) Photons	b) Molecules	i	c) Photons	c) Photons	c) Photons	c) Photons	c) Photons	c) Photons	
a) 3 Dimensional Space	d) 6 Dimensional Space 0) 6	Dimensional	Dimensional Space a) 3	Olmensional Space d) 6	Dimensional Space	d) 6 Dimensional Space	ional	o) 6 Dimensional	lenoi			Space	
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Dr. L. S. Ladke
PRINCIPAL
Science & Arts Colle

N.S. Science & Arts College Bhadrawati, Dist-Chandrapur

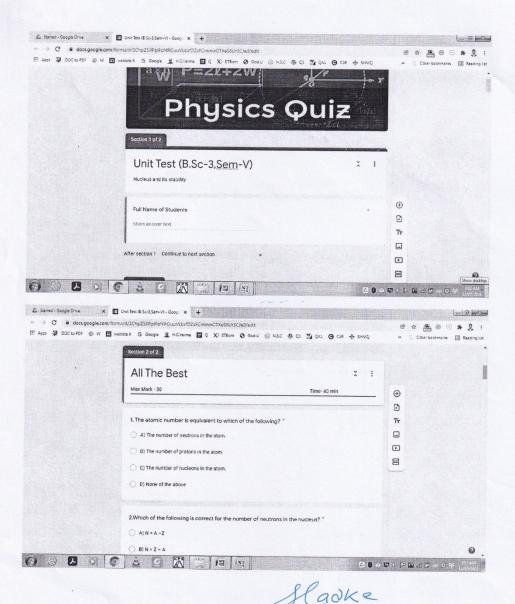
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-	15. Unit of Frequenc y is.	D) Both	A) Hertz	A) Hertz	e A) Hertz	A) Hertz	A) Hertz	A) Hertz	A) Hertz	e A) Hertz	o A) Hertz	e A) Hertz	o e. A) Hertz	al D) Both	A) Hertz	A) Hertz	
	14. Homogen eity maans ?	B) - Translatio ncl Invarience	C) Both	B) Translatio nal Invarience	B) Translatio nal Invarience	C) Both	C) Both	C) Both	C) Both	Translatio nal Invarience			B) Translatio nal Irvarience	A) Rotational Invarience	D) None of these	C) Both	D) None
1000	13. Which of the following leads to the law of conservation of angular momentum?	a) isotropy of space	b) Homogeneity of time	b) Homogenetty of time	b) Homogeneity of time	b) Homogeneity of time	b) Homogeneity of time	b) Homogeneity of time	b) Homogeneity of time	a) isotropy of space	a) Isotropy of space	a) isotropy of space	a) Isotropy of space	c) Nucle ar	d) Gravity	d) Gravity	(q)
	12. Which of the following is unit of Angular Momentum.	C) Kg m³/ sec	C) Kg m³/ sec	C) Kg m²/ sec	C) Kg m²/ sec	C) Kg m²/ sec	C) Kg m³/ sec	C) Kg m²/ sec	C) Kg m³/ sec	A) Kg m² sec.	C) Kg m²/ sec	C) Kg m³/ sec	D) 1/ Kg m² sec	C) Kg m³/ sec	C) Kg m²/ sec	C) Kg m²/ sec	
	11. Which of the following term is defined as. The product of toque and time for which it is applied".	B) Angular Impulse	B) Angular Impulse	B) Angular Impulse	B) Angular Impulse	B) Angular Impulse	B) Angular Impulse	B) Angular Impulse	B) Angular Impulse	B) Angular Impulse	B) Angular Impulse	B) Angular Impulse	B) Angular Impulse	B) Angular Impulse	B) Angular Impulse	B) Angular Impulse	
	10. Joule is the unit of	C) Energy	C) Energy	C) Energy	C) Energy	C) Energy	C) Energy	C) Energy	C) Energy	C) Energy	C) Energy	C) Energy	C) Energy	C) Energy	A) Pressur e	C) Energy	
	9. The dimensions of kinetic energy is	B) [M L*	B) [M L² T-3]	B) [M L² T~3]	B) [M L² T-²]	B) [M L² T-3]	В) [М С <sup>2</sup> Т <sup>-3</sup> ]	B) [M L² T-3]	B) [M L² T-²]	B) [M L? T~2]	B) [M L² T-*]	B) [M L² T-2]	B) [M L² T-²]	B) [M L² T-3]	B) [M L² T*3	B) [M L² T-²]	
	8.Moment of Inertia of solid sphere of about its diameter is.	A) (2/5)MR*	A) (2/5)MR"	A) (2/5)MR*	A) (2/5)MR <sup>2</sup>	A) (2/5)MR*	A) (2/5)MR*	A) (2/5)MR*	A) (2/5)MR*	A) (2/5)MR*	A) (2/5)MR²	A) (2/5)MR*	A) (2/5)MR²	A) (2/5)MR²	A) (2/5)MR²	A) (2/5)MR <sup>2</sup>	
	6. Torque is the rotational analogue 7. Which relation is not analogue 07. Which relation is not of.	D) Linear momentum=Moment of inertia X angular velocity	C) Moment of inertia≕Torque X angularacceleration	C) Moment of inertia=Torque X angularacceleration	A) Torque=Moment of inertia X angular acceleration	C) Moment of inertia=Torque X angularacceleration	A) Torque=Moment of inertia X angular acceleration	C) Moment of inertia=Torque X angularacceleration	D) Linear momentum=Moment of inertia X angular velocity	B) Torque=Dipole moment X magnetic induction	C) Moment of inertia=Torque X angularacceleration	C) Moment of					
	6. Torque is the rotational analogue 7 of.	(D) force	D) force a	(C) C accelerat it	(D) force	(D) force	(D) force	(D) force	(D) force	(D) force	(D) force		(C) accelerat ion	(D) force	(D) force	(D) force	
	5. The moment of nertia of a ring of mass m and radius r about an axis, passing through the certite and perpendicular to the plane of the ring, is:	(a) (1/2) m/2	(b) m/2	(b) m/2	(a) (1/2) m/2	(b) m/2	(b) m/2	(b) mr2	(b) m/2	(b) m/2	(b) m/2	(9)	(b) m/2	e e	(b) m/2	(b) mr2	
	4. What is the formula of theorem of perpendicular axis?	c) tzz – txx = tyy	c) fzz – lxx = lyy	d) None of the mentioned	d) None of the mentioned	d) None of the mentioned	d) None of the mentioned	c) tzz – box = tyy	c) lzz – lox = lyy	c) tzz – bx = lyy	d) None of t the mentioned	d) None of the mentioned	d) None of the mentioned	d) None of			
	3.Which of the following is also known as the law of inertia?	d) Newton's first (c) Izz – box law of motion  = lyy	d) Newton's first of	d) Newton's first law of motion	d) Newton's first law of motion	d) Newton's first law of motion	d) Newton's first law of motion	d) Newton's first law of motion	d) Newton's first law of motion	a) Newton's second law of motion	d) Newton's first 1	d) Newton's first law of motion	d) Newton's first law of motion	d) Newton's first law of motion	d) Newton's first law of motion	d) Newton's first law of motion	
	2. Moment of inertia of an object does not depend upon	c) Angular velocity	c) Angular velocity	c) Angular velocity	c) Angular velocity	c) Angular velocity		c) Angular velocity	c) Angular velocity	c) Angular velocity	c) Angular	c) Angular velocity	c) Angular velocity	c) Angular velocity	b) Mass of distribution	c) Angular velocity	
1	1. Analogue of mass in rotational motion is	a) Moment of inertia	a) Moment of inertia	a) Moment of inertia	a) Moment of inertia	a) Moment of inertia	nent of	a) Moment of inertia	a) Moment of inertia	a) Moment of inertia	a) Moment of inertia	nent of	a) Moment of inertia	a) Moment of inertia	a) Moment of inertia	a) Moment of inertia	
	Name	Anushka Raju Jawale.	Rahul sanjay masalkar	Rajeshwari shankar sundargiri	Dipali Sunil Boinwar	KUNAL SUNIL DONGE	Pranjali Gamtidas Raipure	Prasad Subhash vidhate	Vipul Govinda Nande	Prajakta Natthu 7 Bodhane	Vibha santosh Tated	Alishanaaz Abdul Salam Sheikh	Divya Dipak 6 Vyavahare	Trwari Alshwarya 7 Rajivkumar	Mahesh chindhuji 2 bawane	Achal Ravindra 5 Jidgalwar	
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D) None		-	A) Rotational Invarience	B) Translatio nal	A) Rotational	A) Rotational	B) Translatio	40	(a) Rotational	9	3) ranslatio	() () () () () () () () () () () () () (	() totational	ranslatio	) ranslatio	) ranslatio al	anslatio	Invarience A) Hertz 16/30 A) Rotational
opy of	opy of		b) Homogeneity of time	opy of	Jo Ade			b) Homogenetty	neity		o) Nuclear	70		I) Isotropy of n				
C) Kg m²/ sec	C) Kg m²/							D) 1/ Kg m²					0	A) Kg m² sec s	C) Kg m²/ o		1	
B) Angular Impulse	B) Angular Impulse	B) Angular Impulse	B) Angular Impulse	B) Angular Impulse	A) Momentum	A) Momentum	3) Angular mpulse	A) Momentum	3) Angular moulse									
C) Energy	C) Energy	A) Pressur e	C) Energy									Pressur B			;)	;)	ressur B	A) Pressur
B) [M L?	B) [M L²		3) [M <sup>C</sup>		THE COURSE								1 10000		N.C.	[M L* T*] E	IML" PA	× 0
A) (2/5)MR*	A) (2/5)MR*	A) (2/5)MR*	A) (2/5)MR*	A) (2/5)MR*			A) (2/5)MR*		8 4) (2/5)MR* T					C) (2/3)MR* T*		) (2/5)MR° D)	(A)	
4) Torque=Moment of nertia X angular scceleration	A) Torque=Moment of nertia X angular scceleration	) Moment of nertia=Torque X ingularacceleration	3) Torque=Dipole noment X magnetic nduction	) Torque=Dipole noment X magnetic nduction	u		) Linear tomentum=Moment of tertia X angular elocity	The last	on		m=Moment of ingular	tio	op					
(C) accelerat on		7					3) linear nomentu ir		3) linear Comentu in	C in		-				celerat ine	B) me	(C) A) accelerat ine
(b) mr2	(b) m/2	(b) m/2	(b) mr2	(b) mr2 (d)	(b) mr2 (d)	(b) m/2 (d)	(b) mr2	тіг	(B) m/2	) m2			m/2	mr2	m/2	mr2		O)
d) None of the mentioned	d) None of the mentioned	c) lzz – lox = lyy	c) lzz – lox = lyy	a) lzz = bx - lyy	c) izz – lox = iyy	c) [22 – lox = lyy	c) lzz – bx = lyy	c) lzz – lxx = lyy	c) tzz – bx = lyy	d) None of the mentioned	lone of				) None of he nentioned (			
d) Newton's first law of motion	d) Newton's first law of motion	d) Newton's first law of motion	d) Newton's first law of motion	a) Newton's second law of motion	d) Newton's first law of motion	d) Newton's first law of motion	d) Newton's first law of motion	d) Newton's first law of motion	d) Newton's first law of motion	d) Newton's first taw of motion				d) Newton's first a	) Newton's first ti aw of motion	I) Newton's first b aw of motion +	) Newton's first b aw of motion +	d) Newton's first c) izz – lxx
c) Angular velocity	b) Mass of distribution	c) Angular velocity	a) Mass of object	c) Angular velocity	a) Mass of object	a) Mass of object	c) Angular velocity	c) Angular velocity		c) Angular relocity			) Angular elocity			1 200		c) Angular
a) Moment of inertia	a) Moment of inertia	a) Moment of nertia	a) Moment of nertia	a) Moment of nertia	I) Moment of nertia			) Moment of sertia	) Moment of ertia	Moment of ertia	Moment of c	Moment of c	Moment of certia		Moment of d	Moment of c	Moment of d	b) Angular c
Prachi suresh bagade	Khushi Kharwade		Priya Raju Ambilkar	=	7					Vikanksha a			pitika Kishor a) lurge	Aayuri dipak b)		cep	ayalbodhe inc	-
3/4/2021 15:28:38	3/4/2021 21:16:08	3/5/2021 10:20:10	3/5/2021 11:11:00	3/5/2021 11:30:42	3/5/2021 11:36:02	3/5/2021 11:36:06	3/5/2021 13:10:45 y	3/5/2021 13:15:31 k	3/5/2021 13:44.23 T	3/5/2021 15:58:28 P	3/5/2021 20:01:31 8	3/6/2021 8:32:32 kt	3/6/2021 10:10:38 di	3/6/2021 16:56:37 by	3/72021 10:33:27 St	94 Sh 3/9/2021 14:11:45 ka	3/9/2021 19:00:27 Pa	Jyoti kashinati 3/9/2021 19:34:08 kannake
	Perbot a) Moment of c) Angular of Newton's rest the first and restored from mentioned (b) m/2 on a sceleration investigation of the first and restored from the first and	Public   P	Public   P	Solution   Solution	Solution   Solution	A	Background   Colored   C	Bottometic of Angular   Otherwise first	Bottometric   Coloregia   Co	Barbon   Control   Contr	Prach   Strong   Other   Colored   Other   Colored   Other   Other   Colored   Other   Other	Property   Character   Chara	President   Marchella   Marc	Part	Part	Particular   1 December   1 D	Particle   Notice   Organization   Notice   Organization   Organ	Part   Part

N.S. Science & Arts College
Bhadrawati, Dist-Chandrapur

Sandeep Surendra 3/9/2021 21:40:13 Singh	m .=	) Moment of a) Mass of nertia object	b) Mass of d) Newton's first c) lzz – bx bect law of motion = lyy (b) m/2	c) tzz – bx = lyy	(b) m(2	(D) force	A) Torque=Moment of inertia X angular acceleration	A) (2/5)MR*	9) [M L	C) Energy	C) B) Angular Energy Impulse	A) Kg m² sec	a) Isotropy of Rotational Space Invarience	A) Rotational Invarience A) H	Hertz 16/7	0
Gaurav 3/10/2021 7:10:35 Gajanan jogi		) Moment of d) Axis of retria	b) Newton's third law of motion	c) izz – ixx = iyy (b) mr2	b) m/2	(A) kinetic	A) Torque=Moment of inertia X angular acceleration	B) (2/5)MR*	C) [M L*	A) Pressur e	A) Pressur B) Angular I mpulse	C) Kg m²/ sec	b) Homogeneity of time	A) Rotational Invarience A) H	Hertz 147.	0
Atik 110/2021 17:57:58 kshirsad	a) Moment of	c) Angular	c) Angular d) Newton's first c) izz – bx velocity law of motion = lyy (t	c) 122 – 10x = 1yv	(b) m/2	(D) force	C) Moment of inertia=Torque X angularacceleration	A) (2/5)MR*	B) [M L?	C) Energy	B) [M L² C) B) Angular T-1 Energy Impulse	C) Kg m³/ sec	C) Kg m³/ a) isotropy of r	B) Translatio nal Inverience D) E	Both 267	Q

(Mr K.V. Bhowfale)



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	15. What makes the best moderator s in nuclear power plant?	d) Materials with high mass	c) Materials with high atomic number	b) Materials with low atomic	b) Materials with low atomic	a) Material with low atomic	a) Material With low atomic number	a) Material with low atomic	The second second	1
	14. What type of Reaction takes place in sun?	b) Nuclear fission		a) Nuclear fusion	a) Nuclear fusion	a) Nuclear frision	a) Nuclear fusion	a) Nuclear	1	1
LITY)	13. Atoms of different chemical elements that have the same number of nucleons	d) Isotopes	a) Isobars	a) Isobars	a) Isobars		a) Nuclea	b) kotones	a) sobars	
SSTAB	12. Combin ing of two light nuclei of low mass to produce a heavy nucleus is	a) Nuclear fusion	ear	a) Nuclear fusion		ear	ea	a) Nuclear	ear	ear
S AND IT	11. Who invented nuclear fission?	b) Hans Bethe	c) Otto Hahn	c) Otto Hahn	c) Otto Hahn		c) Otto	c) Otto		
<b>JCLEUS</b>	10. Fission fragment are radioacti ve. What is the most	b) neutron	b) neutron	c) B*	) B+	b) c	0 B+	b) c	e 6	
UNIT TEST - B.SC-3, (SEM-5), (NUCLEUS AND ITS STABILITY)	9. What is the purpose with the moderator water in a fission reactor?	d) transports the released energy out from the reactor.	b) slows down neutrons to thermal velocities	b) slows down neutrons to thermal velocities	d) transports the released energy out from the reactor.	b) slows down neutrons to thermal velocities	b) slows down neutrons to thermal velocities	b) slows down neutrons to thermal velocities	b) slows down neutrons to thermal velocities	0
C-3,(SI	8. When an α- particle is emitted by an unstable nucleus	D) Decreas fed by 2	C) Increase d by 4	D) Decreas ed by 2	A) Increase d by 1	C) Increase d by 4		D) Decreas t ed by 2	D) Correas the ed by 2 v	b) slows down D) neutrons to Decreas thermal
ST - B.S	7. When 8. When an	A) increased by 1	A) Increased by 1	A) Increased by 1	A) Increased by 1	A) Increased by 1	C) D) Doesn't Increase change d by 4	A) Increased by 1	A) Increased by 1	A) Increased
UNIT TE	6. When a nucleus is divided into its constituents . energy is:	C) transformed into visible light.	D) absorbed by the nucleus which then breaks it apart.	D) absorbed by the nucleus which then breaks it apart.	D) absorbed by the nucleus which then breaks it	D) absorbed by the nucleus which then breaks it lapton appart.	D) absorbed by the nucleus which then breaks it D	D) absorbed by the nucleus which then breaks it last apart.	D) absorbed by the nucleus which then breaks it is apart.	C) transformed A s into visible Ir
	5. Binding energy is:		A) the amount of energy required to break a nucleus apart into protons and neutrons.	C) the amount of energy required to break a nucleus apart into electrons and neutrons.	B) the amount of energy required to break a nucleus apart into protons and electrons.	10		U)	vo	rgy
	4. Isotopes of an element:	A) have the same number (B) the amount of energy to forchors and electrons, required to break a but a different number of nucleas epart into proton neutrons.	A) have the same number of protons and electrons, but a different number of neutrons.	A) have the same number (c) the amount of energy of protons and electrons, required to break a but a different number of nucleus spart into neutrons.	B) have the same number of protons and neutrons, but a different number of electrons.	A) have the same number A) the amount of energy of protons and electrons, required to break a but a different number of nucles spent into proton meturons.  I and neutrons.	A) have the same number (A) the amount of energy of protons and electrons, required to break a but a different number of nucleus apart into protons neutrons.	A) have the same number (A) the amount of energy of protons and electrons, required to break a but a different number of nucleus spart into proton neutrons.	A) have the same number (A) the amount of energy of protons and electrons, required to break a but a different number of mucles apair fait person relutions.  And neurons.  And neurons.	A) have the same number A) the amount of energy of protons and electrons, required to break a but a different number of nucleas part into protons and neutrons.
	1, The of the month of the mumber is scored quivalent for the 3. Which of the the mumber of following about the following in the force is true?	C) It is much weaker than the electromagnetic force.	D) It is a strong, short-range, attractive force between the nucleons.	D) It is a strong, short-range, attractive force between the nucleons.	B) It is an attractive force between electrons and neutrons in an atom.	D) It is a strong, short-range, attractive force between the nucleons.	B) It is an attractive force between electrons and neutrons in an atom.	D) It is a strong, short-range, attractive force between the nucleons.	trong, e, force	A) It is an attractive force between electrons and protons in an atom.
2 Which	of the following is correct for the number o neutrons in the	B) N = Z - A	A) N = A	C) N = Z	C) N = Z	A) N = A	D) N = Z	A) N = A	A) N = A	A) N = A
	1. The atomic number is equivalent to which of the following?	B) The number of protons in the atom.	B) The number of protons in the atom.	B) The number of protons in the atom.		B) The number of protons in /	B) The number of protons in the atom. [	B) The number of protons in /	B) The number of protons in A	B) The number of protons in A the atom.
	Full Name of Students		war	Kanchan Kisan Nanne	Karishma Ratshma tatyaji ghugal t	Sunii	E Madhukar p Pise	0		~ \
No. and No.	Timestamp	Priya Ramesh 3/4/2021 10:32:14 (Shorude	Gauri Ist 3/4/2021 10:48:03 Sharma	Kanchan 3/4/2021 11:08:08 Kisan Nanne	B) The number of number of Karlshma protons in 3/4/2021 15:28:36 taryaji ghuqal the atom.	Saksh 3/4/2021 20:54:46 Awari	7 N 3/5/2021 13:12:58	Saumya Chandrash 3/5/2021 14:44:39 khar Singh	Neha Pramod 3/5/2021 19:10:18 Ambeka	GAYATF KESHAN 3/5/2021 20:53:01 BONDE

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(attentials with high					90		terials n high ss		m		low
			-	_			_	1	in .	_	c) Material Spontane with low ous atomic
82	otopes fi			(a) N	tones fus		(a)	a) (a)		-	c) Sponta ous
neo	io o										
William Control		a) Nucl				fusion	a) Nuclea		fusion	b) Nuclear fission	c) Spontaneo
b) Hans	Bethe	c) Otto Hahn	b) Hans Bethe	b) Hans	d) Marie	Curi	c) Otto		Jahn	Otto	b) Hans
		c) B+	c) B+	5	, (q	neutron					
b) slows down neutrons to thermal	b) slows down neutrons to	velocities b) slows down	thermal	c) absorbs the released	ot o		) slows own eufrons to nermal	W D	eactor.	nergy d	B) c) absorbs Increased Increased Increased b)
B) Increas	6	ed by 2	Increase d by 2	C) Increase d bv 4	D) Decreas		D) Decreas	O) (	d by 2	d by 2 e	B) c)
A) Increased bv 1	) Doesn'	change A)	Increased by 1	B) Increased by 2	)) Doesn't	nange	) Icreased	pese			B) B Increased Ir
C) transformed into visible light.	D) absorbed by the nucleus which then	ipart. ;)	nto visible ght.	-	D e	1	9 5	TI			transformed B) sinto visible Incre
	A) the amount of energy required to break.a nucleus apart into protons	and neutrons.  B) the amount of ener	nucleus apart into pro and electrons.	B) the amount of energy required to break a nucleus apart into protons d and electrons.	rrgy		10		) the amount of energy C) quired to break a tran-	nd electrons. ligh	weaker than the of protons and memorary from amount of energy (C) relectromagnetic but a different number of mucleus apart into protons into visible force.
	A) have the same number of protons and electrons, but a different number of	A) have the same number of protons and electrons.	neutrons.	C) have the same number of neutrons and protons.	A) have the same number of protons and electrons, but a different number of neutrons.		A) have the same number of protons and electrons, rout a different number of releutrons.	3) have the same number of protons and neutrons. Divide different number of prefections.	) have the same number B f protons and electrons, re	ar eutrons.	weaker than the of protons and electrons, electromagnetic but a different number of nu force.
between electrons and protons in an atom.		B) It is an attractive force between electrons and peutrons in an	atom. A) It is an attractive force	between electrons and protons in an atom.	D) It is a strong, short-range, attractive force between the nucleons.		0 0	trong. le. force	ive force en ons and ns in an	this much A	weaker than the o electromagnetic b force.
A B N = Z	× × × × × × × × × × × × × × × × × × ×		4	A) N = A	A) N = A		A = N 2			-	
D) None of the above	B) The number of protons in he atom.	3) The number of rotons in	he atom.	umber of rotons in se atom.	The Imber of otons in or atom.		The mber of stons in A stom.		The nber of lons in atom		D) None of C) N = Z the above +A
Karan Ratan Durge	njey			ar a			B) nu nu kita Kailas pri kude the	<b>5.</b>	-		laji.
3/5/2021 23:14:20	3/6/2021 8:26:59		3/6/2021 9:52:18	3/6/2021 17:37:58 g	A 3/6/2021 20:19:37 H		3/6/2021 23:08:40 Ur	Vr 3772021 18:10:44 To	Bh 981 3772021 18:14:35 bav	d	3/10/2021 11:45:58 There
	A) have the same number (b) the amount of energy (c) and offer off	State of Share of Share the same number (Share of Share o	Ration Di None of B) N = 2 - protone na descritoria and descritoria. Required forces a front of the above of B) N = 2 - protone na descritoria and descritoria and descritoria. A) have the same number of human reductions in A) N = A protone and descritoria. A) have the same number of nu	Ration D None of 5) N = 2 protocols and electrons, required to beside a protocols and electrons and of other pages.  By The above A atom.  By The above B	Ration Di None of B) N = 2 - protocos in an entrope. To be a socio de control	Frain Di None of Di N = 2 protons in an electrons. Required to the amount of energy Control of the atom of the atom of Di N = 2 protons in an electrons. Required to the atom of Di N = 2 protons in an electrons. Required to the atom of Di N = 2 protons in an electrons. Required to the atom of Di N = 2 protons in an electrons and electrons. Required to the atom of Di N = 2 protons in an electrons. Required to the atom of Di N = 2 protons in an electrons. Required to the atom of Di N = 2 protons in an electrons and electrons. Required to the atom of Di N = 2 protons in an electrons and electrons. Required to the atom of Di N = 2 protons in an electrons and electrons. Required to the atom of Di N = 2 protons in an electrons and electrons. Required to the atom of Di N = 2 protons in an electrons and electrons. Required to the atom of Di A = 2 protons in an electrons and electrons. Required to the atom of Di A = 2 protons in an electrons and electrons. Required to the atom of Di N = 2 protons in an electrons and electrons. Required to the atom of Di R = 2 protons in an electrons and electrons. Required to the atom of Di R = 2 protons in an electrons and electrons. Required to the atom of Di R = 2 protons in an electrons and electrons. Required to the atom of Di R = 2 protons in an electrons and electrons. Required to the atom of Di R = 2 protons in an electrons and electrons. Required to the atom of Di R = 2 protons in an electrons and electrons. Required to the atom of Di R = 2 protons in an electrons and electrons. Required to the atom of Di R = 2 protons in an electrons and electrons. Required to the atom of Di R = 2 protons in an electrons and electrons. Required to the atom of Di R = 2 protons in an electrons and electrons. Required to the atom of Di R = 2 protons in an electrons and electrons. Required to the atom of Di R = 2 protons in an electrons and electrons. Required to the atom of Di R = 2 protons in an electrons and electrons. Required to the atom of Di R = 2 protons in an electrons and electrons. Requi	State of Sta	Districtive of Sin a 2 between the protocol and the atoms number of the atoms of the atoms of Sin and Sin an	Place   Discose   Control of Discose   Disco	n family bloom of [3] N-z – sections in an included by the amount of energy C). But shows a few control in the same number (5) the amount of energy C) included by the amount	in Figure 10 Notice of 10 N = 2 - controvant in a manufacture for the same number of 10 the amount of eventy of 10 the same number of 10 the amount of eventy of 10 the same number of 10 the amount of eventy of 10 the same number of 10 the sam

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(Mr. K. V. Bhowpole)

Nilkantheau Shinde Sei And Ast's college Bhadsawati (2020-21) & Unit Test BSC - I (Sexo-I) Sub: Mathematics (Paper -II) Mark-20 Time: 30 min Quel. Prove that it limit of a tunation tox, y) (n,y) -> (no,yo) exists, then it is unique Que 2. Prove that lim (3x-2y) = 14 Que3. If U=F(x-y, y-z, z-x) then Prove that  $\frac{84}{8x} + \frac{84}{84} + \frac{84}{8z} = 0$ QUE 4. It a real valued function & m, y) is continuous at P. (2001yo) then there is a neighbourhood of Po in which fory) to 6 bounded @ Unbounded QUES It Z= + (my) has a differential dz= a dx + b dy at (xo, go) thegradue 5x (xo1 to) , by (xo, to) enisto. L. stadke
PRINCIPAL

A.S. Science & April College

a) is unique

(b) es enterawati, Dist-Chandrapur @ is unique (A) Mone of these

### Nilkanthrao Shinde Science and Arts College, Bhadrawati dist. Chandrapur B.Sc.-I (sem-I)

# UNIT TEST Result

# Paper-II(Differential Calculus and Trigonometry)

## Sub:-MATHEMATICS

Sr.no	Name of Students	Marks(20)
1	Gitika Durge	14
2	Divya v. vyvahare	13
3	prasad s vidhate	11
4	Atik B. Kshhirsagar	11
5	Supriya v. Balpane	13
6	Pallavi S. Gowardipe	12
7	Priti G. yelane	15
8	Saniya s. sheikh	16
9	Darshana D. Nagore	15
10	Achal R. Jidgalwar	16
11	Prachi S. Bagde	12
12	Prachi H. Kakde	14
13	Dipali S. Boinwar	13
14	Alishanaaz A. sheikh	16
15	Janvi P. Nimbalkar	15
16	Tumeshwari S Jiwtode	12
17	Kunal S. donge	14
18	Achal a. Awale	16
19	Sahil V. Nakshine	15
20	Khushi m.Kharwade	17
21	Saifaddin sayyad	15
22	Krupali A. Tonge	16
23	Komal D. Shivarkar	14
24	Ajay C. doppala	12
25	Priti V. Karekar	13
26	Vibha s. Tated	18
27	Hitashi M. Thegane	15
28	Tejas d. Telang	14
29	Ashwariya R. Tiwari	13
30	Komal c. thakare	14
31	Aishwariy R. Pijdurkar	15
32	Prajot N. Tangapendi	12
33	Anushka R. Jawade	18
34	Nikita V. Jambhulkar	16
35	Arti S. Kakade	14
36	Pranjali G. Raipure	13
37	Iramnaj S. Ali	12
38	Mahesh C. bawane	14
39	Prajakta N. bodhane	13
40	Mahek M. Sayyad	14
41	Mayuri D.Bodhe	15
42	Rahul S. Masalkar	12
43	Pratik J. Daogaokar	13
44	Rajeshwari S. Sundargiri	12

45	Sairamshi Y. Allakatla	10
46	Keshav anil Kumar	10
47	Nikanksha P. Patil	12
48	Hitesh A. Dahule	13
49	Vipul . Nande	15
50	Pallavi S. Kale	13
51	Rajeshri S. Dixit	12
52	Swati S. Kamble	12
53	Dipak R. Soyam	11

Sladke

Nilkantheau shinde Sci. and Aet's college Bhadrawati (2020-21) Unit Test Bsc-I (sem-II) Sub &-Mathemertics (paper-I) Time: 30 min Quel: - solve any two. A solve (1-x2)(1-y)dx = xy(1-y)dy (B) Find the ofthogonal trajectory of En = an cosno. @ solve (x2+y2+2x)dx + 2ydy =0 Que 2:- $A = \frac{d^2y}{dx^2} + x^2 \frac{dy}{dx} - y \sin x = 0$  is 1 Partial differential Equation 6) Ordinary differential equation. @ None of these 1 5 dy + 22 + 342 =0 is 1 Homogeneous DE (b) Non homogeneous DE N.S. Science & Arts College @ Mone of these Bhadrawati, Dist-Chandrapur

3(2)

0	It un	4
1	adx+No	20

y) is an integrating factor of y, then Il satisfies the pastial DE

1) Nex+ Muy - M(Nx+My) =0

6 N Mx - My - M(Nx-My) =0

② NAIX + MALY - M(NX+My) ≠0 → ②

If the DE Mdx + Ndy =0 is homogeneous and

Mx+Ny \$0 then integrating factor is

@ \frac{1}{N\chi.My} \overline{D} \frac{1}{M\chi.Ny} \overline{O} \frac{1}{M\chi.Hy} \overline{\text{0}} \frac{1}{M\chi.Hy} -

€ If P=In (px-y) its Primitive es

 $\hat{\partial} y = (x - e^{c})$   $\hat{\partial} y = cx^{2} - e^{-c}$ 

@ y = - cx - ec @ y = - cx2+ec \_ 0

(F) Bernoulli's equation is

@ Y'+Py =Q

6 y'+Py=y o

O · y' + py = 0

None of these \_\_\_\_

# Nilkanthrao Shinde Science and Arts College, Bhadrawati dist. Chandrapur B.Sc.-I (sem-II)

# UNIT TEST Result

# Paper-I(Ordinary DE and Difference Equations)

### Sub:-MATHEMATICS

r.no	Name of Students	Marks(20)
1	Achal R. Jidgalwar	14
2	Aishwariy R. Pijdurkar	13
3	Ajay C. Doppala	12
4	Alishanaaz A. sheikh	12
5	Anushka R. Jawade	18
6	Arti S. Kakade	12
7	Ashwariya R. Tiwari	15
8	Atik B. Kshhirsagar	16
9	Darshana D. Nagpure	15
10	Dipak R. Soyam	16
11	Dipali S. Boinwar	12
12	Divya v. vyvahare	14
13	Gitika Durge	13
14	Hitashi M. Thegane	16
15	Hitesh A. Dahule	15
16	Janvi P. Nimbalkar	12
17	Keshav anil Kumar	14
18	Khushi m.Kharwade	18
19	Komal c. Thakur	15
20	Komal D. Shivarkar	11
21	Krupali A. Tonge	15
22	Kunal S. donge	16
23	Mahek M. Sayyad	14
24	Mayuri D.Bodhe	12
25	Nikanksha P. Patil	13
26	Pallavi S. Gowardipe	14
27	Pallavi S. Kale	15
28	Prachi H. Kakde	14
29	Prachi S. Bagde	13
30	Prajakta N. bodhane	14
31	Pranjali G. Raipure	15
32	prasad S. vidhate	12
33	Priti G. yelane	14
34	Priti V. Karekar	16
35	Priya R. Ambilkar	14
36	Rahul S. Masalkar	13

37	Rajeshri S. Dixit	12
38	Rajeshwari S. Sundargiri	14
39	Sahil V. Nakshine	13
40	Saifaddin sayyad	16
41	Saniya s. sheikh	15
42	Supriya v. Balpane	12
43	Tejas d. Telang	13
44	Tumeshwari S Jiwtode	12
45	Vibha s. Tated	18
46	Vipul . Nande	12
47	Sameer R. Yadav	12
48	Rachi Katker	13
49	Krunali A Ghugul	15
50	Keshav anil Kumar	12

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Milkanthead shindle Sci. and Het's college Bhadrawati (2020-21) UNIT TEST BSC-II (Sem-III) Sub: - Mathematics (paper - I) mark - 20 Time 30 min Que 1:- A sequence can have at most one limit Or it lim on exists, it must be unique Que-2:- Find the limit of the sequence (5n)  $S_n = \frac{1}{n^2+1} + \frac{1}{n^2+2} + \cdots + \frac{1}{n^2+n}$ Que 3: - Show that the sequence of decimal digits is convergent. Que 4: - Every convergent sequence of real number es @ cauchys sequence (b) divergent sequence @ unbounded sequence 1 convergent sequence QUE 5:- It I'm Sn = +00, the sequence LSn'y i @ consegent 6 divergent Stadke a None of above N.S. Science & Arts College

Bhadrawati, Dist-Chandrapur

QUE 6: If monotone sequence of real numbers is bounded its and only it a divergent.

(a) Unique (b) divergent

(c) convergent (d) None of these

(d) \land \frac{1}{1} is a bounded sequence

(d) \land \frac{1}{1} is a cauchy sequence

(e) \land \frac{1}{1} is a cauchy sequence

(f) \land \frac{1}{1} is talse

(f) \land \frac{1}{1} is talse

(g) \land \frac{1}{1} is talse

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#### Nilkanthrao Shinde Science and Arts College, Bhadrawati dist. Chandrapur B.Sc.-II (sem-III)

#### **UNIT TEST RESULT**

Paper-I(Real Analysis) Sub:-MATHEMATICS

Sr.no	Name of Students	Marks(20)
1	Hemant M. Farkade	14
2	Tushar D. Asutkar	13
3	Sidhanti Y. Sonewane	13
4	Pooja D. Modot	15
5	Samir A. Nasnurkar	13
6	Nandak R. Kamde	12
7	Pranjal M. Bobade	15
8	Surai G. Bobade	16
9	Prem S. Ramteke	15
10	Amrita K. Prasad	16
11	Priya U. Shah	12
12	Gayatri B. Donge	14
13	Shruti B. Raipure	13
14	Shreya S. Junarkar	16
15	Prachi v. Patil	15
16	Ritul S. Sharma	12
17	Madhuri P. Kandale	14
18	Sejal S. Petkar	16
19	Kajal A. Varbhe	15
20	sejal S. Deogade	15
21	Shruti M. Ramteke	15
22	Bhavana W. Thakare	16
23	Chirag M. Patil	14
24	shruti A. Deogade	12
25	Dipti V. Bhajankar	13
26	Aman S. Patil	16
27	Payal S. Bavane	15
28	Ritu D. Ingade	14
29	Sapna M. Mandal	13
30	Kiran D. Satpute	14
31	Aniket P. Walke	15
32	Amisha S. Kakde	12
33	Dhanashree k nagrade	14
34	Nikita v kurai	16
35	Vaishnvi M Dhambare	14
36	dhanshree r chende	13
37	snehal s sukhdeve	12
38	vaishnvi p dhage	14
39	Kiran D. Ghosare	13
40	Smita N. Wadhai	14
41	Neha S. Manthanwar	15
42	Mrunali M. Donge	12
43	Sejal R. Nandurkar	13
44	Achal D. Doye	12

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45	Prajakta M. Gedam	15
46	Prashant B. Narwade	14
47	Santosh R. More	12
48	Sushma P. Kokude	13

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Nilkanthrao shindle Sci and Alt's college, Bhadlawati (2020-21) Unit Test BSC-II (Sem-IV)

Sub: Mathematics (paper-1)

Time 30 min

mark-20

Que1: - solve any two

- A Group definations
- B Invesse element definations
- O Identity element definations
- Que 2: shone that the binary operation, defined on R by a.b = a+b+ab is commutative and associative.
- Que 3:- It G is a group such that  $(ab)^2 = a^2b^2$ , if  $a_1b \in G$  show that G must be abelian
- Que 4!- A group & is finite & it o(9) is @ Infinite & Finite @ commutative @ xone of these
- Que5!- The identity of a group & is @ Unique & None of these

Que 6: - It a nonempty subset H of the group G is a subgroup of G it and only et.

@ abe a → able a @ abe H → able #H

@ a, b ∈ H ⇒ a 1 b + ∈ H

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#### Nilkanthrao Shinde Science and Arts College, Bhadrawati dist. Chandrapur B.Sc.-II (sem-IV)

#### UNIT TEST Result

#### Paper-I(Algebra) Sub:-MATHEMATICS

Sr.no	Name of Students	Marks(20)
1	Achal D. Doye	14
2	Amisha S. Kakde	13
3	Amrita K. Prasad	14
4	Aniket P. Walke	15
5	Bhavana W. Thakare	13
6	Chirag M. Patil	12
7	Dhanashree k Nagrade	15
8	Dhanshree R Chende	16
9	Gayatri B. Donge	15
10	Hemant M. Farkade	16
11	Kajal A. Warbhe	12
12	Kiran D. Ghosare	14
14	Kiran D. Satpute Madhuri P. Khandale	13
15	Nandak R. Kamle	16
16	Nikita v Kurai	15
17	Pooja D. Modot	14
18	Prachi v. Patil	16
19	Prajakta M. Gedam	15
20	Pranjal M. Bobade	15
21	Prashant B. Narwade	15
22	Prem S. Ramteke	
23	Priya U. Shah	16
24		14
25	Ritul S. Sharma	12
	Samir A. Nasnurkar	13
26	Santosh R. More	16
27	Sapna M. Mandal	15
28	Sejal R. Nandurkar	14
29	sejal S. Deogade	13
30	Sejal S. Petkar	14
31	Shreya S. Junarkar	15
32	shruti A. Deogade	12
33	Shruti B. Raipure	15
34	Shruti M. Ramteke	16
35	Sidhanti Y. Sonewane	14
36	Smita N. Wadhai	13
37	Snehal S. Sukhadeve	12
38	Suraj G. Bobade	14
39	Sushma P. Kokude	
40	Tushar D. Asutkar	13
_		14
41	Vaishnvi P Dhage	15

42	Runail S Amate	12
43	Palash D. Ghate	13
44	Pratiksha D Parchake	13
45	Payal S. Bavane	15
46	Ritu D. Ingade	14
47	Vaishnvi M Dhambare	13
48	Aman S. Patil	15

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Wilkantheao Shindle Sa and Art's college Bhadlawati (2020-21) Unit Test BSC-III (Sem-V) Sub: - Mathematics (paper-1) mark-20 Time 30 min Quel: In (2 define, for u= (d, 2) and  $\mathcal{N} = (\beta_1, \beta_2), (u, v) = 2 \alpha_1 \overline{\beta}_1 + \alpha_1 \overline{\beta}_2 + \alpha_2 \overline{\beta}_1 + \alpha_2 \overline{\beta}_2$ show that this defines an Inner product 2000 (2

Quez: Let V be a set of all continuous complex-valued functions on the closed interval [0,1], It &ct), get) &V, define (fu), gu) = stu). gu) dt Shoul that this defines an Inner product on V.

Que3: Let V be an Inner product space over F. In V detine the distance d(u,v) from 4 tore by d(u,v) = 1/4-vell. prove that (1) d(4,22) no and d(4,20) =0 => 4=2 @ dcure) = d(re,u)

Que 4: - Evely Osthogonal set is slauke

(a) Linearly Dependent

(b) Linearly Independent

(c) Linearly Space

(d) Linear Space

(e) Bhadrawati, Dist-Chandrapur

Que: 5 select the correct answer

@ WNW+=1

B WNW+= {00}

0 W10 W1 = {0}

a) M W N W = {0,1}

Que 6: Let V be an Inner product space then u, v ∈ V and X ∈ F

@ 11 x u11 = 1x / 11 u1/

B 11 4+1211 < 141+121

O 11411 =0 ♦ 4+0

@ None of these

Que 7: It u and ve be vectors in an Inner product space such that 114+111 = 8, 114-1211 = 6 and 11411=7 then the value of 11411 \$

@ 551

(b) 1

052

@ 2

#### Nilkanthrao Shinde Science and Arts College, Bhadrawati dist. Chandrapur B.Sc.-III (sem-V) UNIT TEST Result

## Paper-I(Linear Algebra) Sub:-MATHEMATICS

Sr.no	Name of Students	Marks(20)
1	Tejaswini M. Pise	15
2	Priyanka N. Lingayat	14
3	Hemlata B. Gathe	16
4	Kanchan K. Nanne	15
5	Jayesh J. Bodhe	15
6	Saurabh S. Vidhate	16
7	Shivam S. Saxena	16
8	Prajwal V.Pandhare	14
9	Pallavi V. Khamankar	15
10	Gayatri K. Bonde	16
11	Somya C. Singh	19
12	Rahul A. Shah	14
13	Gandhi M. Pothagani	13
14	Sakshi D. Masirkar	16
15	Priyanka S. Tarale	15
16	Atul L. Mohurle	15
17	Jai S. Pandhare	14
18	Ankita K. Urkude	16
19	Pornima S. There	15
20	Shivani Y. Jogi	16
21	Neha P. Ambekar	18
22	Prajkta R. Hatzade	17
23	Akshay M. Bhusari	14
24	Sakshi S. Awari	14
25	Pritija A. Kodape	13
26	Samiksha S. Pawar	16
27	Karishma T. Ghugal	15
28	Priya R. Ghorude	14
29	Pooja A. Pal	13
30	Sahil P. Deogadde	14
31	Kunal P. Shindew	15
32	Sarika S. Nagpure	16
33	Tanmay A. Kolhe	16
34	Prashant S. Gupta	16
35	Nikhil B. Atram	14
36	Subodh S. Raut	13
37	Jayshree R. Durve	15
38	Nishas Mendhe	14
39	Ankita R. Jawade	16
40	Ashitosh K. Prasad	14
41	Pritam S. Harshe	15
42	Gauri I. Sharma	15
43	Akshay A. Aaut	13
44	Harshal B. Khangar	12

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45	Deepak V. Wailkar	15
46	Jitesh J. Mujumdar	16
47	Karan R. Durge	14
48	Chetan S. Datey	16
49	Bhagyashree Y. Parkhi	15
50	Pranali R. Thak	16
51	Shivani S. Ramteke	14
52	Shubhangi N. Dhavale	16
53	Pranay B. Zalwade	15
54	Bhushan G. Bavane	14
55	Pankaj P. Khilare	17
56	Gurudeo P. Kumbhare	16
57	Vishwjit V. Tonge	15
58	Shobit P. Bura	14
59	Simran S. Sheikh	16
60	Chitalee N. Darvankar	15
61	Yogita S. Bahure	16
62	Amit D. Sapat	15
63	Pallavi R. Deharkar	14
64	Sanisagar A. Nikhade	16
65	Suraj B. Mankar	16
66	Esha V. Choudhary	14
67	Vaishali M. Khade	15
68	Pallavi M. Adate	16
69	Ranjana P. Paswan	15
70	Vanshree S. Kamble	15
71	Komal C. Ranvir	14
72	Pronoti C. Salame	15
73	Priti G. Punvatkar	16
74	Shital T. Shiwarkar	14
75	Samiksha B. Dongare	16
76	Ganesh U. Masharkar	13
77	Durgeswari P. Mallelwar	15
78	Aman R. Rajbhar	14
79	Pallavi V. Vidhate	16

NELKantheau Shinde Sci. and Arts college Bhadeawati (2020-21) Moit Test BSC-III (Sem-VI) Sub:- Mathematics (paper-1) mask-20 Time 3000io Que 1: The decivative of any constant vector function is a) one @ Zelo @ None Quez:- It t and of are two differentiable vector functions of & then @ of (t+g) = of g. + t. of Que3:- It of tinds =0, t is called @ I 2 20 tational (b) Estational Solenoidal flade None of these

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Que 4:- A vector field t is conservative 6 A VX \$ =0 0 VX 5 =1 -12 Que 5: If v = 60x 2. Prove that  $w = \frac{1}{2} \text{ cuelv}$ , where we is a constant Que 6: - If t= +3; + + +2; - ++k find the value of St. de teom (0,0,0) to (2,4,8) an' along the path C: x=t, y=t2, z=t3 QUET:- Evaluate the surface integral SS (yzi+ zxj+xyk).ds, where s is the sustaine of the sphere x2+y2+22=a2 in the tiest octant.

> V.S. Science & Arts College Madrawatt, Dist-Chandrapur

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#### Nilkanthrao Shinde Science and Arts College, Bhadrawati dist. Chandrapur B.Sc.-III (sem-VI)

#### UNIT TEST Result

#### Paper-I(Complex Analysis and Vector Calculus) Sub:-MATHEMATICS

Sr.no	Name of Students	Marks(20)	
1	Akshay A. Raut	14	
2	Akshay M. Bhusari	13	
3	Aman R. Rajbhar	15	
4	Amit D. Sapat	14	
5	Ankita K. Urkude	13	
6	Ankita R. Jawade	14	
7	Ashitosh K. Prasad	15	
8	Atul L. Mohurle	16	
9	Bhagyashree Y. Parkhi	15	
10	Bhushan G. Bavane	16	
11	Chetan S. Datey	12	
12	Chitalee N. Darvankar	14	
13	Dipak V. Wailkar	13	
14	Durgeswari P. Mallelwar	16	
15	Isha V. Choudhary	15	
16	Gandhi M. Pothagani	12	
17	Ganesh U. Masharkar	14	
18	Gauri I. Sharma	16	
19	Gayatri K. Bonde	15	
20	Gurudeo P. Kumbhare	16	
21	Harshal B. Khangar	15	
22	Hemlata B. Gathe	16	
23	Jai S. Pandhare	14	
24	Jayesh J. Bodhe	12	
25	Jayshree R. Durve	13	
26	Jitesh J. Mujumdar	18	
27	Karan R. Durge	15	
28	Karishma T. Ghugal	14	
29	Komal C. Ranvir	13	
30	Kunal P. Shinde	14	
31	Neha P. Ambekar	19	
32	Nikhil B. Atram	12	
33	Nishas Mendhe	18	
34	Pallavi M. Adate	16	
35	Pallavi R. Deharkar	14	
36	Pallavi V. Khamankar	13	
37	Pankaj P. Khilare	16	
38	Pooja A. Pal	14	
39	Pornima S. There	13	
40	Prajkta R. Hatzade	14	
41	Prajwal V.Pandhare	15	

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42	Pranali R. Thak	12
43	Pranay B. Zalwade	13
44	Prashant S. Gupta	12
45	Pritam S. Harshe	14
46	Priti G. Punwatkar	15
47	Pritija A. Kodape	12
48	Priya R. Ghorude	13
49	Priyanka S. Tarale	15
50	Pronoti C. Salame	13
51	Rahul A. Shah	12
52	Ranjana P. Paswan	14
53	Sakshi D. Masirkar	16
54	Sakshi S. Awari	15
55	Samiksha S. Pawar	14
56	Shanisagar A. Nikhade	13
57	Sarika S. Nagpure	15
58	Saurabh S. Vidhate	14
59	Shital T. Shiwarkar	16
60	Shivam S. Saxena	16
61	Shivani S. Ramteke	14
62	Shivani Y. Jogi	14
63	Shobit P. Bura	15
64	Shubhangi N. Dhavale	16
65	Somya C. Singh	20
56	Subodh S. Raut	14
57	Suraj B. Mankar	13
58	Tanmay A. Kolhe	14
59	Vaishali M. Khade	15
70	Vishwjit V. Tonge	16
71	Yogita S. Bahure	15

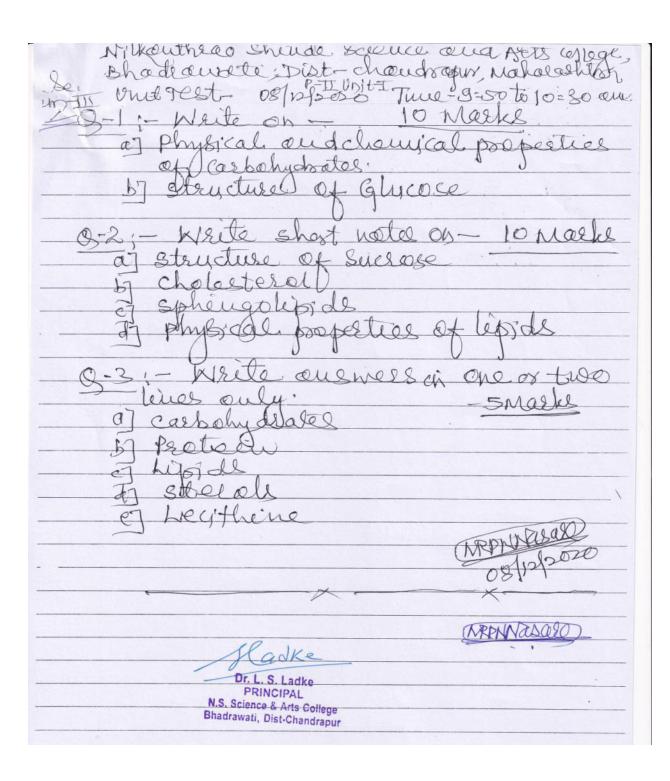
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Name of students	Distained mark	Totalang
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(1) Marshall Asker	15	115
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10 Midhi A Balki	10 15
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-14 FOR 18 12 20 10 10 10 10 10 10 10 10 10 10 10 10 10		PERSONAL PROPERTY AND ADDRESS OF THE PERSON NAMED IN COLUMN 1
02 Crive the economic	importance	of Algae 21
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Name of Students	obtained marks	marks.
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college shadrawate; Dist-chaudrager college Unit rest - True-3-50-10-3000 Blockers
3/3/ Unit 1880- 11me-9=50-10=30 cm 15 80 00 mil
2008-1:- Write on - 10 Marks
by mRNA
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a) SHRNA
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57 Types Of RNA
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1. Likitice siescilla
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Dr. L. S. Ladke PRINCIPAL
N.S. Science & Arts College Bhadrawati, Dist-Chandrapur

hikantheau Sheinde Serence and Acts college schadewali
Test, Until sale 01/2/2021 outnie Exam
Topic - Karyonga
3-1:- Write on- 10 marks
a] Aneuploidy.
5] Inversion
0-2: Write shot notes on - 10 marks - 2/2 Maskeach
a) Aldetron or Departmenty
Duplication complex
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9 Nuclius 1 1 as the soline
0-3: - Write ourvers is two or three lines only. 5 Marks a mark each
eug.
a] Karyotybe
by Harray
I Diploid sladke
Or.L. S. Ladke PRINCIPAL N.S. Science & Arts College  Company of the Description of the Chandrapur
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Nilkenstroad sheuda Science and
rest P-II Unid-II Sew-III 12/10/2000
rest is se sem-III 12/ppose
Time - 10=90 to 11=20 cm.
Q-1:- Write on - 10 Marks
a) IUB system of Nomenciature
ef Enzymes.
b) Characteristics of the
Enzyhire.
S-2: - Write short noteson-19
a) Holoepzywe Mask
b) Apoensyme
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- or three level only 5 mark
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PRINCIPAL
N.S. Science & Arts College Shadrawati, Dist-Chandrapur

Nilkenthead sheirde service and Agets college; shadrawate; Dist-chandrapue.  Pet - 07/01/2021 PIL Time - g. 20 am - 9.4° am
Topic - Ascent of Sop
3ue-1: Weite on 10 Marks
a] Ascent of sap b] Root Pressure Theory.
Que-2; - Write short notes on 10 Marks of of water conduction through xylent; washed
6] cohesion Adves
on exemple for Ascent of sap.
Que-3: - Write onswers in two or three 5 Marks
VALUE - I
a) Transportation. b) xylen
a Ascent of sap
e) phloengraske aronnaise
N.S. Science & Arts College Bhadrawati, Dist-Chandrapur

Nilkentheau Shinde Science and Arts collège,
Bhadrawati; Dist-chaudrapm-442,982
Voit rest - Max Marks 25
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17 Note - Click in the At LOLD one level Dill
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9-3: Write answer in two or three lines only. 5 marks (I markegen)
a) Photosyntheeir
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PRINCIPAL
N.S. Science & Arts College Bhadrawati, Dist-Chandrapur

Nilkantheau Sheinde Science and Arts College Bhadrow Dist-chandrague, Maharashtia State	
BSe Part - I Sem-II BOTANY online Exa	L
Papels-II; Unit-I; Unit rest-I	
Question Papes - Mex. Morks - 25	
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albennettitalian Theory - 5 marks	
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a] Sahanianthus flower 2/2 each	
17 th sheastus	
b] Theopheastus	
c] Magnolia	
d] Liedotype	
Que-3: - Write enswers in two or three lenes	
ordy -5 merksp	
al Linnaeus ornarkeach	
5] Isotype	
1 Binomial Nomenclature (1974)	
e) Neotupe (Mannasare) (Do. P. N. Nasale)	
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Madke sommano	
Dr. L. S. Ladke PRINCIPAL	
N.S. Science & Arts College Bhadrawati, Dist-Chandrapur	

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t-chardeapur, Maharashtia State.	O au
Se. Part - I, Sem-II, BOTANY	o nline stat
Deb-II: Unit-I; Unit rest-I	
uestion Papes - Mex. Marks	-25
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	5 Marks
Que-2:- Weite short notes on-	10 Marks
a) Sahanianthus flower 2	beach
b) Theopheastus	
c] Magnolia	
d] Liedotype	
	or three lines
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D. DETWEET,
Papele-II; Unit-II; Unit Test-II
Question Papels. Max. Marks-25
10 Marks
Que-1:- Write on- 10 Marks
1> Vietual (Digital) Herbarium -5M
2) Types of classification of Anglospeans
Que-2: - Write short Doles OB - 2/2 Marks each
2) Types of classification of Angiospeems 5 M  Que-2: - Write short notes on - 40 Marks each  a) Merits of Bentham & Hooker System  a) Merits of Bentham and Hooker System
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e) Angrosperme (DE.P.N. Nasare)
Stacke ARAMAGASO
Or. L.S. Ladka
N.S. Science & Arts College Bhadrawati, Dist-Chandrapur

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Bho de awate; Disto-chandrapus.
B.S. Pasts-I Sem-II BOTANY Online Exam
Papele-II; Unit-III, UnitTest-III
Max. Marks-25
Question Papez Max. Marks-25
10.1.1.0
Que-1:- Write on - 10 Marks
a Regraductive characters of Brassicaceae  Di Morphological characters of Malvaceae
15 de la mient al astase de Malutacogo
b) Motphological characters of return and
348-2: - Write short notes on- 10 Marks
of Masphological characters at Boassicaceae
b) Reproductive characters of Malvaceae c) Reproductive characters of Fabaceae
7 - Liting charactors of Frebaroas
c) Reproductive characters of teclorage
of Moephological characters of solary coa
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Que-3: Write abswers is two or three lines
only.
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c) Capsule 1.1 recondition
c) Capsure Collabores condition
d) Moroaderphi Conity Solonaceae
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MANAGER
Dr. L. S. Ladke PRINCIPAL
N.S. Science & Arts College

Bhadrawati, Dist-Chandrapur

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Bhadrawale; Dist-chandrapus. Online Exom
Bre. Part-I, Sem-II BOTANY Churchen
Papele-II, Unit-IV, Unit-Test-IV
Question Paper - Mar. Marks-25
Que-1: - Write on: - 10 Marks
a) Reproductive characters of Ascleipiadaceae
1 Morphology of beteraceae
S Moepho
Que-2:= Write short notes on - 10 marks
a) Ray florest
c] Morphological characters of Pay toxel
c) Morphological characters of the productive characters of Poaceae
gue-3: Write answers is two or three lines 5 marks, only.
Sue-3: While abstrais of the Mark each.
a corner Gynostegium Olmark eger.
b) pappus
c] Palea obligaçõe
d) Inflorescence of libraceae
e Economic insportance of Po accompanyasaso
sladke (sonwoods)
Dr. L. S. Ladke PRINCIPAL
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Bhadrawati, Dist-Chandrapur

## Nilkantharao Shinde Science and Arts College, Bhadrawati

Unit Test

Session- 2020-2021

Class- B.Sc-II Sem-III

Subject- Inorganic Chemistry (Unit-I)

Time-45 Minutes

Mark-12

Q-1 what are polyhalide? Give the classification of polyhalides.

(5Mark)

Or

What is Carbides? Discuss classification of carbides with suitable example.

Q-2 (Any Two)

- A) Explain Chain silicates with examples (2.5 marks)
- B) Explain basic nature of Iodine. (2.5 mark)
- C) Describe the structure of Borazine. (2.5 mark)

Q-3. Any Two

- 1. What are interhalogen compounds? (1 mark)
- 2. Draw the structure of borazine. (1 mark)
- 3. What are Silicates? (1 mark)

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## Nilkantharao Shinde Science and Arts College, Bhadrawati

Unit Test

Session- 2020-2021

Class- B.Sc-I Sem-I

Subject- Inorganic Chemistry (Unit-I)

Time-45 Minutes

Mark-12

Q-1. Define Ionization potential. Describe the factors affecting on it. Explain the Trend of Ionization potential in group and period. (5Mark)

OI

Define Quantum number. Discuss the orbital quantum number and magnetic quantum number in detail

Q-2 (Any Two)

- A) Define electronegativity. Why is electron affinity of fluorine iess than chlorine? (2.5 mark)
- B) State and explain Hunds rule of maximum multiplicity. (2.5 mark)
- C) Calculate the effective nuclear charge for 3d electron in Chromium.

(2.5 mark)

Q-3. Any Two

- 1. Define Careening Constant. (1 mark)
- 2. Al is good reducing agent Explain. (1 mark)
- 3. Write Schrodinger equation. (1 mark)

Dr. L. S. Ladke

N.S. Science & Arts College Bhadrawati, Dist-Chandrapur

#### Nilkantharao Shinde Science and Arts College, Bhadrawati

Unit Test

Session-2020-2021

Class- B.Sc-III Sem-IV

Subject- Inorganic Chemistry (Unit-I)

Time-45 Minutes

Mark-12

Q-1 What is error? Explain determinate and indeterminate error in detail. (5Mark)

Or

Explain principle and various types of interferences in Flame Photomerty.

Q-2 (Any Two)

- A) Explain i) Significant figure ii) Accuracy (2.5 marks)
- B) Explain F-Test and T-Test. (2.5 mark)
- C) Distinguished between Accuracy and Precision. (2.5 mark)
- Q-3. Any Two
- 1. Define i) Mean ii) Median (1 mark)
- 2. Calculate Significant figure i) 6.0213 ii) 0.00215 (1 mark)

3. Write two advantages of Flame Photometry (1 mark)

Bhadrawati, Dist-Chandrapur

## Nilkanthrao Shinde Science and Arts College, Bhadrawati Dist. Chandrapur

## Department of Chemistry

## **Unit Test**

Topic:- NMR Spectroscopy

M 20

a. An organic compound having molecular formula ${\rm C_3H_6O}$ shows following NMR dates	a
i) 3H – Triplet δ-1.5	
ii) 2H- Quartet δ-2.6	
iii) 1H- Singlet δ-7.2	
Deduce the structure.	5
b. Write brief note on-	
i) The Role of TMS in NMR Spectroscopy.	
ii) Equivalent and Non-equivalent protons.	5
c.What is chemical Shift ?	21/2
d. Define coupling contant 'J'.	21/2
e. What is shielding and deshielding of protons in NMR spectroscopy?	21/2
f. What is spin – spin coupling.	21/2

Lt. Sachin H. shrisame

# Nilkanthrao Shinde Science and Arts College, Bhadrawati Dist. Chandrapur Department of Chemistry

## Unit Test

Topic:- Organic Synthesis Via Enolate

M 20 Date 10/11/2020

Q.1 What are enolates? Give two examples with their structures.	5 M
Q.2 How is acetoacetic ester prepared by Claisen condensation? St	arting from acetoacetc
ester, how will you prepare the following?	
(I) 4-methyl uracil & (II) Ethyl methyl ketone	5M
Q.3 Give short account of Keto-Enol Tautomerism.	2.5 M
Q.4 How will you obtain following compounds from diethyl malona	ate
(I) n-valeric acid & (II) adipic acid	2.5 M
Q.5 How is diethyl malonate prepared from acetic acid	2.5 M
Q.6 Write nte on Acidity of α-hydrogen atom	2.5M

Lt. Sachin H. Shiname

## Nilkanthrao Shinde Science and Arts College, Bhadrawati Dist. Chandrapur

## Department of Chemistry

## **Unit Test**

Topic:- Green Chemistry

M 20

a. Define Green Chemistry. Discuss the main purpose of Green Chemistry.	5
b. Explain the alternative solvent or replacement of solvents in green chemistry.	5
c. What do you mean by cardle to cardle design?	21/2
d. Discuss scientific areas for practical applications of green chemistry.	21/2
e. Write short note on Reduction of solvent toxicity.	21/2
f. Write a note on Phase transfer catalysis.	21/2

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Bhadrawati, Dist-Chandrapur

Lt. Sachin H. Smrome

## Nilkanthrao Shinde Science & Arts college, Bhadrawati Department of Chemistry

## Unit Test(online) Popic :- Polymer

Marks 20 Date:- 17/08/2020

Q.1 How do you Classify polymer in various way?	(5M)
Q.2 What are cross-linking process? Give any two example.	(5M)
Q.3 What is phenol formaldehyde resin? Give it's properties and application	(2.5M)
Q.4 What is ruber? Explain neoprene & Buna-S with application.	(2.5M)
Q.5 What is vulcanization process?	(2.5M)
Q.6 What is conductiong polymer?	(2.5M)

Dr. L. S. Ladke
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N.S. Science & Arts College
Bhadrawati, Dist-Chandrapur

It. Sachin H. Shrivame

## Nilkanthrao Shinde Science & Arts college, Bhadrawati Department of Chemistry

Unit Test(online)
Popic :- Polymer

Marks 20

Date:- 17/08/2020

#### Test result

Sr.no	Name	Mark obtained	Out of	date	remark
1	Aramnaz Sheik	17	20	17/08/2020	
2	Atul Mohurle	14	20	17/08/2020	
3	diksha deogade	15	20	17/08/2020	
4	Durgeshwari Mallelwar	13	20	17/08/2020	
5	gayatri Bonde	18	20	17/08/2020	
6	Mrunal Dadmal	17	20	17/08/2020	
7	Neha Ambekar	19	20	17/08/2020	
8	Nidhi Balki	13	20	17/08/2020	
9	nikita dhengale	14	20	17/08/2020	
10	pooja vidhate	13	20	17/08/2020	
11	Priya ghorude	12	20	17/08/2020	
12	Saujanya Chukkwar	16	20	17/08/2020	
13	shivani patarange	15	20	17/08/2020	
14	somya singh	19	20	17/08/2020	
15	Swati Giradkar	15	20	17/08/2020	

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PRINCIPAL

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Bhadrawati, Dist-Cilandapur

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NS Science & Arts College

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N.S. Science & Arts College
Bhatlrawati, Dist-Chandrapur

# BAM - Sem X Sub - History

2v.	Name of students	Mark out of 20.
0	Bobacke Pranali	12_
(2)	Bodhe Kiran	13
(3)	Deharkas Utkarsha.	14
(a)	Dongase Pratiksha.	13
1	Girlpunse Divya.	12
6	Kakade Bharati	029
(F)	matte mayuri	08.
(8)	Nibude papita	16
0	Klagh Rashami	08.
(10)	Kalawah Khangar	09
1	Pimpalkas Vaishnavi	08.

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N.S. Science & Arts College Bhadrawati, Dist-Chandrapur (a) a harre

Prof. Kundan D. Shahare Department of History

N-S allege Bhadraidati

## UNIT Test Result

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0	Amane T. Pranalika.	14	
2	Bhoyar M. Prajakta.	13	
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(3)	Gaikwad Amol.	16	
6	Kakde C. Damini	12	
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8	Kosare A. Sweta.	13	
(9)	Kullermaie Varsha B.	0.8	
0	Milmile Karishama. B.	07	
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N.S. Science & Arts Coll

N.S. Science & Arts College Bhadrawati, Dist-Chandrapur प्रा कुहन दे काहार इतिहास विभाग प्रमुख नि ।शे विवक महा भूकावती

	UNIT Test = Sem I				
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Subject - Sociology,  Marks - 20
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4.3 "आप्रवादी COMMITTE क्याय भी देवायनी जिल्ला
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N.S. Science & Arts College  Bhadrawati, Dist-Chandrapur

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	नि चेहपूर
	Unit Test
	Class - B. M. II Sem. IV
	Subject - Sociology
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9.24	, युद्धालारामा अर्थ स्ट्रस्ट ३८७.
9.34	् लोडदेव्या भारीला उरा किहा.
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	Dr. L. S. Ladke
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Dr. L. S. Ladke PRINCIPAL N.S. Science & Arts College			3 (7 (7) 390 110		
Dr. L. S. Ladke  PRINCIPAL  N.S. Science & Arts College					
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N.S. Science & Arts College		Dr. L. S. Ladke			
Bhadrawati, Dist-Chandrapur		N.S. Science & Arte College			
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	NIl Kanthrao Shinde Science and I Bhadruwoon. Dist. Chomdo	art's college,
	Unit Test Result	
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1	ku. Patti Suryqwanshi	18
2	ku Pranali D. Bobade	17
3	raishagri a. Tarale	18
4	Sonaii O. Mondhe	16
5	Manash H. Kande	19
6	Mishau B. Jivane	15
(7	Manjala S. Taine	16
8	Sanket B. Lode	18
9	Savita O. Asekar	17
lo	yogest wars m. shende	17
	Dr.L. S. Ladke PRINCIPAL N.S. Science & Add College	January may
	Bhadrawan, Distremativi aper	िन्द्रागेर् लाहुजी अवरे

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## DEPARTMENT OF SOCIOLOGY

### CERTIFICATE

This is to certify that Shri	i/Ku./Smt. Akshatd
Shalik Umge	is a bonafide student of
Class BA and year Roll	NoHe/She has done
V	/ Good / Excellent Assignment work in
the subject Sociology	in Academic Session 2020 - 2021

Incharge Teacher

Head of the Deptt.

Principal

N. S. Sci. & Arts College, Bhadrawati

N. S. Sci. & Arts College, Bhadrawati

Dr. L. S. Ladke PRINCIPAL

N.S. Science & Arts College Shadrawati, Dist-Chandrapur Anand Screen, Bhadrawati 9850372295

**BHADRAWATI, DIST. CHANDRAPUR** 



# CERTIFICATE

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Class BA. TI Year Roll No.	He/She has done ood / Excellent Assignment work in
Unsatisfactory / Satisfactory / Go	n Academic Session 2000-901.

Incharge Teacher

Head of the Deptt.

Principal

N. S. Sci. & Arts College, Bhadrawatig

**BHADRAWATI, DIST. CHANDRAPUR** 



## DEPARTMENT OF CERTIFICATE

This is to certify that Shri / Ku. / Smt.	Honalisa	Bandy
Randoke	is a bona	fide student of
Class A A H Yeer Roll No.	Не	She has done
Unsatisfactory / Satisfactory / Good / Exc	cellent Assign	ment work in
the subject History in Acader	nic Session _	1020-21.

Dr. K. D. Shahare. Incharge Teacher

Head of the Deptt.

Principal

N. S. Sci. & Arts College, Bhadrawati

N. S. Sci. & Arts College, Bhadrawati

Anand Screen, Bhadrawati 9850372295

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## NILKANTHRAO SHINDE SCIENCE AND ARTS COLLEGE, CENTRE FOR HIGHER LEARNING AND RESEARCH BHADRAWATI, DIST. CHANDRAPUR



## DEPARTMENT OF \_ History

## CERTIFICATE

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Gairwas	is a bonafide student of
Class B. A. IT Roll No.	He/She has done
Unsatisfactory / Satisfactory / Good /	Excellent Assignment work in
the subject History in Acc	ademic Session 2020-21.

Dr. K.D Shahark Incharge Teacher

Head of the Deptt.

Principal

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N. S. Sci. & Arts College, Bhadrawati

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Bhadrawati, Dist-Chandrapur

BHADRAWATI, DIST. CHANDRAPUR



## Assignment / Seminar / Unit Test Book

#### ACADEMIC SESSION 2020 - 2021

Name of the Student: Nitu kum	an only Famay Jadan
Class: Bsc. 1 year 2 nd Jem	Section:
Subject: Chemistry	Roll No.
Univ. Enrolment No :	AUG. 110.

Dr. L. S. Ladke
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N.S. Science & Arts College
Bhadrawati, Dist-Chandrapur

BHADRAWATI, DIST. CHANDRAPUR



# Assignment / Seminar / Unit Test Book

ACADEMIC SESSION 2020 - 2024

Class: B. Sc-II Senn-4	Section: (PCM)
Subject: Chemistry Corganio	1
Univ. Enrolment No:	Roll No.

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Thidrawati, Dist-Chandrapur

BHADRAWATI, DIST. CHANDRAPUR



## Assignment / Seminar / Unit Test Book

ACADEMIC SESSION

Class: Bu - line ( Herr 1)	Section:	
Subject: Aumfing	Roll No.	9
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Dr. L. S. Ladke

Dr. L. S. Ladke

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N.S. Science & Arts College

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#### Nilkanthrao Shinde Science and Arts College, Bhadrawati Preliminary Examination-2021

Preliminary Exam was not taken during the session-2020-21 due to corona pandemic.

Dr. L. S. Ladke PRINCIPAL

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